

Solutions Manual for
Thermodynamics: An Engineering Approach
Seventh Edition
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Chapter 13

GAS MIXTURES

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Composition of Gas Mixtures

13-1C The ratio of the mass of a component to the mass of the mixture is called the mass fraction (mf), and the ratio of the mole number of a component to the mole number of the mixture is called the mole fraction (y).

13-2C The mass fractions will be identical, but the mole fractions will not.

13-3C Yes.

13-4C Yes, because both CO_2 and N_2O has the same molar mass, $M = 44$ kg/kmol.

13-5C No. We can do this only when each gas has the same mole fraction.

13-6C It is the average or the equivalent molar mass of the gas mixture. No.

13-7 From the definition of mass fraction,

$$mf_i = \frac{m_i}{m_m} = \frac{N_i M_i}{N_m M_m} = y_i \left(\frac{M_i}{M_m} \right)$$

13-8 A mixture consists of two gases. Relations for mole fractions when mass fractions are known are to be obtained .

Analysis The mass fractions of A and B are expressed as

$$\text{mf}_A = \frac{m_A}{m_m} = \frac{N_A M_A}{N_m M_m} = y_A \frac{M_A}{M_m} \quad \text{and} \quad \text{mf}_B = y_B \frac{M_B}{M_m}$$

Where m is mass, M is the molar mass, N is the number of moles, and y is the mole fraction. The apparent molar mass of the mixture is

$$M_m = \frac{m_m}{N_m} = \frac{N_A M_A + N_B M_B}{N_m} = y_A M_A + y_B M_B$$

Combining the two equation above and noting that $y_A + y_B = 1$ gives the following convenient relations for converting mass fractions to mole fractions,

$$y_A = \frac{M_B}{M_A(1/\text{mf}_A - 1) + M_B} \quad \text{and} \quad y_B = 1 - y_A$$

which are the desired relations.

13-9 The definitions for the mass fraction, weight, and the weight fractions are

$$(\text{mf})_i = \frac{m_i}{m_{\text{total}}}$$

$$W = mg$$

$$(\text{wf})_i = \frac{W_i}{W_{\text{total}}}$$

Since the total system consists of one mass unit, the mass of the i th component in this mixture is m_i . The weight of this one component is then

$$W_i = g(\text{mf})_i$$

Hence, the weight fraction for this one component is

$$(\text{wf})_i = \frac{g(\text{mf})_i}{\sum g(\text{mf})_i} = (\text{mf})_i$$

13-10E The moles of components of a gas mixture are given. The mole fractions and the apparent molecular weight are to be determined.

Properties The molar masses of He, O₂, N₂, and H₂O are 4.0, 32.0, 28.0 and 18.0 lbm/lbmol, respectively (Table A-1).

Analysis The total mole number of the mixture is

$$N_m = N_{\text{He}} + N_{\text{O}_2} + N_{\text{H}_2\text{O}} + N_{\text{N}_2} = 3 + 1.5 + 0.3 + 2.5 = 7.3 \text{ lbmol}$$

and the mole fractions are

$$y_{\text{He}} = \frac{N_{\text{He}}}{N_m} = \frac{3 \text{ lbmol}}{7.3 \text{ lbmol}} = \mathbf{0.411}$$

$$y_{\text{O}_2} = \frac{N_{\text{O}_2}}{N_m} = \frac{1.5 \text{ lbmol}}{7.3 \text{ lbmol}} = \mathbf{0.206}$$

$$y_{\text{H}_2\text{O}} = \frac{N_{\text{H}_2\text{O}}}{N_m} = \frac{0.3 \text{ lbmol}}{7.3 \text{ lbmol}} = \mathbf{0.0411}$$

$$y_{\text{N}_2} = \frac{N_{\text{N}_2}}{N_m} = \frac{2.5 \text{ lbmol}}{7.3 \text{ lbmol}} = \mathbf{0.343}$$

3 lbmol He 1.5 lbmol O ₂ 0.3 lbmol H ₂ O 2.5 lbmol N ₂
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The total mass of the mixture is

$$\begin{aligned}
m_m &= m_{\text{He}} + m_{\text{O}_2} + m_{\text{H}_2\text{O}} + m_{\text{N}_2} \\
&= N_{\text{He}}M_{\text{He}} + N_{\text{O}_2}M_{\text{O}_2} + N_{\text{H}_2\text{O}}M_{\text{H}_2\text{O}} + N_{\text{N}_2}M_{\text{N}_2} \\
&= (3 \text{ lbm})(4 \text{ lbm/lbmol}) + (1.5 \text{ lbm})(32 \text{ lbm/lbmol}) + (0.3 \text{ lbm})(18 \text{ lbm/lbmol}) + (2.5 \text{ lbm})(28 \text{ lbm/lbmol}) \\
&= 135.4 \text{ lbm}
\end{aligned}$$

Then the apparent molecular weight of the mixture becomes

$$M_m = \frac{m_m}{N_m} = \frac{135.4 \text{ lbm}}{7.3 \text{ lbmol}} = \mathbf{18.6 \text{ lbm/lbmol}}$$

13-11 The masses of the constituents of a gas mixture are given. The mass fractions, the mole fractions, the average molar mass, and gas constant are to be determined.

Properties The molar masses of O₂, N₂, and CO₂ are 32.0, 28.0 and 44.0 kg/kmol, respectively (Table A-1)

Analysis (a) The total mass of the mixture is

$$m_m = m_{\text{O}_2} + m_{\text{N}_2} + m_{\text{CO}_2} = 5 \text{ kg} + 8 \text{ kg} + 10 \text{ kg} = 23 \text{ kg}$$

Then the mass fraction of each component becomes

$$\text{mf}_{\text{O}_2} = \frac{m_{\text{O}_2}}{m_m} = \frac{5 \text{ kg}}{23 \text{ kg}} = \mathbf{0.217}$$

$$\text{mf}_{\text{N}_2} = \frac{m_{\text{N}_2}}{m_m} = \frac{8 \text{ kg}}{23 \text{ kg}} = \mathbf{0.348}$$

$$\text{mf}_{\text{CO}_2} = \frac{m_{\text{CO}_2}}{m_m} = \frac{10 \text{ kg}}{23 \text{ kg}} = \mathbf{0.435}$$

<p>5 kg O₂ 8 kg N₂ 10 kg CO₂</p>

(b) To find the mole fractions, we need to determine the mole numbers of each component first,

$$N_{\text{O}_2} = \frac{m_{\text{O}_2}}{M_{\text{O}_2}} = \frac{5 \text{ kg}}{32 \text{ kg/kmol}} = 0.156 \text{ kmol}$$

$$N_{\text{N}_2} = \frac{m_{\text{N}_2}}{M_{\text{N}_2}} = \frac{8 \text{ kg}}{28 \text{ kg/kmol}} = 0.286 \text{ kmol}$$

$$N_{\text{CO}_2} = \frac{m_{\text{CO}_2}}{M_{\text{CO}_2}} = \frac{10 \text{ kg}}{44 \text{ kg/kmol}} = 0.227 \text{ kmol}$$

Thus,

$$N_m = N_{\text{O}_2} + N_{\text{N}_2} + N_{\text{CO}_2} = 0.156 \text{ kmol} + 0.286 \text{ kmol} + 0.227 \text{ kmol} = 0.669 \text{ kmol}$$

and

$$y_{\text{O}_2} = \frac{N_{\text{O}_2}}{N_m} = \frac{0.156 \text{ kmol}}{0.669 \text{ kmol}} = \mathbf{0.233}$$

$$y_{\text{N}_2} = \frac{N_{\text{N}_2}}{N_m} = \frac{0.286 \text{ kmol}}{0.669 \text{ kmol}} = \mathbf{0.428}$$

$$y_{\text{CO}_2} = \frac{N_{\text{CO}_2}}{N_m} = \frac{0.227 \text{ kmol}}{0.669 \text{ kmol}} = \mathbf{0.339}$$

(c) The average molar mass and gas constant of the mixture are determined from their definitions:

$$M_m = \frac{m_m}{N_m} = \frac{23 \text{ kg}}{0.669 \text{ kmol}} = \mathbf{34.4 \text{ kg/kmol}}$$

and

$$R_m = \frac{R_u}{M_m} = \frac{8.314 \text{ kJ/kmol} \cdot \text{K}}{34.4 \text{ kg/kmol}} = \mathbf{0.242 \text{ kJ/kg} \cdot \text{K}}$$

13-12 The mass fractions of the constituents of a gas mixture are given. The mole fractions of the gas and gas constant are to be determined.

Properties The molar masses of CH₄, and CO₂ are 16.0 and 44.0 kg/kmol, respectively (Table A-1)

Analysis For convenience, consider 100 kg of the mixture. Then the number of moles of each component and the total number of moles are

$$m_{\text{CH}_4} = 75 \text{ kg} \longrightarrow N_{\text{CH}_4} = \frac{m_{\text{CH}_4}}{M_{\text{CH}_4}} = \frac{75 \text{ kg}}{16 \text{ kg/kmol}} = 4.688 \text{ kmol}$$

$$m_{\text{CO}_2} = 25 \text{ kg} \longrightarrow N_{\text{CO}_2} = \frac{m_{\text{CO}_2}}{M_{\text{CO}_2}} = \frac{25 \text{ kg}}{44 \text{ kg/kmol}} = 0.568 \text{ kmol}$$

$$N_m = N_{\text{CH}_4} + N_{\text{CO}_2} = 4.688 \text{ kmol} + 0.568 \text{ kmol} = 5.256 \text{ kmol}$$

mass
75% CH ₄
25% CO ₂

Then the mole fraction of each component becomes

$$y_{\text{CH}_4} = \frac{N_{\text{CH}_4}}{N_m} = \frac{4.688 \text{ kmol}}{5.256 \text{ kmol}} = 0.892 \text{ or } \mathbf{89.2\%}$$

$$y_{\text{CO}_2} = \frac{N_{\text{CO}_2}}{N_m} = \frac{0.568 \text{ kmol}}{5.256 \text{ kmol}} = 0.108 \text{ or } \mathbf{10.8\%}$$

The molar mass and the gas constant of the mixture are determined from their definitions,

$$M_m = \frac{m_m}{N_m} = \frac{100 \text{ kg}}{5.256 \text{ kmol}} = 19.03 \text{ kg/kmol}$$

and

$$R_m = \frac{R_u}{M_m} = \frac{8.314 \text{ kJ/kmol} \cdot \text{K}}{19.03 \text{ kg/kmol}} = \mathbf{0.437 \text{ kJ/kg} \cdot \text{K}}$$

13-13 The mole numbers of the constituents of a gas mixture are given. The mass of each gas and the apparent gas constant are to be determined.

Properties The molar masses of H₂, and N₂ are 2.0 and 28.0 kg/kmol, respectively (Table A-1)

Analysis The mass of each component is determined from

$$N_{\text{H}_2} = 5 \text{ kmol} \longrightarrow m_{\text{H}_2} = N_{\text{H}_2} M_{\text{H}_2} = (5 \text{ kmol})(2.0 \text{ kg/kmol}) = \mathbf{10 \text{ kg}}$$

$$N_{\text{N}_2} = 4 \text{ kmol} \longrightarrow m_{\text{N}_2} = N_{\text{N}_2} M_{\text{N}_2} = (4 \text{ kmol})(28 \text{ kg/kmol}) = \mathbf{112 \text{ kg}}$$

The total mass and the total number of moles are

$$m_m = m_{\text{H}_2} + m_{\text{N}_2} = 10 \text{ kg} + 112 \text{ kg} = 122 \text{ kg}$$

$$N_m = N_{\text{H}_2} + N_{\text{N}_2} = 5 \text{ kmol} + 4 \text{ kmol} = 9 \text{ kmol}$$

5 kmol H ₂
4 kmol N ₂

The molar mass and the gas constant of the mixture are determined from their definitions,

$$M_m = \frac{m_m}{N_m} = \frac{122 \text{ kg}}{9 \text{ kmol}} = 13.56 \text{ kg/kmol}$$

and

$$R_m = \frac{R_u}{M_m} = \frac{8.314 \text{ kJ/kmol} \cdot \text{K}}{13.56 \text{ kg/kmol}} = \mathbf{0.613 \text{ kJ/kg} \cdot \text{K}}$$

13-14 The mass fractions of the constituents of a gas mixture are given. The volumetric analysis of the mixture and the apparent gas constant are to be determined.

Properties The molar masses of O₂, N₂ and CO₂ are 32.0, 28, and 44.0 kg/kmol, respectively (Table A-1)

Analysis For convenience, consider 100 kg of the mixture. Then the number of moles of each component and the total number of moles are

$$m_{\text{O}_2} = 20 \text{ kg} \longrightarrow N_{\text{O}_2} = \frac{m_{\text{O}_2}}{M_{\text{O}_2}} = \frac{20 \text{ kg}}{32 \text{ kg/kmol}} = 0.625 \text{ kmol}$$

$$m_{\text{N}_2} = 20 \text{ kg} \longrightarrow N_{\text{N}_2} = \frac{m_{\text{N}_2}}{M_{\text{N}_2}} = \frac{30 \text{ kg}}{28 \text{ kg/kmol}} = 1.071 \text{ kmol}$$

$$m_{\text{CO}_2} = 50 \text{ kg} \longrightarrow N_{\text{CO}_2} = \frac{m_{\text{CO}_2}}{M_{\text{CO}_2}} = \frac{50 \text{ kg}}{44 \text{ kg/kmol}} = 1.136 \text{ kmol}$$

$$N_m = N_{\text{O}_2} + N_{\text{N}_2} + N_{\text{CO}_2} = 0.625 + 1.071 + 1.136 = 2.832 \text{ kmol}$$

mass
20% O ₂
30% N ₂
50% CO ₂

Noting that the volume fractions are same as the mole fractions, the volume fraction of each component becomes

$$y_{\text{O}_2} = \frac{N_{\text{O}_2}}{N_m} = \frac{0.625 \text{ kmol}}{2.832 \text{ kmol}} = 0.221 \text{ or } \mathbf{22.1\%}$$

$$y_{\text{N}_2} = \frac{N_{\text{N}_2}}{N_m} = \frac{1.071 \text{ kmol}}{2.832 \text{ kmol}} = 0.378 \text{ or } \mathbf{37.8\%}$$

$$y_{\text{CO}_2} = \frac{N_{\text{CO}_2}}{N_m} = \frac{1.136 \text{ kmol}}{2.832 \text{ kmol}} = 0.401 \text{ or } \mathbf{40.1\%}$$

The molar mass and the gas constant of the mixture are determined from their definitions,

$$M_m = \frac{m_m}{N_m} = \frac{100 \text{ kg}}{2.832 \text{ kmol}} = 35.31 \text{ kg/kmol}$$

and

$$R_m = \frac{R_u}{M_m} = \frac{8.314 \text{ kJ/kmol} \cdot \text{K}}{35.31 \text{ kg/kmol}} = \mathbf{0.235 \text{ kJ/kg} \cdot \text{K}}$$

***P-v-T* Behavior of Gas Mixtures**

13-15C Normally yes. Air, for example, behaves as an ideal gas in the range of temperatures and pressures at which oxygen and nitrogen behave as ideal gases.

13-16C The pressure of a gas mixture is equal to the sum of the pressures each gas would exert if existed alone at the mixture temperature and volume. This law holds exactly for ideal gas mixtures, but only approximately for real gas mixtures.

13-17C The volume of a gas mixture is equal to the sum of the volumes each gas would occupy if existed alone at the mixture temperature and pressure. This law holds exactly for ideal gas mixtures, but only approximately for real gas mixtures.

13-18C The P - v - T behavior of a component in an ideal gas mixture is expressed by the ideal gas equation of state using the properties of the individual component instead of the mixture, $P_i v_i = R_i T_i$. The P - v - T behavior of a component in a real gas mixture is expressed by more complex equations of state, or by $P_i v_i = Z_i R_i T_i$, where Z_i is the compressibility factor.

13-19C Component pressure is the pressure a component would exert if existed alone at the mixture temperature and volume. Partial pressure is the quantity $y_i P_m$, where y_i is the mole fraction of component i . These two are identical for ideal gases.

13-20C Component volume is the volume a component would occupy if existed alone at the mixture temperature and pressure. Partial volume is the quantity $y_i V_m$, where y_i is the mole fraction of component i . These two are identical for ideal gases.

13-21C The one with the highest mole number.

13-22C The partial pressures will decrease but the pressure fractions will remain the same.

13-23C The partial pressures will increase but the pressure fractions will remain the same.

13-24C No. The correct expression is “the volume of a gas mixture is equal to the sum of the volumes each gas would occupy if existed alone at the mixture temperature and pressure.”

13-25C No. The correct expression is “the temperature of a gas mixture is equal to the temperature of the individual gas components.”

13-26C Yes, it is correct.

13-27C With Kay's rule, a real-gas mixture is treated as a pure substance whose critical pressure and temperature are defined in terms of the critical pressures and temperatures of the mixture components as

$$P'_{cr,m} = \sum y_i P_{cr,i} \quad \text{and} \quad T'_{cr,m} = \sum y_i T_{cr,i}$$

The compressibility factor of the mixture (Z_m) is then easily determined using these pseudo-critical point values.

13-28 The partial pressure of R-134a in atmospheric air to form a 100-ppm contaminant is to be determined.

Analysis Noting that volume fractions and mole fractions are equal, the molar fraction of R-134a in air is

$$y_{R134a} = \frac{100}{10^6} = 0.0001$$

The partial pressure of R-134a in air is then

$$P_{R134a} = y_{R134a} P_m = (0.0001)(100 \text{ kPa}) = \mathbf{0.01 \text{ kPa}}$$

13-29 A tank contains a mixture of two gases of known masses at a specified pressure and temperature. The mixture is now heated to a specified temperature. The volume of the tank and the final pressure of the mixture are to be determined.

Assumptions Under specified conditions both Ar and N₂ can be treated as ideal gases, and the mixture as an ideal gas mixture.

Analysis The total number of moles is

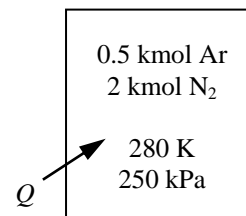
$$N_m = N_{Ar} + N_{N_2} = 0.5 \text{ kmol} + 2 \text{ kmol} = 2.5 \text{ kmol}$$

and

$$V_m = \frac{N_m R_u T_m}{P_m} = \frac{(2.5 \text{ kmol})(8.314 \text{ kPa} \cdot \text{m}^3/\text{kmol} \cdot \text{K})(280 \text{ K})}{250 \text{ kPa}} = \mathbf{23.3 \text{ m}^3}$$

Also,

$$\frac{P_2 V_2}{T_2} = \frac{P_1 V_1}{T_1} \longrightarrow P_2 = \frac{T_2}{T_1} P_1 = \frac{400 \text{ K}}{280 \text{ K}} (250 \text{ kPa}) = \mathbf{357.1 \text{ kPa}}$$



13-30 The volume fractions of components of a gas mixture are given. The mass fractions and apparent molecular weight of the mixture are to be determined.

Properties The molar masses of H_2 , He, and N_2 are 2.0, 4.0, and 28.0 kg/kmol, respectively (Table A-1).

Analysis We consider 100 kmol of this mixture. Noting that volume fractions are equal to the mole fractions, mass of each component are

$$\begin{aligned}m_{H_2} &= N_{H_2} M_{H_2} = (30 \text{ kmol})(2 \text{ kg/kmol}) = 60 \text{ kg} \\m_{He} &= N_{He} M_{He} = (40 \text{ kmol})(4 \text{ kg/kmol}) = 160 \text{ kg} \\m_{N_2} &= N_{N_2} M_{N_2} = (30 \text{ kmol})(28 \text{ kg/kmol}) = 840 \text{ kg}\end{aligned}$$

The total mass is

$$m_m = m_{H_2} + m_{He} + m_{N_2} = 60 + 160 + 840 = 1060 \text{ kg}$$

Then the mass fractions are

$$\begin{aligned}\text{mf}_{H_2} &= \frac{m_{H_2}}{m_m} = \frac{60 \text{ kg}}{1060 \text{ kg}} = \mathbf{0.05660} \\ \text{mf}_{He} &= \frac{m_{He}}{m_m} = \frac{160 \text{ kg}}{1060 \text{ kg}} = \mathbf{0.1509} \\ \text{mf}_{N_2} &= \frac{m_{N_2}}{m_m} = \frac{840 \text{ kg}}{1060 \text{ kg}} = \mathbf{0.7925}\end{aligned}$$

The apparent molecular weight of the mixture is

$$M_m = \frac{m_m}{N_m} = \frac{1060 \text{ kg}}{100 \text{ kmol}} = \mathbf{10.60 \text{ kg/kmol}}$$

<p>30% H_2 40% He 30% N_2 (by volume)</p>

13-31 The partial pressures of a gas mixture are given. The mole fractions, the mass fractions, the mixture molar mass, the apparent gas constant, the constant-volume specific heat, and the specific heat ratio are to be determined.

Properties The molar masses of CO_2 , O_2 and N_2 are 44.0, 32.0, and 28.0 kg/kmol, respectively (Table A-1). The constant-volume specific heats of these gases at 300 K are 0.657, 0.658, and 0.743 kJ/kg·K, respectively (Table A-2a).

Analysis The total pressure is

$$P_{\text{total}} = P_{\text{CO}_2} + P_{\text{O}_2} + P_{\text{N}_2} = 12.5 + 37.5 + 50 = 100 \text{ kPa}$$

The volume fractions are equal to the pressure fractions. Then,

$$y_{\text{CO}_2} = \frac{P_{\text{CO}_2}}{P_{\text{total}}} = \frac{12.5}{100} = \mathbf{0.125}$$

$$y_{\text{O}_2} = \frac{P_{\text{O}_2}}{P_{\text{total}}} = \frac{37.5}{100} = \mathbf{0.375}$$

$$y_{\text{N}_2} = \frac{P_{\text{N}_2}}{P_{\text{total}}} = \frac{50}{100} = \mathbf{0.50}$$

Partial
pressures
 CO_2 , 12.5 kPa
 O_2 , 37.5 kPa
 N_2 , 50 kPa

We consider 100 kmol of this mixture. Then the mass of each component are

$$m_{\text{CO}_2} = N_{\text{CO}_2} M_{\text{CO}_2} = (12.5 \text{ kmol})(44 \text{ kg/kmol}) = 550 \text{ kg}$$

$$m_{\text{O}_2} = N_{\text{O}_2} M_{\text{O}_2} = (37.5 \text{ kmol})(32 \text{ kg/kmol}) = 1200 \text{ kg}$$

$$m_{\text{N}_2} = N_{\text{N}_2} M_{\text{N}_2} = (50 \text{ kmol})(28 \text{ kg/kmol}) = 1400 \text{ kg}$$

The total mass is

$$m_m = m_{\text{N}_2} + m_{\text{O}_2} + m_{\text{Ar}} = 550 + 1200 + 1400 = 3150 \text{ kg}$$

Then the mass fractions are

$$\text{mf}_{\text{CO}_2} = \frac{m_{\text{CO}_2}}{m_m} = \frac{550 \text{ kg}}{3150 \text{ kg}} = \mathbf{0.1746}$$

$$\text{mf}_{\text{O}_2} = \frac{m_{\text{O}_2}}{m_m} = \frac{1200 \text{ kg}}{3150 \text{ kg}} = \mathbf{0.3810}$$

$$\text{mf}_{\text{N}_2} = \frac{m_{\text{N}_2}}{m_m} = \frac{1400 \text{ kg}}{3150 \text{ kg}} = \mathbf{0.4444}$$

The apparent molecular weight of the mixture is

$$M_m = \frac{m_m}{N_m} = \frac{3150 \text{ kg}}{100 \text{ kmol}} = \mathbf{31.50 \text{ kg/kmol}}$$

The constant-volume specific heat of the mixture is determined from

$$\begin{aligned} c_v &= \text{mf}_{\text{CO}_2} c_{v,\text{CO}_2} + \text{mf}_{\text{O}_2} c_{v,\text{O}_2} + \text{mf}_{\text{N}_2} c_{v,\text{N}_2} \\ &= 0.1746 \times 0.657 + 0.3810 \times 0.658 + 0.4444 \times 0.743 \\ &= \mathbf{0.6956 \text{ kJ/kg} \cdot \text{K}} \end{aligned}$$

The apparent gas constant of the mixture is

$$R = \frac{R_u}{M_m} = \frac{8.314 \text{ kJ/kmol} \cdot \text{K}}{31.50 \text{ kg/kmol}} = \mathbf{0.2639 \text{ kJ/kg} \cdot \text{K}}$$

The constant-pressure specific heat of the mixture and the specific heat ratio are

$$c_p = c_v + R = 0.6956 + 0.2639 = \mathbf{0.9595 \text{ kJ/kg} \cdot \text{K}}$$

$$k = \frac{c_p}{c_v} = \frac{0.9595 \text{ kJ/kg} \cdot \text{K}}{0.6956 \text{ kJ/kg} \cdot \text{K}} = \mathbf{1.379}$$

13-32 The mole numbers of combustion gases are given. The partial pressure of water vapor is to be determined.

Analysis The total mole of the mixture and the mole fraction of water vapor are

$$N_{\text{total}} = 0.75 + 1.66 + 5.65 = 8.06 \text{ kmol}$$

$$y_{\text{H}_2\text{O}} = \frac{N_{\text{H}_2\text{O}}}{N_{\text{total}}} = \frac{1.66}{8.06} = 0.2060$$

Noting that molar fraction is equal to pressure fraction, the partial pressure of water vapor is

$$P_{\text{H}_2\text{O}} = y_{\text{H}_2\text{O}} P_{\text{total}} = (0.2060)(101.3 \text{ kPa}) = \mathbf{20.86 \text{ kPa}}$$

13-33 An additional 5% of oxygen is mixed with standard atmospheric air. The molecular weight of this mixture is to be determined.

Properties The molar masses of N_2 and O_2 are 28.0 and 32.0 kg/kmol, respectively (Table A-1).

Analysis Standard air is taken as 79% nitrogen and 21% oxygen by mole. That is,

$$y_{\text{O}_2} = 0.21$$

$$y_{\text{N}_2} = 0.79$$

Adding another 0.05 moles of O_2 to 1 kmol of standard air gives

$$y_{\text{O}_2} = \frac{0.26}{1.05} = 0.2476$$

$$y_{\text{N}_2} = \frac{0.79}{1.05} = 0.7524$$

Then,

$$M_m = y_{\text{O}_2} M_{\text{O}_2} + y_{\text{N}_2} M_{\text{N}_2} = 0.2476 \times 32 + 0.7524 \times 28 = \mathbf{28.99 \text{ kg/kmol}}$$

13-34 The masses, temperatures, and pressures of two gases contained in two tanks connected to each other are given. The valve connecting the tanks is opened and the final temperature is measured. The volume of each tank and the final pressure are to be determined.

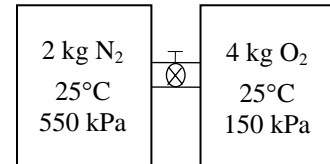
Assumptions Under specified conditions both N_2 and O_2 can be treated as ideal gases, and the mixture as an ideal gas mixture

Properties The molar masses of N_2 and O_2 are 28.0 and 32.0 kg/kmol, respectively. The gas constants of N_2 and O_2 are 0.2968 and 0.2598 kPa·m³/kg·K, respectively (Table A-1).

Analysis The volumes of the tanks are

$$V_{N_2} = \left(\frac{mRT}{P} \right)_{N_2} = \frac{(2 \text{ kg})(0.2968 \text{ kPa} \cdot \text{m}^3/\text{kg} \cdot \text{K})(298 \text{ K})}{550 \text{ kPa}} = \mathbf{0.322 \text{ m}^3}$$

$$V_{O_2} = \left(\frac{mRT}{P} \right)_{O_2} = \frac{(4 \text{ kg})(0.2598 \text{ kPa} \cdot \text{m}^3/\text{kg} \cdot \text{K})(298 \text{ K})}{150 \text{ kPa}} = \mathbf{2.065 \text{ m}^3}$$



$$V_{\text{total}} = V_{N_2} + V_{O_2} = 0.322 \text{ m}^3 + 2.065 \text{ m}^3 = 2.386 \text{ m}^3$$

Also,

$$N_{N_2} = \frac{m_{N_2}}{M_{N_2}} = \frac{2 \text{ kg}}{28 \text{ kg/kmol}} = 0.07143 \text{ kmol}$$

$$N_{O_2} = \frac{m_{O_2}}{M_{O_2}} = \frac{4 \text{ kg}}{32 \text{ kg/kmol}} = 0.125 \text{ kmol}$$

$$N_m = N_{N_2} + N_{O_2} = 0.07143 \text{ kmol} + 0.125 \text{ kmol} = 0.1964 \text{ kmol}$$

Thus,

$$P_m = \left(\frac{NR_u T}{V} \right)_m = \frac{(0.1964 \text{ kmol})(8.314 \text{ kPa} \cdot \text{m}^3/\text{kmol} \cdot \text{K})(298 \text{ K})}{2.386 \text{ m}^3} = \mathbf{204 \text{ kPa}}$$

13-35 The masses of components of a gas mixture are given. The apparent molecular weight of this mixture, the volume it occupies, the partial volume of the oxygen, and the partial pressure of the helium are to be determined.

Properties The molar masses of O₂, CO₂, and He are 32.0, 44.0, and 4.0 kg/kmol, respectively (Table A-1).

Analysis The total mass of the mixture is

$$m_m = m_{\text{O}_2} + m_{\text{CO}_2} + m_{\text{He}} = 0.1 + 1 + 0.5 = 1.6 \text{ kg}$$

The mole numbers of each component are

$$N_{\text{O}_2} = \frac{m_{\text{O}_2}}{M_{\text{O}_2}} = \frac{0.1 \text{ kg}}{32 \text{ kg/kmol}} = 0.003125 \text{ kmol}$$

$$N_{\text{CO}_2} = \frac{m_{\text{CO}_2}}{M_{\text{CO}_2}} = \frac{1 \text{ kg}}{44 \text{ kg/kmol}} = 0.02273 \text{ kmol}$$

$$N_{\text{He}} = \frac{m_{\text{He}}}{M_{\text{He}}} = \frac{0.5 \text{ kg}}{4 \text{ kg/kmol}} = 0.125 \text{ kmol}$$

0.1 kg O ₂ 1 kg CO ₂ 0.5 kg He
--

The mole number of the mixture is

$$N_m = N_{\text{O}_2} + N_{\text{CO}_2} + N_{\text{He}} = 0.003125 + 0.02273 + 0.125 = 0.15086 \text{ kmol}$$

Then the apparent molecular weight of the mixture becomes

$$M_m = \frac{m_m}{N_m} = \frac{1.6 \text{ kg}}{0.15086 \text{ kmol}} = \mathbf{10.61 \text{ kg/kmol}}$$

The volume of this ideal gas mixture is

$$V_m = \frac{N_m R_u T}{P} = \frac{(0.1509 \text{ kmol})(8.314 \text{ kPa} \cdot \text{m}^3/\text{kmol} \cdot \text{K})(300 \text{ K})}{100 \text{ kPa}} = \mathbf{3.764 \text{ m}^3}$$

The partial volume of oxygen in the mixture is

$$V_{\text{O}_2} = y_{\text{O}_2} V_m = \frac{N_{\text{O}_2}}{N_m} V_m = \frac{0.003125 \text{ kmol}}{0.1509 \text{ kmol}} (3.764 \text{ m}^3) = \mathbf{0.07795 \text{ m}^3}$$

The partial pressure of helium in the mixture is

$$P_{\text{He}} = y_{\text{He}} P_m = \frac{N_{\text{He}}}{N_m} P_m = \frac{0.125 \text{ kmol}}{0.1509 \text{ kmol}} (100 \text{ kPa}) = \mathbf{82.84 \text{ kPa}}$$

13-36 The mass fractions of components of a gas mixture are given. The volume occupied by 100 kg of this mixture is to be determined.

Properties The molar masses of CH_4 , C_3H_8 , and C_4H_{10} are 16.0, 44.0, and 58.0 kg/kmol, respectively (Table A-1).

Analysis The mole numbers of each component are

$$N_{\text{CH}_4} = \frac{m_{\text{CH}_4}}{M_{\text{CH}_4}} = \frac{60 \text{ kg}}{16 \text{ kg/kmol}} = 3.75 \text{ kmol}$$

$$N_{\text{C}_3\text{H}_8} = \frac{m_{\text{C}_3\text{H}_8}}{M_{\text{C}_3\text{H}_8}} = \frac{25 \text{ kg}}{44 \text{ kg/kmol}} = 0.5682 \text{ kmol}$$

$$N_{\text{C}_4\text{H}_{10}} = \frac{m_{\text{C}_4\text{H}_{10}}}{M_{\text{C}_4\text{H}_{10}}} = \frac{15 \text{ kg}}{58 \text{ kg/kmol}} = 0.2586 \text{ kmol}$$

60% CH_4 25% C_3H_8 15% C_4H_{10} (by mass)

The mole number of the mixture is

$$N_m = N_{\text{CH}_4} + N_{\text{C}_3\text{H}_8} + N_{\text{C}_4\text{H}_{10}} = 3.75 + 0.5682 + 0.2586 = 4.5768 \text{ kmol}$$

The apparent molecular weight of the mixture is

$$M_m = \frac{m_m}{N_m} = \frac{100 \text{ kg}}{4.5768 \text{ kmol}} = 21.85 \text{ kg/kmol}$$

Then the volume of this ideal gas mixture is

$$V_m = \frac{N_m R_u T}{P} = \frac{(4.5768 \text{ kmol})(8.314 \text{ kPa} \cdot \text{m}^3/\text{kmol} \cdot \text{K})(310 \text{ K})}{3000 \text{ kPa}} = \mathbf{3.93 \text{ m}^3}$$

13-37E The mass fractions of components of a gas mixture are given. The mass of 7 ft³ of this mixture and the partial volumes of the components are to be determined.

Properties The molar masses of N₂, O₂, and He are 28.0, 32.0, and 4.0 lbm/lbmol, respectively (Table A-1E).

Analysis We consider 100 lbm of this mixture for calculating the molar mass of the mixture. The mole numbers of each component are

$$N_{\text{N}_2} = \frac{m_{\text{N}_2}}{M_{\text{N}_2}} = \frac{45 \text{ lbm}}{28 \text{ lbm/lbmol}} = 1.607 \text{ lbmol}$$

$$N_{\text{O}_2} = \frac{m_{\text{O}_2}}{M_{\text{O}_2}} = \frac{35 \text{ lbm}}{32 \text{ lbm/lbmol}} = 1.094 \text{ lbmol}$$

$$N_{\text{He}} = \frac{m_{\text{He}}}{M_{\text{He}}} = \frac{20 \text{ lbm}}{4 \text{ lbm/lbmol}} = 5 \text{ lbmol}$$

<p style="text-align: center;">7 ft³ 45% N₂ 35% O₂ 20% He (by mass)</p>
--

The mole number of the mixture is

$$N_m = N_{\text{N}_2} + N_{\text{O}_2} + N_{\text{He}} = 1.607 + 1.094 + 5 = 7.701 \text{ lbmol}$$

The apparent molecular weight of the mixture is

$$M_m = \frac{m_m}{N_m} = \frac{100 \text{ lbm}}{7.701 \text{ lbmol}} = 12.99 \text{ lbm/lbmol}$$

Then the mass of this ideal gas mixture is

$$m = \frac{P \mathcal{V} M_m}{R_u T} = \frac{(300 \text{ psia})(7 \text{ ft}^3)(12.99 \text{ lbm/lbmol})}{(10.73 \text{ psia} \cdot \text{ft}^3/\text{lbmol} \cdot \text{R})(520 \text{ R})} = \mathbf{4.887 \text{ lbm}}$$

The mole fractions are

$$y_{\text{N}_2} = \frac{N_{\text{N}_2}}{N_m} = \frac{1.607 \text{ lbmol}}{7.701 \text{ lbmol}} = 0.2087$$

$$y_{\text{O}_2} = \frac{N_{\text{O}_2}}{N_m} = \frac{1.094 \text{ lbmol}}{7.701 \text{ lbmol}} = 0.142$$

$$y_{\text{He}} = \frac{N_{\text{He}}}{N_m} = \frac{5 \text{ lbmol}}{7.701 \text{ lbmol}} = 0.6493$$

Noting that volume fractions are equal to mole fractions, the partial volumes are determined from

$$\mathcal{V}_{\text{N}_2} = y_{\text{N}_2} \mathcal{V}_m = (0.2087)(7 \text{ ft}^3) = \mathbf{1.461 \text{ ft}^3}$$

$$\mathcal{V}_{\text{O}_2} = y_{\text{O}_2} \mathcal{V}_m = (0.142)(7 \text{ ft}^3) = \mathbf{0.994 \text{ ft}^3}$$

$$\mathcal{V}_{\text{He}} = y_{\text{He}} \mathcal{V}_m = (0.6493)(7 \text{ ft}^3) = \mathbf{4.545 \text{ ft}^3}$$

13-38 The mass fractions of components of a gas mixture are given. The partial pressure of ethane is to be determined.

Properties The molar masses of CH_4 and C_2H_6 are 16.0 and 30.0 kg/kmol, respectively (Table A-1).

Analysis We consider 100 kg of this mixture. The mole numbers of each component are

$$N_{\text{CH}_4} = \frac{m_{\text{CH}_4}}{M_{\text{CH}_4}} = \frac{70 \text{ kg}}{16 \text{ kg/kmol}} = 4.375 \text{ kmol}$$

$$N_{\text{C}_2\text{H}_6} = \frac{m_{\text{C}_2\text{H}_6}}{M_{\text{C}_2\text{H}_6}} = \frac{30 \text{ kg}}{30 \text{ kg/kmol}} = 1.0 \text{ kmol}$$

The mole number of the mixture is

$$N_m = N_{\text{CH}_4} + N_{\text{C}_2\text{H}_6} = 4.375 + 1.0 = 5.375 \text{ kmol}$$

The mole fractions are

$$y_{\text{CH}_4} = \frac{N_{\text{CH}_4}}{N_m} = \frac{4.375 \text{ kmol}}{5.375 \text{ kmol}} = 0.8139$$

$$y_{\text{C}_2\text{H}_6} = \frac{N_{\text{C}_2\text{H}_6}}{N_m} = \frac{1.0 \text{ kmol}}{5.375 \text{ kmol}} = 0.1861$$

The final pressure of ethane in the final mixture is

$$P_{\text{C}_2\text{H}_6} = y_{\text{C}_2\text{H}_6} P_m = (0.1861)(130 \text{ kPa}) = \mathbf{24.19 \text{ kPa}}$$

70% CH_4
30% C_2H_6
(by mass)
100 m^3
130 kPa, 25°C

13-39 A container contains a mixture of two fluids. The volume of the container and the total weight of its contents are to be determined.

Assumptions The volume of the mixture is the sum of the volumes of the two constituents.

Properties The specific volumes of the two fluids are given to be 0.001 m^3/kg and 0.008 m^3/kg .

Analysis The volumes of the two fluids are given by

$$V_A = m_A v_A = (1 \text{ kg})(0.001 \text{ m}^3/\text{kg}) = 0.001 \text{ m}^3$$

$$V_B = m_B v_B = (2 \text{ kg})(0.008 \text{ m}^3/\text{kg}) = 0.016 \text{ m}^3$$

The volume of the container is then

$$V = V_A + V_B = 0.001 + 0.016 = \mathbf{0.017 \text{ m}^3}$$

The total mass is

$$m = m_A + m_B = 1 + 2 = 3 \text{ kg}$$

and the weight of this mass will be

$$W = mg = (3 \text{ kg})(9.6 \text{ m/s}^2) = 28.8 \text{ kg} \cdot \text{m/s}^2 = \mathbf{28.8 \text{ N}}$$

1 kg fluid A
2 kg fluid B

13-40E A mixture consists of liquid water and another fluid. The specific weight of this mixture is to be determined.

Properties The densities of water and the fluid are given to be 62.4 lbf/ft^3 and 50.0 lbf/ft^3 , respectively.

Analysis We consider 1 ft^3 of this mixture. The volume of the water in the mixture is 0.7 ft^3 which has a mass of

$$m_w = \rho_w \mathcal{V}_w = (62.4 \text{ lbf/ft}^3)(0.7 \text{ ft}^3) = 43.68 \text{ lbf}$$

The weight of this water is

$$W_w = m_w g = (43.68 \text{ lbf})(31.9 \text{ ft/s}^2) \left(\frac{1 \text{ lbf}}{32.174 \text{ lbf} \cdot \text{ft/s}^2} \right) = 43.31 \text{ lbf}$$

0.7 ft³ water
0.3 ft³ fluid

Similarly, the volume of the second fluid is 0.3 ft^3 , and the mass of this fluid is

$$m_f = \rho_f \mathcal{V}_f = (50 \text{ lbf/ft}^3)(0.3 \text{ ft}^3) = 15 \text{ lbf}$$

The weight of the fluid is

$$W_f = m_f g = (15 \text{ lbf})(31.9 \text{ ft/s}^2) \left(\frac{1 \text{ lbf}}{32.174 \text{ lbf} \cdot \text{ft/s}^2} \right) = 14.87 \text{ lbf}$$

The specific weight of this mixture is then

$$\gamma = \frac{W_w + W_f}{\mathcal{V}_w + \mathcal{V}_f} = \frac{(43.31 + 14.87) \text{ lbf}}{(0.7 + 0.3) \text{ ft}^3} = \mathbf{58.2 \text{ lbf/ft}^3}$$

13-41 The mole fractions of components of a gas mixture are given. The mass flow rate of the mixture is to be determined.

Properties The molar masses of air and CH_4 are 28.97 and 16.0 kg/kmol, respectively (Table A-1).

Analysis The molar fraction of air is

$$y_{\text{air}} = 1 - y_{\text{CH}_4} = 1 - 0.15 = 0.85$$

The molar mass of the mixture is determined from

$$\begin{aligned} M_m &= y_{\text{CH}_4} M_{\text{CH}_4} + y_{\text{air}} M_{\text{air}} \\ &= 0.15 \times 16 + 0.85 \times 28.97 \\ &= 27.02 \text{ kg/kmol} \end{aligned}$$

15% CH_4
85% air

(by mole)

Given the engine displacement and speed and assuming that this is a 4-stroke engine (2 revolutions per cycle), the volume flow rate is determined from

$$\dot{\mathcal{V}} = \frac{\dot{n} \mathcal{V}_d}{2} = \frac{(3000 \text{ rev/min})(0.005 \text{ m}^3)}{2 \text{ rev/cycle}} = 7.5 \text{ m}^3/\text{min}$$

The specific volume of the mixture is

$$\nu = \frac{R_u T}{M_m P} = \frac{(8.314 \text{ kPa} \cdot \text{m}^3/\text{kmol} \cdot \text{K})(293 \text{ K})}{(27.02 \text{ kg/kmol})(80 \text{ kPa})} = 1.127 \text{ m}^3/\text{kg}$$

Hence the mass flow rate is

$$\dot{m} = \frac{\dot{\mathcal{V}}}{\nu} = \frac{7.5 \text{ m}^3/\text{min}}{1.127 \text{ m}^3/\text{kg}} = \mathbf{6.65 \text{ kg/min}}$$

13-42E The volumetric fractions of components of a natural gas mixture are given. The mass and volume flow rates of the mixture are to be determined.

Properties The molar masses of CH_4 and C_2H_6 are 16.0 and 30.0 lbm/lbmol, respectively (Table A-1E).

Analysis The molar mass of the mixture is determined from

$$M_m = y_{\text{CH}_4}M_{\text{CH}_4} + y_{\text{C}_2\text{H}_6}M_{\text{C}_2\text{H}_6} = 0.95 \times 16 + 0.05 \times 30 = 16.70 \text{ lbm/lbmol}$$

The specific volume of the mixture is

$$v = \frac{R_u T}{M_m P} = \frac{(10.73 \text{ psia} \cdot \text{ft}^3/\text{lbmol} \cdot \text{R})(520 \text{ R})}{(16.70 \text{ lbm/lbmol})(100 \text{ psia})} = 3.341 \text{ ft}^3/\text{lbm}$$


→	95% CH_4 5% C_2H_6 (by volume)
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The volume flow rate is

$$\dot{V} = AV = \frac{\pi D^2}{4} V = \frac{\pi (36/12 \text{ ft})^2}{4} (10 \text{ ft/s}) = \mathbf{70.69 \text{ ft}^3/\text{s}}$$

and the mass flow rate is

$$\dot{m} = \frac{\dot{V}}{v} = \frac{70.69 \text{ ft}^3/\text{s}}{3.341 \text{ ft}^3/\text{lbm}} = \mathbf{21.16 \text{ lbm/s}}$$

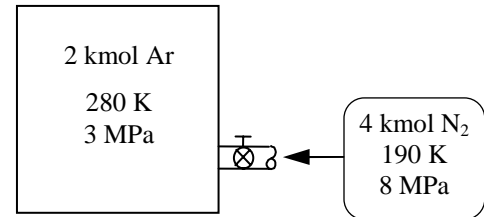
13-43  The mole numbers, temperatures, and pressures of two gases forming a mixture are given. The final temperature is also given. The pressure of the mixture is to be determined using two methods.

Analysis (a) Under specified conditions both Ar and N₂ will considerably deviate from the ideal gas behavior. Treating the mixture as an ideal gas,

$$\left. \begin{array}{l} \text{Initial state: } P_1 V_1 = N_1 R_u T_1 \\ \text{Final state: } P_2 V_2 = N_2 R_u T_2 \end{array} \right\} P_2 = \frac{N_2 T_2}{N_1 T_1} P_1 = \frac{(6)(230 \text{ K})}{(2)(280 \text{ K})} (3 \text{ MPa}) = \mathbf{7.39 \text{ MPa}}$$

(b) Initially,

$$\left. \begin{array}{l} T_R = \frac{T_1}{T_{cr,Ar}} = \frac{280 \text{ K}}{151.0 \text{ K}} = 1.854 \\ P_R = \frac{P_1}{P_{cr,Ar}} = \frac{3 \text{ MPa}}{4.86 \text{ MPa}} = 0.6173 \end{array} \right\} Z_{Ar} = 0.985 \text{ (Fig. A-15 or EES)}$$



Then the volume of the tank is

$$V = \frac{Z N_{Ar} R_u T}{P} = \frac{(0.985)(2 \text{ kmol})(8.314 \text{ kPa} \cdot \text{m}^3/\text{kmol} \cdot \text{K})(280 \text{ K})}{3000 \text{ kPa}} = 1.529 \text{ m}^3$$

After mixing,

$$\left. \begin{array}{l} T_{R,Ar} = \frac{T_m}{T_{cr,Ar}} = \frac{230 \text{ K}}{151.0 \text{ K}} = 1.523 \\ \text{Ar: } v_{R,Ar} = \frac{v_{Ar}}{R_u T_{cr,Ar} / P_{cr,Ar}} = \frac{v_m / N_{Ar}}{R_u T_{cr,Ar} / P_{cr,Ar}} \\ \quad = \frac{(1.529 \text{ m}^3)/(2 \text{ kmol})}{(8.314 \text{ kPa} \cdot \text{m}^3/\text{kmol} \cdot \text{K})(151.0 \text{ K})/(4860 \text{ kPa})} = 2.96 \end{array} \right\} P_R = 0.496 \text{ (Fig. A-15 or EES)}$$

$$\left. \begin{array}{l} T_{R,N_2} = \frac{T_m}{T_{cr,N_2}} = \frac{230 \text{ K}}{126.2 \text{ K}} = 1.823 \\ \text{N}_2: v_{R,N_2} = \frac{v_{N_2}}{R_u T_{cr,N_2} / P_{cr,N_2}} = \frac{v_m / N_{N_2}}{R_u T_{cr,N_2} / P_{cr,N_2}} \\ \quad = \frac{(1.529 \text{ m}^3)/(4 \text{ kmol})}{(8.314 \text{ kPa} \cdot \text{m}^3/\text{kmol} \cdot \text{K})(126.2 \text{ K})/(3390 \text{ kPa})} = 1.235 \end{array} \right\} P_R = 1.43 \text{ (Fig. A-15 or EES)}$$

Thus,

$$P_{Ar} = (P_R P_{cr})_{Ar} = (0.496)(4.86 \text{ MPa}) = 2.41 \text{ MPa}$$

$$P_{N_2} = (P_R P_{cr})_{N_2} = (1.43)(3.39 \text{ MPa}) = 4.85 \text{ MPa}$$

and

$$P_m = P_{Ar} + P_{N_2} = 2.41 \text{ MPa} + 4.85 \text{ MPa} = \mathbf{7.26 \text{ MPa}}$$



13-44 Problem 13-43 is reconsidered. The effect of the moles of nitrogen supplied to the tank on the final pressure of the mixture is to be studied using the ideal-gas equation of state and the compressibility chart with Dalton's law.

Analysis The problem is solved using EES, and the solution is given below.

"Input Data"

$R_u = 8.314$ [kJ/kmol-K] "universal Gas Constant"

$T_{Ar} = 280$ [K]

$P_{Ar} = 3000$ [kPa] "Pressure for only Argon in the tank initially."

$N_{Ar} = 2$ [kmol]

{ $N_{N2} = 4$ [kmol]}

$T_{mix} = 230$ [K]

$T_{cr_Ar} = 151.0$ [K] "Critical Constants are found in Table A.1 of the text"

$P_{cr_Ar} = 4860$ [kPa]

$T_{cr_N2} = 126.2$ [K]

$P_{cr_N2} = 3390$ [kPa]

"Ideal-gas Solution:"

$P_{Ar} * V_{Tank_IG} = N_{Ar} * R_u * T_{Ar}$ "Apply the ideal gas law the gas in the tank."

$P_{mix_IG} * V_{Tank_IG} = N_{mix} * R_u * T_{mix}$ "Ideal-gas mixture pressure"

$N_{mix} = N_{Ar} + N_{N2}$ "Moles of mixture"

"Real Gas Solution:"

$P_{Ar} * V_{Tank_RG} = Z_{Ar_1} * N_{Ar} * R_u * T_{Ar}$ "Real gas volume of tank"

$T_R = T_{Ar} / T_{cr_Ar}$ "Initial reduced Temp. of Ar"

$P_R = P_{Ar} / P_{cr_Ar}$ "Initial reduced Press. of Ar"

$Z_{Ar_1} = \text{COMPRESS}(T_R, P_R)$ "Initial compressibility factor for Ar"

$P_{Ar_mix} * V_{Tank_RG} = Z_{Ar_mix} * N_{Ar} * R_u * T_{mix}$ "Real gas Ar Pressure in mixture"

$T_{R_Ar_mix} = T_{mix} / T_{cr_Ar}$ "Reduced Temp. of Ar in mixture"

$P_{R_Ar_mix} = P_{Ar_mix} / P_{cr_Ar}$ "Reduced Press. of Ar in mixture"

$Z_{Ar_mix} = \text{COMPRESS}(T_{R_Ar_mix}, P_{R_Ar_mix})$ "Compressibility factor for Ar in mixture"

$P_{N2_mix} * V_{Tank_RG} = Z_{N2_mix} * N_{N2} * R_u * T_{mix}$ "Real gas N2 Pressure in mixture"

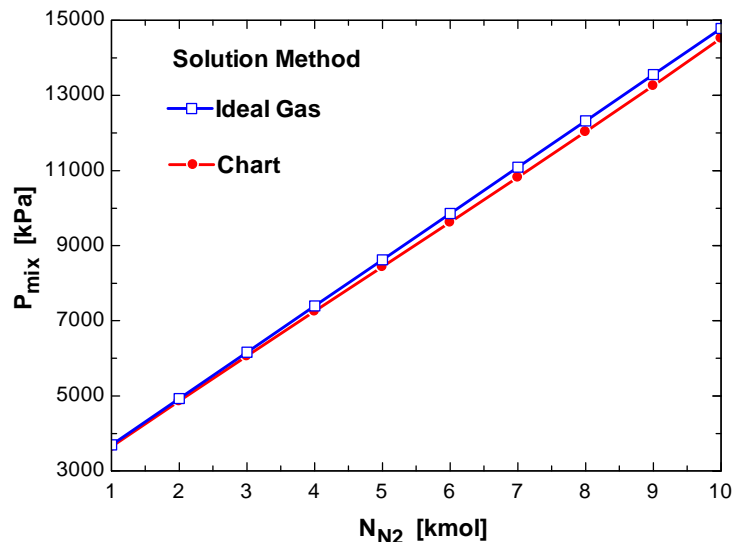
$T_{R_N2_mix} = T_{mix} / T_{cr_N2}$ "Reduced Temp. of N2 in mixture"

$P_{R_N2_mix} = P_{N2_mix} / P_{cr_N2}$ "Reduced Press. of N2 in mixture"

$Z_{N2_mix} = \text{COMPRESS}(T_{R_N2_mix}, P_{R_N2_mix})$ "Compressibility factor for N2 in mixture"

$P_{mix} = P_{R_Ar_mix} * P_{cr_Ar} + P_{R_N2_mix} * P_{cr_N2}$ "Mixture pressure by Dalton's law. 23800"

N_{N2} [kmol]	P_{mix} [kPa]	$P_{mix,IG}$ [kPa]
1	3647	3696
2	4863	4929
3	6063	6161
4	7253	7393
5	8438	8625
6	9626	9857
7	10822	11089
8	12032	12321
9	13263	13554
10	14521	14786



13-45E The mass fractions of gases forming a mixture at a specified pressure and temperature are given. The mass of the gas mixture is to be determined using four methods.

Properties The molar masses of CH_4 and C_2H_6 are 16.0 and 30.0 lbm/lbmol, respectively (Table A-1E).

Analysis (a) We consider 100 lbm of this mixture. Then the mole numbers of each component are

$$N_{\text{CH}_4} = \frac{m_{\text{CH}_4}}{M_{\text{CH}_4}} = \frac{75 \text{ lbm}}{16 \text{ lbm/lbmol}} = 4.6875 \text{ lbmol}$$

$$N_{\text{C}_2\text{H}_6} = \frac{m_{\text{C}_2\text{H}_6}}{M_{\text{C}_2\text{H}_6}} = \frac{25 \text{ lbm}}{30 \text{ lbm/lbmol}} = 0.8333 \text{ lbmol}$$

The mole number of the mixture and the mole fractions are

$$N_m = 4.6875 + 0.8333 = 5.5208 \text{ lbmol}$$

$$y_{\text{CH}_4} = \frac{N_{\text{CH}_4}}{N_m} = \frac{4.6875 \text{ lbmol}}{5.5208 \text{ lbmol}} = 0.8491$$

$$y_{\text{C}_2\text{H}_6} = \frac{N_{\text{C}_2\text{H}_6}}{N_m} = \frac{0.8333 \text{ lbmol}}{5.5208 \text{ lbmol}} = 0.1509$$

75% CH_4
25% C_2H_6
(by mass)
2000 psia
300°F

Then the apparent molecular weight of the mixture becomes

$$M_m = \frac{m_m}{N_m} = \frac{100 \text{ lbm}}{5.5208 \text{ lbmol}} = 18.11 \text{ lbm/lbmol}$$

The apparent gas constant of the mixture is

$$R = \frac{R_u}{M_m} = \frac{10.73 \text{ psia} \cdot \text{ft}^3/\text{lbmol} \cdot \text{R}}{18.11 \text{ lbm/lbmol}} = 0.5925 \text{ psia} \cdot \text{ft}^3/\text{lbm} \cdot \text{R}$$

The mass of this mixture in a 1 million ft^3 tank is

$$m = \frac{P\mathcal{V}}{RT} = \frac{(2000 \text{ psia})(1 \times 10^6 \text{ ft}^3)}{(0.5925 \text{ psia} \cdot \text{ft}^3/\text{lbm} \cdot \text{R})(760 \text{ R})} = \mathbf{4.441 \times 10^6 \text{ lbm}}$$

(b) To use the Amagat's law for this real gas mixture, we first need the compressibility factor of each component at the mixture temperature and pressure. The compressibility factors are obtained using Fig. A-15 to be

$$\left. \begin{aligned} T_{R,\text{CH}_4} &= \frac{T_m}{T_{\text{cr},\text{CH}_4}} = \frac{760 \text{ R}}{343.9 \text{ R}} = 2.210 \\ P_{R,\text{CH}_4} &= \frac{P_m}{P_{\text{cr},\text{CH}_4}} = \frac{2000 \text{ psia}}{673 \text{ psia}} = 2.972 \end{aligned} \right\} Z_{\text{CH}_4} = 0.98$$

$$\left. \begin{aligned} T_{R,\text{C}_2\text{H}_6} &= \frac{760 \text{ R}}{549.8 \text{ R}} = 1.382 \\ P_{R,\text{C}_2\text{H}_6} &= \frac{1500 \text{ psia}}{708 \text{ psia}} = 2.119 \end{aligned} \right\} Z_{\text{C}_2\text{H}_6} = 0.77$$

Then,

$$Z_m = \sum y_i Z_i = y_{\text{CH}_4} Z_{\text{CH}_4} + y_{\text{C}_2\text{H}_6} Z_{\text{C}_2\text{H}_6} = (0.8491)(0.98) + (0.1509)(0.77) = 0.9483$$

$$m = \frac{P\mathcal{V}}{Z_m RT} = \frac{(2000 \text{ psia})(1 \times 10^6 \text{ ft}^3)}{(0.9483)(0.5925 \text{ psia} \cdot \text{ft}^3/\text{lbm} \cdot \text{R})(760 \text{ R})} = \mathbf{4.684 \times 10^6 \text{ lbm}}$$

(c) To use Dalton's law with compressibility factors: (Fig. A-15)

$$T_{R,\text{CH}_4} = 2.210$$

$$\left. v_{R,\text{CH}_4} = \frac{v_m/m_{\text{CH}_4}}{R_{\text{CH}_4} T_{\text{cr},\text{CH}_4} / P_{\text{cr},\text{CH}_4}} = \frac{(1 \times 10^6 \text{ ft}^3)/(4.441 \times 10^6 \times 0.75 \text{ lbm})}{(0.6688 \text{ psia} \cdot \text{ft}^3/\text{lbm} \cdot \text{R})(343.9 \text{ R})/(673 \text{ psia})} = 0.8782 \right\} Z_{\text{CH}_4} = 0.98$$

$$\left. \begin{aligned} T_{R,C_2H_6} &= 1.382 \\ \nu_{R,C_2H_6} &= \frac{\mathcal{V}_m/m_{C_2H_6}}{RT_{cr,C_2H_6}/P_{cr,C_2H_6}} = \frac{(1 \times 10^6 \text{ ft}^3)/(4.441 \times 10^6 \times 0.25 \text{ lbm})}{(0.3574 \text{ psia} \cdot \text{ft}^3/\text{lbm} \cdot \text{R})(549.8 \text{ R})/(708 \text{ psia})} = 3.244 \end{aligned} \right\} Z_{CH_4} = 0.92$$

Note that we used $m = 4.441 \times 10^6 \times 0.25 \text{ lbm}$ in above calculations, the value obtained by ideal gas behavior. The solution normally requires iteration until the assumed and calculated mass values match. The mass of the component gas is obtained by multiplying the mass of the mixture by its mass fraction. Then,

$$Z_m = \sum y_i Z_i = y_{CH_4} Z_{CH_4} + y_{C_2H_6} Z_{C_2H_6} = (0.8491)(0.98) + (0.1509)(0.92) = 0.9709$$

$$m = \frac{P\mathcal{V}}{Z_m RT} = \frac{(2000 \text{ psia})(1 \times 10^6 \text{ ft}^3)}{(0.9709)(0.5925 \text{ psia} \cdot \text{ft}^3/\text{lbm} \cdot \text{R})(760 \text{ R})} = \mathbf{4.575 \times 10^6 \text{ lbm}}$$

This mass is sufficiently close to the assumed mass value of $4.441 \times 10^6 \times 0.25 \text{ lbm}$. Therefore, there is no need to repeat the calculations at this calculated mass.

(d) To use Kay's rule, we need to determine the pseudo-critical temperature and pseudo-critical pressure of the mixture using the critical point properties of gases.

$$\begin{aligned} T'_{cr,m} &= \sum y_i T_{cr,i} = y_{CH_4} T_{cr,CH_4} + y_{C_2H_6} T_{cr,C_2H_6} \\ &= (0.8491)(343.9 \text{ R}) + (0.1509)(549.8 \text{ R}) = 375.0 \text{ R} \end{aligned}$$

$$\begin{aligned} P'_{cr,m} &= \sum y_i P_{cr,i} = y_{CH_4} P_{cr,CH_4} + y_{C_2H_6} P_{cr,C_2H_6} \\ &= (0.8491)(673 \text{ psia}) + (0.1509)(708 \text{ psia}) = 678.3 \text{ psia} \end{aligned}$$

Then,

$$\left. \begin{aligned} T_R &= \frac{T_m}{T'_{cr,m}} = \frac{760 \text{ R}}{375.0 \text{ R}} = 2.027 \\ P_R &= \frac{P_m}{P'_{cr,m}} = \frac{2000 \text{ psia}}{678.3 \text{ psia}} = 2.949 \end{aligned} \right\} Z_m = 0.97 \quad (\text{Fig. A-15})$$

$$m = \frac{P\mathcal{V}}{Z_m RT} = \frac{(2000 \text{ psia})(1 \times 10^6 \text{ ft}^3)}{(0.97)(0.5925 \text{ psia} \cdot \text{ft}^3/\text{lbm} \cdot \text{R})(760 \text{ R})} = \mathbf{4.579 \times 10^6 \text{ lbm}}$$

13-46 The volumetric analysis of a mixture of gases is given. The volumetric and mass flow rates are to be determined using three methods.

Properties The molar masses of O₂, N₂, CO₂, and CH₄ are 32.0, 28.0, 44.0, and 16.0 kg/kmol, respectively (Table A-1).

Analysis (a) We consider 100 kmol of this mixture. Noting that volume fractions are equal to the mole fractions, mass of each component are

$$\begin{aligned} m_{\text{O}_2} &= N_{\text{O}_2} M_{\text{O}_2} = (30 \text{ kmol})(32 \text{ kg/kmol}) = 960 \text{ kg} \\ m_{\text{N}_2} &= N_{\text{N}_2} M_{\text{N}_2} = (40 \text{ kmol})(28 \text{ kg/kmol}) = 1120 \text{ kg} \\ m_{\text{CO}_2} &= N_{\text{CO}_2} M_{\text{CO}_2} = (10 \text{ kmol})(44 \text{ kg/kmol}) = 440 \text{ kg} \\ m_{\text{CH}_4} &= N_{\text{CH}_4} M_{\text{CH}_4} = (20 \text{ kmol})(16 \text{ kg/kmol}) = 320 \text{ kg} \end{aligned}$$

30% O₂
 40% N₂
 10% CO₂
 20% CH₄
 (by volume)

The total mass is

$$\begin{aligned} m_m &= m_{\text{O}_2} + m_{\text{N}_2} + m_{\text{CO}_2} + m_{\text{CH}_4} \\ &= 960 + 1120 + 440 + 320 = 2840 \text{ kg} \end{aligned}$$

The apparent molecular weight of the mixture is

$$M_m = \frac{m_m}{N_m} = \frac{2840 \text{ kg}}{100 \text{ kmol}} = 28.40 \text{ kg/kmol}$$

→ Mixture
 8 MPa, 15°C

The apparent gas constant of the mixture is

$$R = \frac{R_u}{M_m} = \frac{8.314 \text{ kJ/kmol} \cdot \text{K}}{28.40 \text{ kg/kmol}} = 0.2927 \text{ kJ/kg} \cdot \text{K}$$

The specific volume of the mixture is

$$\nu = \frac{RT}{P} = \frac{(0.2927 \text{ kPa} \cdot \text{m}^3/\text{kg} \cdot \text{K})(288 \text{ K})}{8000 \text{ kPa}} = 0.01054 \text{ m}^3/\text{kg}$$

The volume flow rate is

$$\dot{V} = AV = \frac{\pi D^2}{4} V = \frac{\pi(0.02 \text{ m})^2}{4} (3 \text{ m/s}) = \mathbf{0.0009425 \text{ m}^3/\text{s}}$$

and the mass flow rate is

$$\dot{m} = \frac{\dot{V}}{\nu} = \frac{0.0009425 \text{ m}^3/\text{s}}{0.01054 \text{ m}^3/\text{kg}} = \mathbf{0.08942 \text{ kg/s}}$$

(b) To use the Amagat's law for this real gas mixture, we first need the mole fractions and the Z of each component at the mixture temperature and pressure. The compressibility factors are obtained using Fig. A-15 to be

$$\left. \begin{aligned} T_{R,\text{O}_2} &= \frac{T_m}{T_{\text{cr},\text{O}_2}} = \frac{288 \text{ K}}{154.8 \text{ K}} = 1.860 \\ P_{R,\text{O}_2} &= \frac{P_m}{P_{\text{cr},\text{O}_2}} = \frac{8 \text{ MPa}}{5.08 \text{ MPa}} = 1.575 \end{aligned} \right\} Z_{\text{O}_2} = 0.95$$

$$\left. \begin{aligned} T_{R,\text{N}_2} &= \frac{288 \text{ K}}{126.2 \text{ K}} = 2.282 \\ P_{R,\text{N}_2} &= \frac{8 \text{ MPa}}{3.39 \text{ MPa}} = 2.360 \end{aligned} \right\} Z_{\text{N}_2} = 0.99$$

$$\left. \begin{aligned} T_{R,\text{CO}_2} &= \frac{288 \text{ K}}{304.2 \text{ K}} = 0.947 \\ P_{R,\text{CO}_2} &= \frac{8 \text{ MPa}}{7.39 \text{ MPa}} = 1.083 \end{aligned} \right\} Z_{\text{CO}_2} = 0.199$$

$$\left. \begin{aligned} T_{R,\text{CH}_4} &= \frac{288 \text{ K}}{191.1 \text{ K}} = 1.507 \\ P_{R,\text{CH}_4} &= \frac{8 \text{ MPa}}{4.64 \text{ MPa}} = 1.724 \end{aligned} \right\} Z_{\text{CH}_4} = 0.85$$

and

$$\begin{aligned} Z_m &= \sum y_i Z_i = y_{\text{O}_2} Z_{\text{O}_2} + y_{\text{N}_2} Z_{\text{N}_2} + y_{\text{CO}_2} Z_{\text{CO}_2} + y_{\text{CH}_4} Z_{\text{CH}_4} \\ &= (0.30)(0.95) + (0.40)(0.99) + (0.10)(0.199) + (0.20)(0.85) = 0.8709 \end{aligned}$$

Then,

$$\nu = \frac{Z_m RT}{P} = \frac{(0.8709)(0.2927 \text{ kPa} \cdot \text{m}^3/\text{kg} \cdot \text{K})(288 \text{ K})}{8000 \text{ kPa}} = 0.009178 \text{ m}^3/\text{kg}$$

$$\dot{V} = \mathbf{0.0009425 \text{ m}^3/\text{s}}$$

$$\dot{m} = \frac{\dot{V}}{\nu} = \frac{0.0009425 \text{ m}^3/\text{s}}{0.009178 \text{ m}^3/\text{kg}} = \mathbf{0.10269 \text{ kg/s}}$$

(c) To use Kay's rule, we need to determine the pseudo-critical temperature and pseudo-critical pressure of the mixture using the critical point properties of mixture gases.

$$\begin{aligned} T'_{cr,m} &= \sum y_i T_{cr,i} = y_{O_2} T_{cr,O_2} + y_{N_2} T_{cr,N_2} + y_{CO_2} T_{cr,CO_2} + y_{CH_4} T_{cr,CH_4} \\ &= (0.30)(154.8 \text{ K}) + (0.40)(126.2 \text{ K}) + (0.10)(304.2 \text{ K}) + (0.20)(191.1 \text{ K}) = 165.6 \text{ K} \end{aligned}$$

$$\begin{aligned} P'_{cr,m} &= \sum y_i P_{cr,i} = y_{O_2} P_{cr,O_2} + y_{N_2} P_{cr,N_2} + y_{CO_2} P_{cr,CO_2} + y_{CH_4} P_{cr,CH_4} \\ &= (0.30)(5.08 \text{ MPa}) + (0.40)(3.39 \text{ MPa}) + (0.10)(7.39 \text{ MPa}) + (0.20)(4.64 \text{ MPa}) = 4.547 \text{ MPa} \end{aligned}$$

and

$$\left. \begin{aligned} T_R &= \frac{T_m}{T'_{cr,m}} = \frac{288 \text{ K}}{165.6 \text{ K}} = 1.739 \\ P_R &= \frac{P_m}{P'_{cr,m}} = \frac{8 \text{ MPa}}{4.547 \text{ MPa}} = 1.759 \end{aligned} \right\} Z_m = 0.92 \quad (\text{Fig. A-15})$$

Then,

$$\nu = \frac{Z_m RT}{P} = \frac{(0.92)(0.2927 \text{ kPa} \cdot \text{m}^3/\text{kg} \cdot \text{K})(288 \text{ K})}{8000 \text{ kPa}} = 0.009694 \text{ m}^3/\text{kg}$$

$$\dot{V} = \mathbf{0.0009425 \text{ m}^3/\text{s}}$$

$$\dot{m} = \frac{\dot{V}}{\nu} = \frac{0.0009425 \text{ m}^3/\text{s}}{0.009694 \text{ m}^3/\text{kg}} = \mathbf{0.009723 \text{ kg/s}}$$

Properties of Gas Mixtures

13-47C Yes. Yes (extensive property).

13-48C No (intensive property).

13-49C The answers are the same for entropy.

13-50C Yes. Yes (conservation of energy).

13-51C We have to use the partial pressure.

13-52C No, this is an approximate approach. It assumes a component behaves as if it existed alone at the mixture temperature and pressure (i.e., it disregards the influence of dissimilar molecules on each other.)

13-53 The volume fractions of components of a gas mixture are given. This mixture is heated while flowing through a tube at constant pressure. The heat transfer to the mixture per unit mass of the mixture is to be determined.

Assumptions All gases will be modeled as ideal gases with constant specific heats.

Properties The molar masses of O_2 , N_2 , CO_2 , and CH_4 are 32.0, 28.0, 44.0, and 16.0 kg/kmol, respectively (Table A-1). The constant-pressure specific heats of these gases at room temperature are 0.918, 1.039, 0.846, and 2.2537 kJ/kg·K, respectively (Table A-2a).

Analysis We consider 100 kmol of this mixture. Noting that volume fractions are equal to the mole fractions, mass of each component are

$$m_{O_2} = N_{O_2} M_{O_2} = (30 \text{ kmol})(32 \text{ kg/kmol}) = 960 \text{ kg}$$

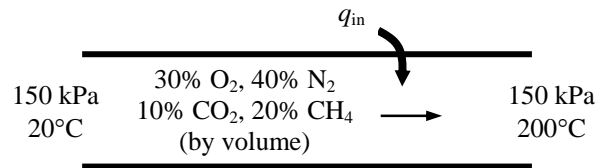
$$m_{N_2} = N_{N_2} M_{N_2} = (40 \text{ kmol})(28 \text{ kg/kmol}) = 1120 \text{ kg}$$

$$m_{CO_2} = N_{CO_2} M_{CO_2} = (10 \text{ kmol})(44 \text{ kg/kmol}) = 440 \text{ kg}$$

$$m_{CH_4} = N_{CH_4} M_{CH_4} = (20 \text{ kmol})(16 \text{ kg/kmol}) = 320 \text{ kg}$$

The total mass is

$$\begin{aligned} m_m &= m_{O_2} + m_{N_2} + m_{CO_2} + m_{CH_4} \\ &= 960 + 1120 + 440 + 320 \\ &= 2840 \text{ kg} \end{aligned}$$



Then the mass fractions are

$$mf_{O_2} = \frac{m_{O_2}}{m_m} = \frac{960 \text{ kg}}{2840 \text{ kg}} = 0.3380$$

$$mf_{N_2} = \frac{m_{N_2}}{m_m} = \frac{1120 \text{ kg}}{2840 \text{ kg}} = 0.3944$$

$$mf_{CO_2} = \frac{m_{CO_2}}{m_m} = \frac{440 \text{ kg}}{2840 \text{ kg}} = 0.1549$$

$$mf_{CH_4} = \frac{m_{CH_4}}{m_m} = \frac{320 \text{ kg}}{2840 \text{ kg}} = 0.1127$$

The constant-pressure specific heat of the mixture is determined from

$$\begin{aligned} c_p &= mf_{O_2} c_{p,O_2} + mf_{N_2} c_{p,N_2} + mf_{CO_2} c_{p,CO_2} + mf_{CH_4} c_{p,CH_4} \\ &= 0.3380 \times 0.918 + 0.3944 \times 1.039 + 0.1549 \times 0.846 + 0.1127 \times 2.2537 \\ &= 1.1051 \text{ kJ/kg} \cdot \text{K} \end{aligned}$$

An energy balance on the tube gives

$$q_{in} = c_p (T_2 - T_1) = (1.1051 \text{ kJ/kg} \cdot \text{K})(200 - 20) \text{ K} = \mathbf{199 \text{ kJ/kg}}$$

13-54E A mixture of helium and nitrogen is heated at constant pressure in a closed system. The work produced is to be determined.

Assumptions **1** Helium and nitrogen are ideal gases. **2** The process is reversible.

Properties The mole numbers of helium and nitrogen are 4.0 and 28.0 lbm/lbmol, respectively (Table A-1E).

Analysis One lbm of this mixture consists of 0.35 lbm of nitrogen and 0.65 lbm of helium or $0.35 \text{ lbm}/(28.0 \text{ lbm/lbmol}) = 0.0125 \text{ lbmol}$ of nitrogen and $0.65 \text{ lbm}/(4.0 \text{ lbm/lbmol}) = 0.1625 \text{ lbmol}$ of helium. The total mole is $0.0125 + 0.1625 = 0.175 \text{ lbmol}$. The constituent mole fraction are then

$$y_{\text{N}_2} = \frac{N_{\text{N}_2}}{N_{\text{total}}} = \frac{0.0125 \text{ lbmol}}{0.175 \text{ lbmol}} = 0.07143$$

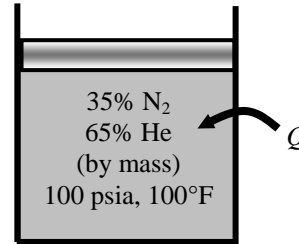
$$y_{\text{He}} = \frac{N_{\text{He}}}{N_{\text{total}}} = \frac{0.1625 \text{ lbmol}}{0.175 \text{ lbmol}} = 0.9286$$

The effective molecular weight of this mixture is

$$\begin{aligned} M &= y_{\text{N}_2} M_{\text{N}_2} + y_{\text{He}} M_{\text{He}} \\ &= (0.07143)(28) + (0.9286)(4) \\ &= 5.714 \text{ lbm/lbmol} \end{aligned}$$

The work done is determined from

$$\begin{aligned} w &= \int_1^2 P d\nu = P_2 \nu_2 - P_1 \nu_1 = R(T_2 - T_1) \\ &= \frac{R_u}{M} (T_2 - T_1) = \frac{1.9858 \text{ Btu/lbmol} \cdot \text{R}}{5.714 \text{ lbm/lbmol}} (500 - 100) \text{R} = \mathbf{139.0 \text{ Btu/lbm}} \end{aligned}$$



13-55 The volume fractions of components of a gas mixture are given. This mixture is expanded isentropically to a specified pressure. The work produced per unit mass of the mixture is to be determined.

Assumptions All gases will be modeled as ideal gases with constant specific heats.

Properties The molar masses of H_2 , He, and N_2 are 2.0, 4.0, and 28.0 kg/kmol, respectively (Table A-1). The constant-pressure specific heats of these gases at room temperature are 14.307, 5.1926, and 1.039 kJ/kg·K, respectively (Table A-2a).

Analysis We consider 100 kmol of this mixture. Noting that volume fractions are equal to the mole fractions, mass of each component are

$$\begin{aligned} m_{H_2} &= N_{H_2} M_{H_2} = (30 \text{ kmol})(2 \text{ kg/kmol}) = 60 \text{ kg} \\ m_{He} &= N_{He} M_{He} = (40 \text{ kmol})(4 \text{ kg/kmol}) = 160 \text{ kg} \\ m_{N_2} &= N_{N_2} M_{N_2} = (30 \text{ kmol})(28 \text{ kg/kmol}) = 840 \text{ kg} \end{aligned}$$

The total mass is

$$m_m = m_{H_2} + m_{He} + m_{N_2} = 60 + 160 + 840 = 1060 \text{ kg}$$

Then the mass fractions are

$$\begin{aligned} \text{mf}_{H_2} &= \frac{m_{H_2}}{m_m} = \frac{60 \text{ kg}}{1060 \text{ kg}} = 0.05660 \\ \text{mf}_{He} &= \frac{m_{He}}{m_m} = \frac{160 \text{ kg}}{1060 \text{ kg}} = 0.1509 \\ \text{mf}_{N_2} &= \frac{m_{N_2}}{m_m} = \frac{840 \text{ kg}}{1060 \text{ kg}} = 0.7925 \end{aligned}$$

<p>30% H_2 40% He 30% N_2 (by volume) 5 MPa, 600°C</p>
--

The apparent molecular weight of the mixture is

$$M_m = \frac{m_m}{N_m} = \frac{1060 \text{ kg}}{100 \text{ kmol}} = 10.60 \text{ kg/kmol}$$

The constant-pressure specific heat of the mixture is determined from

$$\begin{aligned} c_p &= \text{mf}_{H_2} c_{p,H_2} + \text{mf}_{He} c_{p,He} + \text{mf}_{N_2} c_{p,N_2} \\ &= 0.05660 \times 14.307 + 0.1509 \times 5.1926 + 0.7925 \times 1.039 \\ &= 2.417 \text{ kJ/kg} \cdot \text{K} \end{aligned}$$

The apparent gas constant of the mixture is

$$R = \frac{R_u}{M_m} = \frac{8.314 \text{ kJ/kmol} \cdot \text{K}}{10.60 \text{ kg/kmol}} = 0.7843 \text{ kJ/kg} \cdot \text{K}$$

Then the constant-volume specific heat is

$$c_v = c_p - R = 2.417 - 0.7843 = 1.633 \text{ kJ/kg} \cdot \text{K}$$

The specific heat ratio is

$$k = \frac{c_p}{c_v} = \frac{2.417}{1.633} = 1.480$$

The temperature at the end of the expansion is

$$T_2 = T_1 \left(\frac{P_2}{P_1} \right)^{(k-1)/k} = (873 \text{ K}) \left(\frac{200 \text{ kPa}}{5000 \text{ kPa}} \right)^{0.48/1.48} = 307 \text{ K}$$

An energy balance on the adiabatic expansion process gives

$$w_{\text{out}} = c_p (T_1 - T_2) = (2.417 \text{ kJ/kg} \cdot \text{K})(873 - 307) \text{ K} = \mathbf{1368 \text{ kJ/kg}}$$

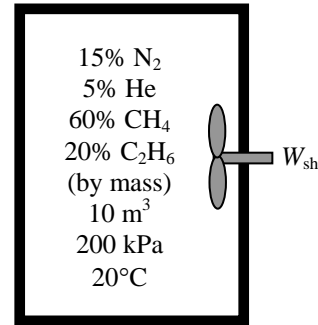
13-56 The mass fractions of components of a gas mixture are given. This mixture is enclosed in a rigid, well-insulated vessel, and a paddle wheel in the vessel is turned until specified amount of work have been done on the mixture. The mixture's final pressure and temperature are to be determined.

Assumptions All gases will be modeled as ideal gases with constant specific heats.

Properties The molar masses of N_2 , He, CH_4 , and C_2H_6 are 28.0, 4.0, 16.0, and 30.0 kg/kmol, respectively (Table A-1). The constant-pressure specific heats of these gases at room temperature are 1.039, 5.1926, 2.2537, and 1.7662 kJ/kg·K, respectively (Table A-2a).

Analysis We consider 100 kg of this mixture. The mole numbers of each component are

$$\begin{aligned} N_{N_2} &= \frac{m_{N_2}}{M_{N_2}} = \frac{15 \text{ kg}}{28 \text{ kg/kmol}} = 0.5357 \text{ kmol} \\ N_{He} &= \frac{m_{He}}{M_{He}} = \frac{5 \text{ kg}}{4 \text{ kg/kmol}} = 1.25 \text{ kmol} \\ N_{CH_4} &= \frac{m_{CH_4}}{M_{CH_4}} = \frac{60 \text{ kg}}{16 \text{ kg/kmol}} = 3.75 \text{ kmol} \\ N_{C_2H_6} &= \frac{m_{C_2H_6}}{M_{C_2H_6}} = \frac{20 \text{ kg}}{30 \text{ kg/kmol}} = 0.6667 \text{ kmol} \end{aligned}$$



The mole number of the mixture is

$$N_m = N_{N_2} + N_{He} + N_{CH_4} + N_{C_2H_6} = 0.5357 + 1.25 + 3.75 + 0.6667 = 6.2024 \text{ kmol}$$

The apparent molecular weight of the mixture is

$$M_m = \frac{m_m}{N_m} = \frac{100 \text{ kg}}{6.2024 \text{ kmol}} = 16.12 \text{ kg/kmol}$$

The constant-pressure specific heat of the mixture is determined from

$$\begin{aligned} c_p &= mf_{N_2} c_{p,N_2} + mf_{He} c_{p,He} + mf_{CH_4} c_{p,CH_4} + mf_{C_2H_6} c_{p,C_2H_6} \\ &= 0.15 \times 1.039 + 0.05 \times 5.1926 + 0.60 \times 2.2537 + 0.20 \times 1.7662 \\ &= 2.121 \text{ kJ/kg} \cdot \text{K} \end{aligned}$$

The apparent gas constant of the mixture is

$$R = \frac{R_u}{M_m} = \frac{8.134 \text{ kJ/kmol} \cdot \text{K}}{16.12 \text{ kg/kmol}} = 0.5158 \text{ kJ/kg} \cdot \text{K}$$

Then the constant-volume specific heat is

$$c_v = c_p - R = 2.121 - 0.5158 = 1.605 \text{ kJ/kg} \cdot \text{K}$$

The mass in the container is

$$m_m = \frac{P_1 V_m}{RT_1} = \frac{(200 \text{ kPa})(10 \text{ m}^3)}{(0.5158 \text{ kPa} \cdot \text{m}^3/\text{kg} \cdot \text{K})(293 \text{ K})} = 13.23 \text{ kg}$$

An energy balance on the system gives

$$W_{sh,in} = m_m c_v (T_2 - T_1) \longrightarrow T_2 = T_1 + \frac{W_{sh,in}}{m_m c_v} = (293 \text{ K}) + \frac{100 \text{ kJ}}{(13.23 \text{ kg})(1.605 \text{ kJ/kg} \cdot \text{K})} = \mathbf{297.7 \text{ K}}$$

Since the volume remains constant and this is an ideal gas,

$$P_2 = P_1 \frac{T_2}{T_1} = (200 \text{ kPa}) \frac{297.7 \text{ K}}{293 \text{ K}} = \mathbf{203.2 \text{ kPa}}$$

13-57 Propane and air mixture is compressed isentropically in an internal combustion engine. The work input is to be determined.

Assumptions Under specified conditions propane and air can be treated as ideal gases, and the mixture as an ideal gas mixture.

Properties The molar masses of C_3H_8 and air are 44.0 and 28.97 kg/kmol, respectively (TableA-1).

Analysis Given the air-fuel ratio, the mass fractions are determined to be

$$mf_{\text{air}} = \frac{AF}{AF+1} = \frac{16}{17} = 0.9412$$

$$mf_{C_3H_8} = \frac{1}{AF+1} = \frac{1}{17} = 0.05882$$

The molar mass of the mixture is determined to be

$$M_m = \frac{1}{\frac{mf_{\text{air}}}{M_{\text{air}}} + \frac{mf_{C_3H_8}}{M_{C_3H_8}}} = \frac{1}{\frac{0.9412}{28.97 \text{ kg/kmol}} + \frac{0.05882}{44.0 \text{ kg/kmol}}} = 29.56 \text{ kg/kmol}$$

The mole fractions are

$$y_{\text{air}} = mf_{\text{air}} \frac{M_m}{M_{\text{air}}} = (0.9412) \frac{29.56 \text{ kg/kmol}}{28.97 \text{ kg/kmol}} = 0.9606$$

$$y_{C_3H_8} = mf_{C_3H_8} \frac{M_m}{M_{C_3H_8}} = (0.05882) \frac{29.56 \text{ kg/kmol}}{44.0 \text{ kg/kmol}} = 0.03944$$

The final pressure is expressed from ideal gas relation to be

$$P_2 = P_1 r \frac{T_2}{T_1} = (95 \text{ kPa})(9.5) \frac{T_2}{(30 + 273.15) \text{ K}} = 2.977T_2 \quad (1)$$

since the final temperature is not known. Using Dalton's law to find partial pressures, the entropies at the initial state are determined from EES to be:

$$T = 30^\circ\text{C}, P = (0.9606 \times 95) = 91.26 \text{ kPa} \longrightarrow s_{\text{air},1} = 5.7417 \text{ kJ/kg}\cdot\text{K}$$

$$T = 30^\circ\text{C}, P = (0.03944 \times 95) = 3.75 \text{ kPa} \longrightarrow s_{C_3H_8,1} = 6.7697 \text{ kJ/kg}\cdot\text{K}$$

The final state entropies cannot be determined at this point since the final pressure and temperature are not known. However, for an isentropic process, the entropy change is zero and the final temperature and the final pressure may be determined from

$$\Delta s_{\text{total}} = mf_{\text{air}} \Delta s_{\text{air}} + mf_{C_3H_8} \Delta s_{C_3H_8} = 0$$

and using Eq. (1). The solution may be obtained using EES to be

$$T_2 = 654.9 \text{ K}, P_2 = 1951 \text{ kPa}$$

The initial and final internal energies are (from EES)

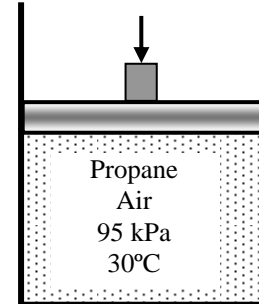
$$T_1 = 30^\circ\text{C} \longrightarrow \begin{array}{l} u_{\text{air},1} = 216.5 \text{ kJ/kg} \\ u_{C_3H_8,1} = -2404 \text{ kJ/kg} \end{array} \quad T_2 = 654.9 \text{ K} \longrightarrow \begin{array}{l} u_{\text{air},2} = 477.1 \text{ kJ/kg} \\ u_{C_3H_8,2} = -1607 \text{ kJ/kg} \end{array}$$

Noting that the heat transfer is zero, an energy balance on the system gives

$$q_{\text{in}} + w_{\text{in}} = \Delta u_m \longrightarrow w_{\text{in}} = \Delta u_m$$

$$\text{where } \Delta u_m = mf_{\text{air}} (u_{\text{air},2} - u_{\text{air},1}) + mf_{C_3H_8} (u_{C_3H_8,2} - u_{C_3H_8,1})$$

$$\text{Substituting, } w_{\text{in}} = \Delta u_m = (0.9412)(477.1 - 216.5) + (0.05882)[(-1607) - (-2404)] = \mathbf{292.2 \text{ kJ/kg}}$$



13-58 The moles, temperatures, and pressures of two gases forming a mixture are given. The mixture temperature and pressure are to be determined.

Assumptions 1 Under specified conditions both CO₂ and H₂ can be treated as ideal gases, and the mixture as an ideal gas mixture. **2** The tank is insulated and thus there is no heat transfer. **3** There are no other forms of work involved.

Properties The molar masses and specific heats of CO₂ and H₂ are 44.0 kg/kmol, 2.0 kg/kmol, 0.657 kJ/kg·°C, and 10.183 kJ/kg·°C, respectively. (Tables A-1 and A-2b).

Analysis (a) We take both gases as our system. No heat, work, or mass crosses the system boundary, therefore this is a closed system with $Q = 0$ and $W = 0$. Then the energy balance for this closed system reduces to

$$\begin{aligned} E_{\text{in}} - E_{\text{out}} &= \Delta E_{\text{system}} \\ 0 &= \Delta U = \Delta U_{\text{CO}_2} + \Delta U_{\text{H}_2} \\ 0 &= [mc_v(T_m - T_1)]_{\text{CO}_2} + [mc_v(T_m - T_1)]_{\text{H}_2} \end{aligned}$$

Using c_v values at room temperature and noting that $m = NM$, the final temperature of the mixture is determined to be

$$\begin{aligned} (2.5 \times 44 \text{ kg})(0.657 \text{ kJ/kg} \cdot ^\circ\text{C})(T_m - 27^\circ\text{C}) + (7.5 \times 2 \text{ kg})(10.183 \text{ kJ/kg} \cdot ^\circ\text{C})(T_m - 40^\circ\text{C}) &= 0 \\ T_m &= \mathbf{35.8^\circ\text{C}} \quad (308.8 \text{ K}) \end{aligned}$$

CO ₂	H ₂
2.5 kmol	7.5 kmol
200 kPa	400 kPa
27°C	40°C

(b) The volume of each tank is determined from

$$\begin{aligned} \mathcal{V}_{\text{CO}_2} &= \left(\frac{NR_u T_1}{P_1} \right)_{\text{CO}_2} = \frac{(2.5 \text{ kmol})(8.314 \text{ kPa} \cdot \text{m}^3/\text{kmol} \cdot \text{K})(300 \text{ K})}{200 \text{ kPa}} = 31.18 \text{ m}^3 \\ \mathcal{V}_{\text{H}_2} &= \left(\frac{NR_u T_1}{P_1} \right)_{\text{H}_2} = \frac{(7.5 \text{ kmol})(8.314 \text{ kPa} \cdot \text{m}^3/\text{kmol} \cdot \text{K})(313 \text{ K})}{400 \text{ kPa}} = 48.79 \text{ m}^3 \end{aligned}$$

Thus,

$$\begin{aligned} \mathcal{V}_m &= \mathcal{V}_{\text{CO}_2} + \mathcal{V}_{\text{H}_2} = 31.18 \text{ m}^3 + 48.79 \text{ m}^3 = 79.97 \text{ m}^3 \\ N_m &= N_{\text{CO}_2} + N_{\text{H}_2} = 2.5 \text{ kmol} + 7.5 \text{ kmol} = 10.0 \text{ kmol} \end{aligned}$$

and

$$P_m = \frac{N_m R_u T_m}{V_m} = \frac{(10.0 \text{ kmol})(8.314 \text{ kPa} \cdot \text{m}^3/\text{kmol} \cdot \text{K})(308.8 \text{ K})}{79.97 \text{ m}^3} = \mathbf{321 \text{ kPa}}$$

13-59 The mass fractions of components of a gas mixture are given. This mixture is compressed in a reversible, isothermal, steady-flow compressor. The work and heat transfer for this compression per unit mass of the mixture are to be determined.

Assumptions All gases will be modeled as ideal gases with constant specific heats.

Properties The molar masses of CH_4 , C_3H_8 , and C_4H_{10} are 16.0, 44.0, and 58.0 kg/kmol, respectively (Table A-1).

Analysis The mole numbers of each component are

$$N_{\text{CH}_4} = \frac{m_{\text{CH}_4}}{M_{\text{CH}_4}} = \frac{60 \text{ kg}}{16 \text{ kg/kmol}} = 3.75 \text{ kmol}$$

$$N_{\text{C}_3\text{H}_8} = \frac{m_{\text{C}_3\text{H}_8}}{M_{\text{C}_3\text{H}_8}} = \frac{25 \text{ kg}}{44 \text{ kg/kmol}} = 0.5682 \text{ kmol}$$

$$N_{\text{C}_4\text{H}_{10}} = \frac{m_{\text{C}_4\text{H}_{10}}}{M_{\text{C}_4\text{H}_{10}}} = \frac{15 \text{ kg}}{58 \text{ kg/kmol}} = 0.2586 \text{ kmol}$$

The mole number of the mixture is

$$N_m = N_{\text{CH}_4} + N_{\text{C}_3\text{H}_8} + N_{\text{C}_4\text{H}_{10}}$$

$$= 3.75 + 0.5682 + 0.2586 = 4.5768 \text{ kmol}$$

The apparent molecular weight of the mixture is

$$M_m = \frac{m_m}{N_m} = \frac{100 \text{ kg}}{4.5768 \text{ kmol}} = 21.85 \text{ kg/kmol}$$

The apparent gas constant of the mixture is

$$R = \frac{R_u}{M_m} = \frac{8.314 \text{ kJ/kmol} \cdot \text{K}}{21.85 \text{ kg/kmol}} = 0.3805 \text{ kJ/kg} \cdot \text{K}$$

For a reversible, isothermal process, the work input is

$$w_{\text{in}} = RT \ln\left(\frac{P_2}{P_1}\right) = (0.3805 \text{ kJ/kg} \cdot \text{K})(293 \text{ K}) \ln\left(\frac{1000 \text{ kPa}}{100 \text{ kPa}}\right) = \mathbf{257 \text{ kJ/kg}}$$

An energy balance on the control volume gives

$$\underbrace{\dot{E}_{\text{in}} - \dot{E}_{\text{out}}}_{\text{Rate of net energy transfer by heat, work, and mass}} = \underbrace{\Delta \dot{E}_{\text{system}}}_{\text{Rate of change in internal, kinetic, potential, etc. energies}} \stackrel{\text{no (steady)}}{=} 0$$

$$\dot{E}_{\text{in}} = \dot{E}_{\text{out}}$$

$$\dot{m}h_1 + \dot{W}_{\text{in}} = \dot{m}h_2 + \dot{Q}_{\text{out}}$$

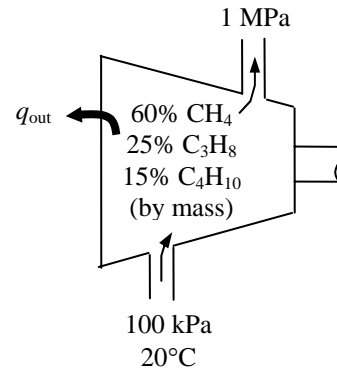
$$\dot{W}_{\text{in}} - \dot{Q}_{\text{out}} = \dot{m}(h_2 - h_1)$$

$$w_{\text{in}} - q_{\text{out}} = c_p(T_2 - T_1) = 0 \quad \text{since } T_2 = T_1$$

$$w_{\text{in}} = q_{\text{out}}$$

That is,

$$q_{\text{out}} = w_{\text{in}} = \mathbf{257 \text{ kJ/kg}}$$



13-60 The volume fractions of components of a gas mixture during the expansion process of the ideal Otto cycle are given. The thermal efficiency of this cycle is to be determined.

Assumptions All gases will be modeled as ideal gases with constant specific heats.

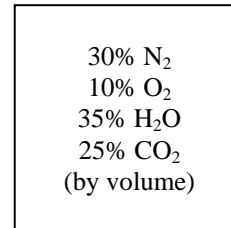
Properties The molar masses of N_2 , O_2 , H_2O , and CO_2 are 28.0, 32.0, 18.0, and 44.0 kg/kmol, respectively (Table A-1). The constant-pressure specific heats of these gases at room temperature are 1.039, 0.918, 1.8723, and 0.846 kJ/kg·K, respectively. The air properties at room temperature are $c_p = 1.005$ kJ/kg·K, $c_v = 0.718$ kJ/kg·K, $k = 1.4$ (Table A-2a).

Analysis We consider 100 kmol of this mixture. Noting that volume fractions are equal to the mole fractions, mass of each component are

$$\begin{aligned} m_{N_2} &= N_{N_2} M_{N_2} = (30 \text{ kmol})(28 \text{ kg/kmol}) = 840 \text{ kg} \\ m_{O_2} &= N_{O_2} M_{O_2} = (10 \text{ kmol})(32 \text{ kg/kmol}) = 320 \text{ kg} \\ m_{H_2O} &= N_{H_2O} M_{H_2O} = (35 \text{ kmol})(18 \text{ kg/kmol}) = 630 \text{ kg} \\ m_{CO_2} &= N_{CO_2} M_{CO_2} = (25 \text{ kmol})(44 \text{ kg/kmol}) = 1100 \text{ kg} \end{aligned}$$

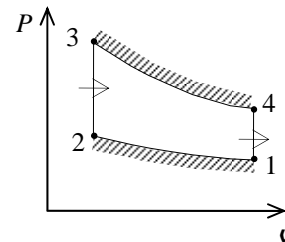
The total mass is

$$\begin{aligned} m_m &= m_{N_2} + m_{O_2} + m_{H_2O} + m_{CO_2} \\ &= 840 + 320 + 630 + 1100 \\ &= 2890 \text{ kg} \end{aligned}$$



Then the mass fractions are

$$\begin{aligned} mf_{N_2} &= \frac{m_{N_2}}{m_m} = \frac{840 \text{ kg}}{2890 \text{ kg}} = 0.2907 \\ mf_{O_2} &= \frac{m_{O_2}}{m_m} = \frac{320 \text{ kg}}{2890 \text{ kg}} = 0.1107 \\ mf_{H_2O} &= \frac{m_{H_2O}}{m_m} = \frac{630 \text{ kg}}{2890 \text{ kg}} = 0.2180 \\ mf_{CO_2} &= \frac{m_{CO_2}}{m_m} = \frac{1100 \text{ kg}}{2890 \text{ kg}} = 0.3806 \end{aligned}$$



The constant-pressure specific heat of the mixture is determined from

$$\begin{aligned} c_p &= mf_{N_2} c_{p,N_2} + mf_{O_2} c_{p,O_2} + mf_{H_2O} c_{p,H_2O} + mf_{CO_2} c_{p,CO_2} \\ &= 0.2907 \times 1.039 + 0.1107 \times 0.918 + 0.2180 \times 1.8723 + 0.3806 \times 0.846 \\ &= 1.134 \text{ kJ/kg} \cdot \text{K} \end{aligned}$$

The apparent molecular weight of the mixture is

$$M_m = \frac{m_m}{N_m} = \frac{2890 \text{ kg}}{100 \text{ kmol}} = 28.90 \text{ kg/kmol}$$

The apparent gas constant of the mixture is

$$R = \frac{R_u}{M_m} = \frac{8.314 \text{ kJ/kmol} \cdot \text{K}}{28.90 \text{ kg/kmol}} = 0.2877 \text{ kJ/kg} \cdot \text{K}$$

Then the constant-volume specific heat is

$$c_v = c_p - R = 1.134 - 0.2877 = 0.846 \text{ kJ/kg} \cdot \text{K}$$

The specific heat ratio is

$$k = \frac{c_p}{c_v} = \frac{1.134}{0.846} = 1.340$$

The average of the air properties at room temperature and combustion gas properties are

$$c_{p,\text{avg}} = 0.5(1.134 + 1.005) = 1.070 \text{ kJ/kg} \cdot \text{K}$$

$$c_{v,\text{avg}} = 0.5(0.846 + 0.718) = 0.782 \text{ kJ/kg} \cdot \text{K}$$

$$k_{\text{avg}} = 0.5(1.34 + 1.4) = 1.37$$

These average properties will be used for heat addition and rejection processes. For compression, the air properties at room temperature and during expansion, the mixture properties will be used. During the compression process,

$$T_2 = T_1 r^{k-1} = (288 \text{ K})(8)^{0.4} = 662 \text{ K}$$

During the heat addition process,

$$q_{\text{in}} = c_{v,\text{avg}}(T_3 - T_2) = (0.782 \text{ kJ/kg} \cdot \text{K})(1373 - 662) \text{ K} = 556 \text{ kJ/kg}$$

During the expansion process,

$$T_4 = T_3 \left(\frac{1}{r} \right)^{k-1} = (1373 \text{ K}) \left(\frac{1}{8} \right)^{0.37} = 636.1 \text{ K}$$

During the heat rejection process,

$$q_{\text{out}} = c_{v,\text{avg}}(T_4 - T_1) = (0.782 \text{ kJ/kg} \cdot \text{K})(636.1 - 288) \text{ K} = 272.2 \text{ kJ/kg}$$

The thermal efficiency of the cycle is then

$$\eta_{\text{th}} = 1 - \frac{q_{\text{out}}}{q_{\text{in}}} = 1 - \frac{272.2 \text{ kJ/kg}}{556 \text{ kJ/kg}} = \mathbf{0.511}$$

13-61 The thermal efficiency of the cycle in the previous problem is to be compared to that predicted by air standard analysis.

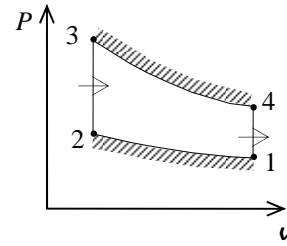
Assumptions Air-standard assumptions are applicable.

Properties The air properties at room temperature are $c_p = 1.005 \text{ kJ/kg}\cdot\text{K}$, $c_v = 0.718 \text{ kJ/kg}\cdot\text{K}$, $k = 1.4$ (Table A-2a).

Analysis In the previous problem, the thermal efficiency of the cycle was determined to be **0.511** (51.1%). The thermal efficiency with air-standard model is determined from

$$\eta_{\text{th}} = 1 - \frac{1}{r^{k-1}} = 1 - \frac{1}{8^{0.4}} = \mathbf{0.565}$$

which is greater than that calculated with gas mixture analysis in the previous problem.



13-62E The volume fractions of components of a gas mixture passing through the turbine of a simple ideal Brayton cycle are given. The thermal efficiency of this cycle is to be determined.

Assumptions All gases will be modeled as ideal gases with constant specific heats.

Properties The molar masses of N_2 , O_2 , H_2O , and CO_2 are 28.0, 32.0, 18.0, and 44.0 lbm/lbmol, respectively (Table A-1E). The constant-pressure specific heats of these gases at room temperature are 0.248, 0.219, 0.445, and 0.203 Btu/lbm·R, respectively. The air properties at room temperature are $c_p = 0.240$ Btu/lbm·R, $c_v = 0.171$ Btu/lbm·R, $k = 1.4$ (Table A-2Ea).

Analysis We consider 100 lbmol of this mixture. Noting that volume fractions are equal to the mole fractions, mass of each component are

$$\begin{aligned} m_{N_2} &= N_{N_2} M_{N_2} = (20 \text{ lbmol})(28 \text{ lbm/lbmol}) = 560 \text{ lbm} \\ m_{O_2} &= N_{O_2} M_{O_2} = (5 \text{ lbmol})(32 \text{ lbm/lbmol}) = 160 \text{ lbm} \\ m_{H_2O} &= N_{H_2O} M_{H_2O} = (35 \text{ lbmol})(18 \text{ lbm/lbmol}) = 630 \text{ lbm} \\ m_{CO_2} &= N_{CO_2} M_{CO_2} = (40 \text{ lbmol})(44 \text{ lbm/lbmol}) = 1760 \text{ lbm} \end{aligned}$$

The total mass is

$$\begin{aligned} m_m &= m_{N_2} + m_{O_2} + m_{H_2O} + m_{CO_2} \\ &= 560 + 160 + 630 + 1760 \\ &= 3110 \text{ lbm} \end{aligned}$$

Then the mass fractions are

$$\begin{aligned} mf_{N_2} &= \frac{m_{N_2}}{m_m} = \frac{560 \text{ lbm}}{3110 \text{ lbm}} = 0.1801 \\ mf_{O_2} &= \frac{m_{O_2}}{m_m} = \frac{160 \text{ lbm}}{3110 \text{ lbm}} = 0.05145 \\ mf_{H_2O} &= \frac{m_{H_2O}}{m_m} = \frac{630 \text{ lbm}}{3110 \text{ lbm}} = 0.2026 \\ mf_{CO_2} &= \frac{m_{CO_2}}{m_m} = \frac{1760 \text{ lbm}}{3110 \text{ lbm}} = 0.5659 \end{aligned}$$

The constant-pressure specific heat of the mixture is determined from

$$\begin{aligned} c_p &= mf_{N_2} c_{p,N_2} + mf_{O_2} c_{p,O_2} + mf_{H_2O} c_{p,H_2O} + mf_{CO_2} c_{p,CO_2} \\ &= 0.1801 \times 0.248 + 0.05145 \times 0.219 + 0.2026 \times 0.445 + 0.5659 \times 0.203 \\ &= 0.2610 \text{ Btu/lbm} \cdot \text{R} \end{aligned}$$

The apparent molecular weight of the mixture is

$$M_m = \frac{m_m}{N_m} = \frac{3110 \text{ lbm}}{100 \text{ lbmol}} = 31.10 \text{ lbm/lbmol}$$

The apparent gas constant of the mixture is

$$R = \frac{R_u}{M_m} = \frac{1.9858 \text{ Btu/lbmol} \cdot \text{R}}{31.10 \text{ lbm/lbmol}} = 0.06385 \text{ Btu/lbm} \cdot \text{R}$$

Then the constant-volume specific heat is

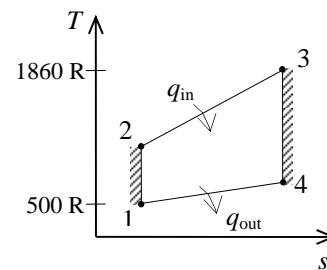
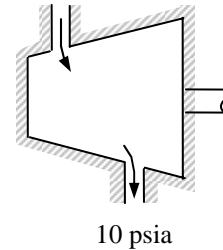
$$c_v = c_p - R = 0.2610 - 0.06385 = 0.1971 \text{ Btu/lbm} \cdot \text{R}$$

The specific heat ratio is

$$k = \frac{c_p}{c_v} = \frac{0.2610}{0.1971} = 1.324$$

The average of the air properties at room temperature and combustion gas properties are

20% N_2 , 5% O_2
35% H_2O , 40% CO_2
(by volume)



$$c_{p,\text{avg}} = 0.5(0.2610 + 0.240) = 0.2505 \text{ Btu/lbm} \cdot \text{R}$$

$$c_{v,\text{avg}} = 0.5(0.1971 + 0.171) = 0.1841 \text{ Btu/lbm} \cdot \text{R}$$

$$k_{\text{avg}} = 0.5(1.324 + 1.4) = 1.362$$

These average properties will be used for heat addition and rejection processes. For compression, the air properties at room temperature and during expansion, the mixture properties will be used. During the compression process,

$$T_2 = T_1 \left(\frac{P_2}{P_1} \right)^{(k-1)/k} = (500 \text{ R})(6)^{0.4/1.4} = 834.3 \text{ R}$$

During the heat addition process,

$$q_{\text{in}} = c_{p,\text{avg}}(T_3 - T_2) = (0.2505 \text{ Btu/lbm} \cdot \text{R})(1860 - 834.3) \text{ R} = 256.9 \text{ Btu/lbm}$$

During the expansion process,

$$T_4 = T_3 \left(\frac{P_4}{P_3} \right)^{(k-1)/k} = (1860 \text{ R}) \left(\frac{1}{6} \right)^{0.362/1.362} = 1155.3 \text{ R}$$

During the heat rejection process,

$$q_{\text{out}} = c_{p,\text{avg}}(T_4 - T_1) = (0.2505 \text{ Btu/lbm} \cdot \text{R})(1155.3 - 500) \text{ R} = 164.2 \text{ Btu/lbm}$$

The thermal efficiency of the cycle is then

$$\eta_{\text{th}} = 1 - \frac{q_{\text{out}}}{q_{\text{in}}} = 1 - \frac{164.2 \text{ Btu/lbm}}{256.9 \text{ Btu/lbm}} = 0.361 = \mathbf{36.1\%}$$

13-63E The thermal efficiency of the cycle in the previous problem is to be compared to that predicted by air standard analysis?

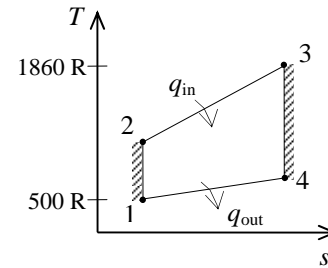
Assumptions Air-standard assumptions are applicable.

Properties The air properties at room temperature are $c_p = 0.240$ Btu/lbm-R, $c_v = 0.171$ Btu/lbm-R, $k = 1.4$ (Table A-2Ea).

Analysis In the previous problem, the thermal efficiency of the cycle was determined to be **0.361** (36.1%). The thermal efficiency with air-standard model is determined from

$$\eta_{\text{th}} = 1 - \frac{1}{r_p^{(k-1)/k}} = 1 - \frac{1}{6^{0.4/1.4}} = 0.401 = \mathbf{40.1\%}$$

which is greater than that calculated with gas mixture analysis in the previous problem.



13-64E The mass fractions of a natural gas mixture at a specified pressure and temperature trapped in a geological location are given. This natural gas is pumped to the surface. The work required is to be determined using Kay's rule and the enthalpy-departure method.

Properties The molar masses of CH_4 and C_2H_6 are 16.0 and 30.0 lbm/lbmol, respectively. The critical properties are 343.9 R, 673 psia for CH_4 and 549.8 R and 708 psia for C_2H_6 (Table A-1E). The constant-pressure specific heats of these gases at room temperature are 0.532 and 0.427 Btu/lbm·R, respectively (Table A-2Ea).

Analysis We consider 100 lbm of this mixture. Then the mole numbers of each component are

$$N_{\text{CH}_4} = \frac{m_{\text{CH}_4}}{M_{\text{CH}_4}} = \frac{75 \text{ lbm}}{16 \text{ lbm/lbmol}} = 4.6875 \text{ lbmol}$$

$$N_{\text{C}_2\text{H}_6} = \frac{m_{\text{C}_2\text{H}_6}}{M_{\text{C}_2\text{H}_6}} = \frac{25 \text{ lbm}}{30 \text{ lbm/lbmol}} = 0.8333 \text{ lbmol}$$

The mole number of the mixture and the mole fractions are

$$N_m = 4.6875 + 0.8333 = 5.5208 \text{ lbmol}$$

$$y_{\text{CH}_4} = \frac{N_{\text{CH}_4}}{N_m} = \frac{4.6875 \text{ lbmol}}{5.5208 \text{ lbmol}} = 0.8491$$

$$y_{\text{C}_2\text{H}_6} = \frac{N_{\text{C}_2\text{H}_6}}{N_m} = \frac{0.8333 \text{ lbmol}}{5.5208 \text{ lbmol}} = 0.1509$$

75% CH_4
25% C_2H_6
(by mass)
2000 psia
300°F

Then the apparent molecular weight of the mixture becomes

$$M_m = \frac{m_m}{N_m} = \frac{100 \text{ lbm}}{5.5208 \text{ lbmol}} = 18.11 \text{ lbm/lbmol}$$

The apparent gas constant of the mixture is

$$R = \frac{R_u}{M_m} = \frac{1.9858 \text{ Btu/lbmol} \cdot \text{R}}{18.11 \text{ lbm/lbmol}} = 0.1097 \text{ Btu/lbm} \cdot \text{R}$$

The constant-pressure specific heat of the mixture is determined from

$$c_p = \text{mf}_{\text{CH}_4} c_{p,\text{CH}_4} + \text{mf}_{\text{C}_2\text{H}_6} c_{p,\text{C}_2\text{H}_6} = 0.75 \times 0.532 + 0.25 \times 0.427 = 0.506 \text{ Btu/lbm} \cdot \text{R}$$

To use Kay's rule, we need to determine the pseudo-critical temperature and pseudo-critical pressure of the mixture using the critical point properties of gases.

$$\begin{aligned} T'_{\text{cr},m} &= \sum y_i T_{\text{cr},i} = y_{\text{CH}_4} T_{\text{cr},\text{CH}_4} + y_{\text{C}_2\text{H}_6} T_{\text{cr},\text{C}_2\text{H}_6} \\ &= (0.8491)(343.9 \text{ R}) + (0.1509)(549.8 \text{ R}) = 375.0 \text{ R} \end{aligned}$$

$$\begin{aligned} P'_{\text{cr},m} &= \sum y_i P_{\text{cr},i} = y_{\text{CH}_4} P_{\text{cr},\text{CH}_4} + y_{\text{C}_2\text{H}_6} P_{\text{cr},\text{C}_2\text{H}_6} \\ &= (0.8491)(673 \text{ psia}) + (0.1509)(708 \text{ psia}) = 678.3 \text{ psia} \end{aligned}$$

The compressibility factor of the gas mixture in the reservoir and the mass of this gas are

$$\left. \begin{aligned} T_R &= \frac{T_m}{T'_{\text{cr},m}} = \frac{760 \text{ R}}{375.0 \text{ R}} = 2.027 \\ P_R &= \frac{P_m}{P'_{\text{cr},m}} = \frac{2000 \text{ psia}}{678.3 \text{ psia}} = 2.949 \end{aligned} \right\} Z_m = 0.963 \quad (\text{Fig. A-15})$$

$$m = \frac{P\mathcal{V}}{Z_m RT} = \frac{(2000 \text{ psia})(1 \times 10^6 \text{ ft}^3)}{(0.963)(0.5925 \text{ psia} \cdot \text{ft}^3/\text{lbm} \cdot \text{R})(760 \text{ R})} = 4.612 \times 10^6 \text{ lbm}$$

The enthalpy departure factors in the reservoir and the surface are (from EES or Fig. A-29)

$$\left. \begin{aligned} T_{R1} &= \frac{T_m}{T'_{cr,m}} = \frac{760 \text{ R}}{375.0 \text{ R}} = 2.027 \\ P_{R1} &= \frac{P_m}{P'_{cr,m}} = \frac{2000 \text{ psia}}{678.3 \text{ psia}} = 2.949 \end{aligned} \right\} Z_{h1} = 0.703$$

$$\left. \begin{aligned} T_{R2} &= \frac{T_m}{T'_{cr,m}} = \frac{660 \text{ R}}{375.0 \text{ R}} = 1.76 \\ P_{R2} &= \frac{P_m}{P'_{cr,m}} = \frac{20 \text{ psia}}{678.3 \text{ psia}} = 0.0295 \end{aligned} \right\} Z_{h2} = 0.0112$$

The enthalpy change for the ideal gas mixture is


$$(h_1 - h_2)_{\text{ideal}} = c_p(T_1 - T_2) = (0.506 \text{ Btu/lbm} \cdot \text{R})(760 - 660)\text{R} = 50.6 \text{ Btu/lbm}$$

The enthalpy change with departure factors is

$$\begin{aligned} h_1 - h_2 &= (h_1 - h_2)_{\text{ideal}} - RT'_{cr,m}(Z_{h1} - Z_{h2}) \\ &= 50.6 - (0.1096)(375)(0.703 - 0.0112) \\ &= 22.12 \text{ Btu/lbm} \end{aligned}$$

The work input is then

$$W_{\text{in}} = m(h_1 - h_2) = (4.612 \times 10^6 \text{ lbm})(22.12 \text{ Btu/lbm}) = \mathbf{1.02 \times 10^8 \text{ Btu}}$$

13-65E  A gas mixture with known mass fractions is accelerated through a nozzle from a specified state to a specified pressure. For a specified isentropic efficiency, the exit temperature and the exit velocity of the mixture are to be determined.

Assumptions **1** Under specified conditions both N_2 and CO_2 can be treated as ideal gases, and the mixture as an ideal gas mixture. **2** The nozzle is adiabatic and thus heat transfer is negligible. **3** This is a steady-flow process. **4** Potential energy changes are negligible.

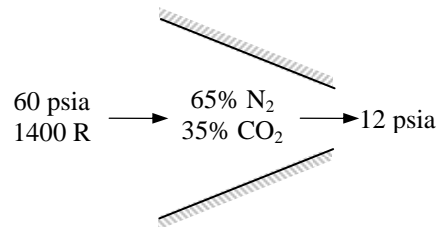
Properties The specific heats of N_2 and CO_2 are $c_{p,N_2} = 0.248$ Btu/lbm·R, $c_{v,N_2} = 0.177$ Btu/lbm·R, $c_{p,CO_2} = 0.203$ Btu/lbm·R, and $c_{v,CO_2} = 0.158$ Btu/lbm·R. (Table A-2E).

Analysis (a) Under specified conditions both N_2 and CO_2 can be treated as ideal gases, and the mixture as an ideal gas mixture. The c_p , c_v , and k values of this mixture are determined from

$$\begin{aligned} c_{p,m} &= \sum mf_i c_{p,i} = mf_{N_2} c_{p,N_2} + mf_{CO_2} c_{p,CO_2} \\ &= (0.65)(0.248) + (0.35)(0.203) \\ &= 0.2323 \text{ Btu/lbm} \cdot \text{R} \end{aligned}$$

$$\begin{aligned} c_{v,m} &= \sum mf_i c_{v,i} = mf_{N_2} c_{v,N_2} + mf_{CO_2} c_{v,CO_2} \\ &= (0.65)(0.177) + (0.35)(0.158) \\ &= 0.1704 \text{ Btu/lbm} \cdot \text{R} \end{aligned}$$

$$k_m = \frac{c_{p,m}}{c_{v,m}} = \frac{0.2323 \text{ Btu/lbm} \cdot \text{R}}{0.1704 \text{ Btu/lbm} \cdot \text{R}} = 1.363$$



Therefore, the N_2 - CO_2 mixture can be treated as a single ideal gas with above properties. Then the isentropic exit temperature can be determined from

$$T_{2s} = T_1 \left(\frac{P_2}{P_1} \right)^{(k-1)/k} = (1400 \text{ R}) \left(\frac{12 \text{ psia}}{60 \text{ psia}} \right)^{0.363/1.363} = 911.7 \text{ R}$$

From the definition of isentropic efficiency,

$$\eta_N = \frac{h_1 - h_2}{h_1 - h_{2s}} = \frac{c_p (T_1 - T_2)}{c_p (T_1 - T_{2s})} \rightarrow 0.88 = \frac{1400 - T_2}{1400 - 911.7} \rightarrow T_2 = \mathbf{970.3 \text{ R}}$$

(b) Noting that, $q = w = 0$, from the steady-flow energy balance relation,

$$\begin{aligned} \dot{E}_{in} - \dot{E}_{out} &= \Delta \dot{E}_{system} \stackrel{\phi^0(\text{steady})}{=} 0 \\ \dot{E}_{in} &= \dot{E}_{out} \end{aligned}$$

$$h_1 + V_1^2 / 2 = h_2 + V_2^2 / 2$$

$$0 = c_p (T_2 - T_1) + \frac{V_2^2 - V_1^2}{2} \stackrel{\phi^0}{}$$

$$V_2 = \sqrt{2c_p (T_1 - T_2)} = \sqrt{2(0.2323 \text{ Btu/lbm} \cdot \text{R})(1400 - 970.3) \text{ R} \left(\frac{25,037 \text{ ft}^2/\text{s}^2}{1 \text{ Btu/lbm}} \right)} = \mathbf{2236 \text{ ft/s}}$$



13-66E Problem 13-65E is reconsidered. The problem is first to be solved and then, for all other conditions being the same, the problem is to be resolved to determine the composition of the nitrogen and carbon dioxide that is required to have an exit velocity of 2200 ft/s at the nozzle exit.

Analysis The problem is solved using EES, and the solution is given below.

"Given"

mf_N2=0.65
mf_CO2=1-mf_N2
P1=60 [psia]
T1=1400 [R]
Vel1=0 [ft/s]
P2=12 [psia]
eta_N=0.88
"Vel2=2200 [ft/s]"

"Properties"

c_p_N2=0.248 [Btu/lbm-R]
c_v_N2=0.177 [Btu/lbm-R]
c_p_CO2=0.203 [Btu/lbm-R]
c_v_CO2=0.158 [Btu/lbm-R]
MM_N2=28 [lbm/lbmol]
MM_CO2=44 [lbm/lbmol]

"Analysis"

c_p_m=mf_N2*c_p_N2+mf_CO2*c_p_CO2
c_v_m=mf_N2*c_v_N2+mf_CO2*c_v_CO2
k_m=c_p_m/c_v_m
T2_s=T1*(P2/P1)^((k_m-1)/k_m)
eta_N=(T1-T2)/(T1-T2_s)
0=c_p_m*(T2-T1)+(Vel2^2-Vel1^2)/2*Convert(ft^2/s^2, Btu/lbm)
N_N2=mf_N2/MM_N2
N_CO2=mf_CO2/MM_CO2
N_total=N_N2+N_CO2
y_N2=N_N2/N_total
y_CO2=N_CO2/N_total

SOLUTION of the stated problem

c_p_CO2=0.203 [Btu/lbm-R]	c_p_m=0.2323 [Btu/lbm-R]	c_p_N2=0.248 [Btu/lbm-R]
c_v_CO2=0.158 [Btu/lbm-R]	c_v_m=0.1704 [Btu/lbm-R]	c_v_N2=0.177 [Btu/lbm-R]
eta_N=0.88	k_m=1.363	mf_CO2=0.35
mf_N2=0.65	MM_CO2=44 [lbm/lbmol]	MM_N2=28 [lbm/lbmol]
N_CO2=0.007955	N_N2=0.02321	N_total=0.03117
P1=60 [psia]	P2=12 [psia]	T1=1400 [R]
T2=970.3 [R]	T2_s=911.7 [R]	Vel1=0 [ft/s]
Vel2=2236 [ft/s]	y_CO2=0.2552	y_N2=0.7448

SOLUTION of the problem with exit velocity of 2200 ft/s

c_p_CO2=0.203 [Btu/lbm-R]	c_p_m=0.2285 [Btu/lbm-R]	c_p_N2=0.248 [Btu/lbm-R]
c_v_CO2=0.158 [Btu/lbm-R]	c_v_m=0.1688 [Btu/lbm-R]	c_v_N2=0.177 [Btu/lbm-R]
eta_N=0.88	k_m=1.354	mf_CO2=0.434
mf_N2=0.566	MM_CO2=44 [lbm/lbmol]	MM_N2=28 [lbm/lbmol]
N_CO2=0.009863	N_N2=0.02022	N_total=0.03008
P1=60 [psia]	P2=12 [psia]	T1=1400 [R]
T2=976.9 [R]	T2_s=919.3 [R]	Vel1=0 [ft/s]
Vel2=2200 [ft/s]	y_CO2=0.3279	y_N2=0.6721

13-67 A mixture of hydrogen and oxygen is considered. The entropy change of this mixture between the two specified states is to be determined.

Assumptions Hydrogen and oxygen are ideal gases.

Properties The gas constants of hydrogen and oxygen are 4.124 and 0.2598 kJ/kg·K, respectively (Table A-1).

Analysis The effective gas constant of this mixture is

$R = mf_{\text{H}_2}R_{\text{H}_2} + mf_{\text{O}_2}R_{\text{O}_2} = (0.33)(4.124) + (0.67)(0.2598) = 1.5350 \text{ kJ/kg} \cdot \text{K}$ Since the temperature of the two states is the same, the entropy change is determined from

$$s_2 - s_1 = -R \ln \frac{P_2}{P_1} = -(1.5350 \text{ kJ/kg} \cdot \text{K}) \ln \frac{150 \text{ kPa}}{750 \text{ kPa}} = \mathbf{2.470 \text{ kJ/kg} \cdot \text{K}}$$

13-68 A piston-cylinder device contains a gas mixture at a given state. Heat is transferred to the mixture. The amount of heat transfer and the entropy change of the mixture are to be determined.

Assumptions 1 Under specified conditions both H_2 and N_2 can be treated as ideal gases, and the mixture as an ideal gas mixture. **2** Kinetic and potential energy changes are negligible.

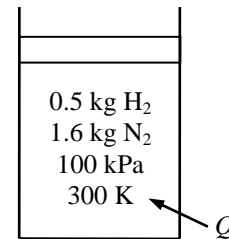
Properties The constant pressure specific heats of H_2 and N_2 at 450 K are 14.501 kJ/kg·K and 1.049 kJ/kg·K, respectively. (Table A-2b).

Analysis (a) Noting that $P_2 = P_1$ and $V_2 = 2V_1$,

$$\frac{P_2 V_2}{T_2} = \frac{P_1 V_1}{T_1} \longrightarrow T_2 = \frac{2V_1}{V_1} T_1 = 2T_1 = (2)(300 \text{ K}) = 600 \text{ K}$$

From the closed system energy balance relation,

$$\begin{aligned} E_{\text{in}} - E_{\text{out}} &= \Delta E_{\text{system}} \\ Q_{\text{in}} - W_{b,\text{out}} &= \Delta U \quad \rightarrow \quad Q_{\text{in}} = \Delta H \end{aligned}$$



since W_b and ΔU combine into ΔH for quasi-equilibrium constant pressure processes.

$$\begin{aligned} Q_{\text{in}} = \Delta H &= \Delta H_{\text{H}_2} + \Delta H_{\text{N}_2} = [mc_{p,\text{avg}}(T_2 - T_1)]_{\text{H}_2} + [mc_{p,\text{avg}}(T_2 - T_1)]_{\text{N}_2} \\ &= (0.5 \text{ kg})(14.501 \text{ kJ/kg} \cdot \text{K})(600 - 300) \text{ K} + (1.6 \text{ kg})(1.049 \text{ kJ/kg} \cdot \text{K})(600 - 300) \text{ K} \\ &= \mathbf{2679 \text{ kJ}} \end{aligned}$$

(b) Noting that the total mixture pressure, and thus the partial pressure of each gas, remains constant, the entropy change of the mixture during this process is

$$\begin{aligned} \Delta S_{\text{H}_2} &= [m(s_2 - s_1)]_{\text{H}_2} = m_{\text{H}_2} \left(c_p \ln \frac{T_2}{T_1} - R \ln \frac{P_2}{P_1} \right)_{\text{H}_2} = m_{\text{H}_2} \left(c_p \ln \frac{T_2}{T_1} \right)_{\text{H}_2} \\ &= (0.5 \text{ kg})(14.501 \text{ kJ/kg} \cdot \text{K}) \ln \frac{600 \text{ K}}{300 \text{ K}} \\ &= 5.026 \text{ kJ/K} \end{aligned}$$

$$\begin{aligned} \Delta S_{\text{N}_2} &= [m(s_2 - s_1)]_{\text{N}_2} = m_{\text{N}_2} \left(c_p \ln \frac{T_2}{T_1} - R \ln \frac{P_2}{P_1} \right)_{\text{N}_2} = m_{\text{N}_2} \left(c_p \ln \frac{T_2}{T_1} \right)_{\text{N}_2} \\ &= (1.6 \text{ kg})(1.049 \text{ kJ/kg} \cdot \text{K}) \ln \frac{600 \text{ K}}{300 \text{ K}} \\ &= 1.163 \text{ kJ/K} \end{aligned}$$

$$\Delta S_{\text{total}} = \Delta S_{\text{H}_2} + \Delta S_{\text{N}_2} = 5.026 \text{ kJ/K} + 1.163 \text{ kJ/K} = \mathbf{6.19 \text{ kJ/K}}$$



13-69 The temperatures and pressures of two gases forming a mixture in a mixing chamber are given. The mixture temperature and the rate of entropy generation are to be determined.

Assumptions 1 Under specified conditions both C_2H_6 and CH_4 can be treated as ideal gases, and the mixture as an ideal gas mixture. **2** The mixing chamber is insulated and thus there is no heat transfer. **3** There are no other forms of work involved. **3** This is a steady-flow process. **4** The kinetic and potential energy changes are negligible.

Properties The specific heats of C_2H_6 and CH_4 are 1.7662 kJ/kg·°C and 2.2537 kJ/kg·°C, respectively. (Table A-2b).

Analysis (a) The enthalpy of ideal gases is independent of pressure, and thus the two gases can be treated independently even after mixing. Noting that

$\dot{W} = \dot{Q} = 0$, the steady-flow energy balance equation reduces to

$$\begin{aligned}\dot{E}_{in} - \dot{E}_{out} &= \Delta \dot{E}_{system}^{\phi 0(\text{steady})} = 0 \\ \dot{E}_{in} &= \dot{E}_{out} \\ \sum \dot{m}_i h_i &= \sum \dot{m}_e h_e \\ 0 &= \sum \dot{m}_e h_e - \sum \dot{m}_i h_i = \dot{m}_{C_2H_6} (h_e - h_i)_{C_2H_6} + \dot{m}_{CH_4} (h_e - h_i)_{CH_4} \\ 0 &= [\dot{m} c_p (T_e - T_i)]_{C_2H_6} + [\dot{m} c_p (T_e - T_i)]_{CH_4}\end{aligned}$$

Using c_p values at room temperature and substituting, the exit temperature of the mixture becomes

$$\begin{aligned}0 &= (6 \text{ kg/s})(1.7662 \text{ kJ/kg} \cdot ^\circ\text{C})(T_m - 15^\circ\text{C}) + (3 \text{ kg/s})(2.2537 \text{ kJ/kg} \cdot ^\circ\text{C})(T_m - 60^\circ\text{C}) \\ T_m &= \mathbf{32.5^\circ\text{C}} \quad (305.5 \text{ K})\end{aligned}$$

(b) The rate of entropy change associated with this process is determined from an entropy balance on the mixing chamber,

$$\begin{aligned}\dot{S}_{in} - \dot{S}_{out} + \dot{S}_{gen} &= \Delta \dot{S}_{system}^{\approx 0} = 0 \\ [\dot{m}(s_1 - s_2)]_{C_2H_6} + [\dot{m}(s_1 - s_2)]_{CH_4} + \dot{S}_{gen} &= 0 \\ \dot{S}_{gen} &= [\dot{m}(s_2 - s_1)]_{C_2H_6} + [\dot{m}(s_2 - s_1)]_{CH_4}\end{aligned}$$

The molar flow rate of the two gases in the mixture is

$$\begin{aligned}\dot{N}_{C_2H_6} &= \left(\frac{\dot{m}}{M}\right)_{C_2H_6} = \frac{6 \text{ kg/s}}{30 \text{ kg/kmol}} = 0.2 \text{ kmol/s} \\ \dot{N}_{CH_4} &= \left(\frac{\dot{m}}{M}\right)_{CH_4} = \frac{4.5 \text{ kg/s}}{16 \text{ kg/kmol}} = 0.1875 \text{ kmol/s}\end{aligned}$$

Then the mole fraction of each gas becomes

$$\begin{aligned}y_{C_2H_6} &= \frac{0.2}{0.2 + 0.1875} = 0.5161 \\ y_{CH_4} &= \frac{0.1875}{0.2 + 0.1875} = 0.4839\end{aligned}$$

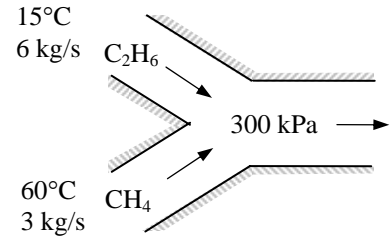
Thus,

$$\begin{aligned}(s_2 - s_1)_{C_2H_6} &= \left(c_p \ln \frac{T_2}{T_1} - R \ln \frac{y P_{m,2}}{P_1}\right)_{C_2H_6} = \left(c_p \ln \frac{T_2}{T_1} - R \ln y\right)_{C_2H_6} \\ &= (1.7662 \text{ kJ/kg} \cdot \text{K}) \ln \frac{305.5 \text{ K}}{288 \text{ K}} - (0.2765 \text{ kJ/kg} \cdot \text{K}) \ln(0.5161) = 0.2872 \text{ kJ/kg} \cdot \text{K}\end{aligned}$$

$$\begin{aligned}(s_2 - s_1)_{CH_4} &= \left(c_p \ln \frac{T_2}{T_1} - R \ln \frac{y P_{m,2}}{P_1}\right)_{CH_4} = \left(c_p \ln \frac{T_2}{T_1} - R \ln y\right)_{CH_4} \\ &= (2.2537 \text{ kJ/kg} \cdot \text{K}) \ln \frac{305.5 \text{ K}}{333 \text{ K}} - (0.5182 \text{ kJ/kg} \cdot \text{K}) \ln(0.4839) = 0.1821 \text{ kJ/kg} \cdot \text{K}\end{aligned}$$

Noting that $P_{m,2} = P_{i,1} = 300 \text{ kPa}$ and substituting,

$$\dot{S}_{gen} = (6 \text{ kg/s})(0.2872 \text{ kJ/kg} \cdot \text{K}) + (3 \text{ kg/s})(0.1821 \text{ kJ/kg} \cdot \text{K}) = \mathbf{2.27 \text{ kW/K}}$$





13-70 Problem 13-69 is reconsidered. The effect of the mass fraction of methane in the mixture on the mixture temperature and the rate of exergy destruction is to be investigated.

Analysis The problem is solved using EES, and the solution is given below.

"Given"

"1: C2H6, 2: CH4"

m_dot_total=9 [kg/s]

"mf_CH4=0.3333"

mf_C2H6=1-mf_CH4

m_dot_1=mf_C2H6*m_dot_total

m_dot_2=mf_CH4*m_dot_total

T1=(15+273) [K]

T2=(60+273) [K]

P=300 [kPa]

T0=(25+273) [K]

"Properties"

c_p_1=1.7662 [kJ/kg-K]

c_p_2=2.2537 [kJ/kg-K]

R_1=0.2765 [kJ/kg-K]

R_2=0.5182 [kJ/kg-K]

MM_1=30 [kg/kmol]

MM_2=16 [kg/kmol]

"Analysis"

0=m_dot_1*c_p_1*(T3-T1)+m_dot_2*c_p_2*(T3-T2)

N_dot_1=m_dot_1/MM_1

N_dot_2=m_dot_2/MM_2

N_dot_total=N_dot_1+N_dot_2

y_1=N_dot_1/N_dot_total

y_2=N_dot_2/N_dot_total

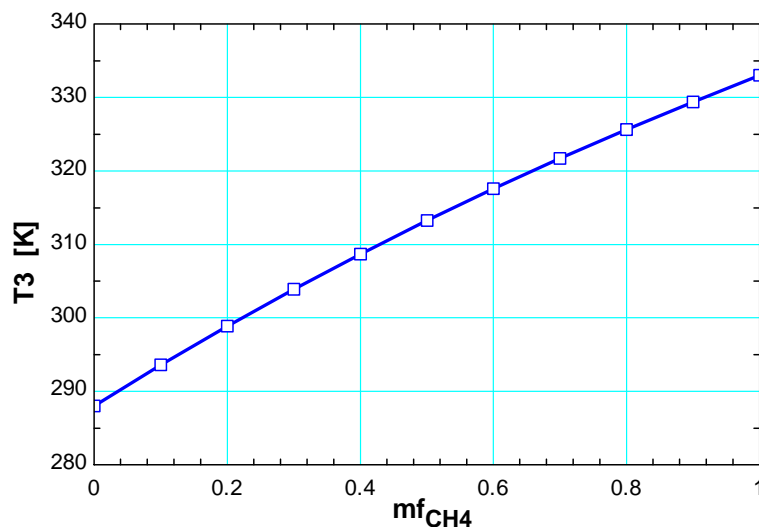
DELTA_s_1=c_p_1*ln(T3/T1)-R_1*ln(y_1)

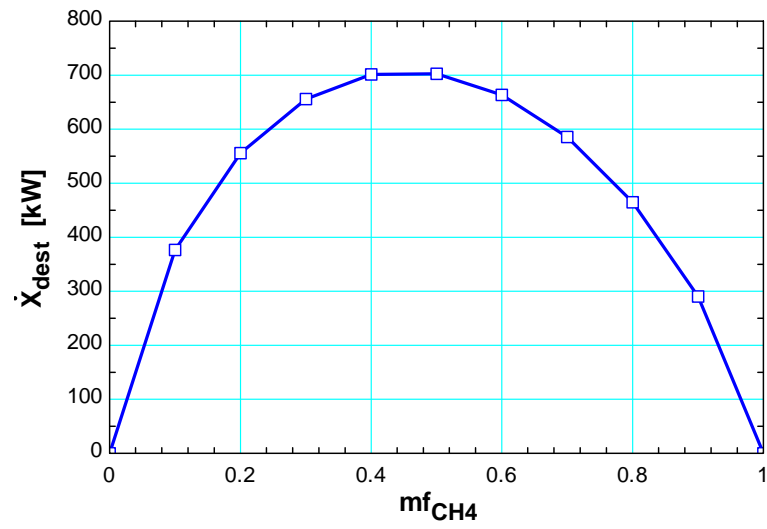
DELTA_s_2=c_p_2*ln(T3/T2)-R_2*ln(y_2)

S_dot_gen=m_dot_1*DELTA_s_1+m_dot_2*DELTA_s_2

X_dot_dest=T0*S_dot_gen

mf _{F2}	T3 [K]	X _{dest} [kW]
0	288	0
0.1	293.6	376.4
0.2	298.9	555.4
0.3	303.9	655.8
0.4	308.7	701.4
0.5	313.2	702.5
0.6	317.6	663.6
0.7	321.7	585.4
0.8	325.6	464.6
0.9	329.4	290.2
1	333	0.09793





13-71E In an air-liquefaction plant, it is proposed that the pressure and temperature of air be adiabatically reduced. It is to be determined whether this process is possible and the work produced is to be determined using Kay's rule and the departure charts.

Assumptions Air is a gas mixture with 21% O₂ and 79% N₂, by mole.

Properties The molar masses of O₂ and N₂ are 32.0 and 28.0 lbm/lbmol, respectively. The critical properties are 278.6 R, 736 psia for O₂ and 227.1 R and 492 psia for N₂ (Table A-1E).

Analysis To use Kay's rule, we need to determine the pseudo-critical temperature and pseudo-critical pressure of the mixture using the critical point properties of gases.

$$\begin{aligned} T'_{cr,m} &= \sum y_i T_{cr,i} = y_{O_2} T_{cr,O_2} + y_{N_2} T_{cr,N_2} \\ &= (0.21)(278.6 \text{ R}) + (0.79)(227.1 \text{ R}) = 237.9 \text{ R} \\ P'_{cr,m} &= \sum y_i P_{cr,i} = y_{O_2} P_{cr,O_2} + y_{N_2} P_{cr,N_2} \\ &= (0.21)(736 \text{ psia}) + (0.79)(492 \text{ psia}) = 543.2 \text{ psia} \end{aligned}$$

21% O₂
79% N₂
(by mole)
1500 psia
40°F

The enthalpy and entropy departure factors at the initial and final states are (from EES)

$$\left. \begin{aligned} T_{R1} &= \frac{T_{m1}}{T'_{cr,m}} = \frac{500 \text{ R}}{237.9 \text{ R}} = 2.102 \\ P_{R1} &= \frac{P_{m1}}{P'_{cr,m}} = \frac{1500 \text{ psia}}{432.2 \text{ psia}} = 3.471 \end{aligned} \right\} \begin{aligned} Z_{h1} &= 0.725 \\ Z_{s1} &= 0.339 \end{aligned}$$

$$\left. \begin{aligned} T_{R2} &= \frac{T_{m2}}{T'_{cr,m}} = \frac{360 \text{ R}}{237.9 \text{ R}} = 1.513 \\ P_{R2} &= \frac{P_{m2}}{P'_{cr,m}} = \frac{15 \text{ psia}}{432.2 \text{ psia}} = 0.0347 \end{aligned} \right\} \begin{aligned} Z_{h2} &= 0.0179 \\ Z_{s2} &= 0.00906 \end{aligned}$$

The enthalpy and entropy changes of the air under the ideal gas assumption is (Properties are from Table A-17E)

$$(h_2 - h_1)_{\text{ideal}} = 85.97 - 119.48 = -33.5 \text{ Btu/lbm}$$

$$(s_2 - s_1)_{\text{ideal}} = s_2^{\circ} - s_1^{\circ} - R \ln \frac{P_2}{P_1} = 0.50369 - 0.58233 - (0.06855) \ln \frac{15}{1500} = 0.2370 \text{ Btu/lbm} \cdot \text{R}$$

With departure factors, the enthalpy change (i.e., the work output) and the entropy change are

$$\begin{aligned} w_{\text{out}} &= h_1 - h_2 = (h_1 - h_2)_{\text{ideal}} - RT'_{cr} (Z_{h1} - Z_{h2}) \\ &= 33.5 - (0.06855)(237.9)(0.725 - 0.0179) = \mathbf{22.0 \text{ Btu/lbm}} \end{aligned}$$

$$\begin{aligned} s_2 - s_1 &= (s_2 - s_1)_{\text{ideal}} - R(Z_{s2} - Z_{s1}) \\ &= 0.2370 - (0.06855)(0.00906 - 0.339) = \mathbf{0.2596 \text{ Btu/lbm} \cdot \text{R}} \end{aligned}$$

The entropy change in this case is equal to the entropy generation during the process since the process is adiabatic. The positive value of entropy generation shows that this process is possible.

13-72 Heat is transferred to a gas mixture contained in a piston cylinder device. The initial state and the final temperature are given. The heat transfer is to be determined for the ideal gas and non-ideal gas cases.

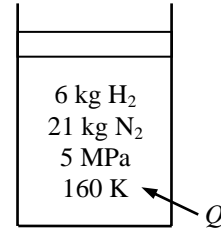
Properties The molar masses of H₂ and N₂ are 2.0, and 28.0 kg/kmol. (Table A-1).

Analysis From the energy balance relation,

$$\begin{aligned} E_{\text{in}} - E_{\text{out}} &= \Delta E \\ Q_{\text{in}} - W_{b,\text{out}} &= \Delta U \\ Q_{\text{in}} &= \Delta H = \Delta H_{\text{H}_2} + \Delta H_{\text{N}_2} = N_{\text{H}_2}(\bar{h}_2 - \bar{h}_1)_{\text{H}_2} + N_{\text{N}_2}(\bar{h}_2 - \bar{h}_1)_{\text{N}_2} \end{aligned}$$

since W_b and ΔU combine into ΔH for quasi-equilibrium constant pressure processes

$$\begin{aligned} N_{\text{H}_2} &= \frac{m_{\text{H}_2}}{M_{\text{H}_2}} = \frac{6 \text{ kg}}{2 \text{ kg/kmol}} = 3 \text{ kmol} \\ N_{\text{N}_2} &= \frac{m_{\text{N}_2}}{M_{\text{N}_2}} = \frac{21 \text{ kg}}{28 \text{ kg/kmol}} = 0.75 \text{ kmol} \end{aligned}$$



(a) Assuming ideal gas behavior, the inlet and exit enthalpies of H₂ and N₂ are determined from the ideal gas tables to be

$$\begin{aligned} \text{H}_2: \quad \bar{h}_1 &= \bar{h}_{@160 \text{ K}} = 4,535.4 \text{ kJ/kmol}, & \bar{h}_2 &= \bar{h}_{@200 \text{ K}} = 5,669.2 \text{ kJ/kmol} \\ \text{N}_2: \quad \bar{h}_1 &= \bar{h}_{@160 \text{ K}} = 4,648 \text{ kJ/kmol}, & \bar{h}_2 &= \bar{h}_{@200 \text{ K}} = 5,810 \text{ kJ/kmol} \end{aligned}$$

Thus, $Q_{\text{ideal}} = 3 \times (5,669.2 - 4,535.4) + 0.75 \times (5,810 - 4,648) = \mathbf{4273 \text{ kJ}}$

(b) Using Amagat's law and the generalized enthalpy departure chart, the enthalpy change of each gas is determined to be

$$\text{H}_2: \left. \begin{aligned} T_{R_1, \text{H}_2} &= \frac{T_{m,1}}{T_{\text{cr}, \text{H}_2}} = \frac{160}{33.3} = 4.805 \\ P_{R_1, \text{H}_2} &= P_{R_2, \text{H}_2} = \frac{P_m}{P_{\text{cr}, \text{H}_2}} = \frac{5}{1.30} = 3.846 \\ T_{R_2, \text{H}_2} &= \frac{T_{m,2}}{T_{\text{cr}, \text{H}_2}} = \frac{200}{33.3} = 6.006 \end{aligned} \right\} \begin{aligned} Z_{h_1} &\cong 0 \\ Z_{h_2} &\cong 0 \end{aligned} \quad (\text{Fig. A-29})$$

Thus H₂ can be treated as an ideal gas during this process.

$$\text{N}_2: \left. \begin{aligned} T_{R_1, \text{N}_2} &= \frac{T_{m,1}}{T_{\text{cr}, \text{N}_2}} = \frac{160}{126.2} = 1.27 \\ P_{R_1, \text{N}_2} &= P_{R_2, \text{N}_2} = \frac{P_m}{P_{\text{cr}, \text{N}_2}} = \frac{5}{3.39} = 1.47 \\ T_{R_2, \text{N}_2} &= \frac{T_{m,2}}{T_{\text{cr}, \text{N}_2}} = \frac{200}{126.2} = 1.58 \end{aligned} \right\} \begin{aligned} Z_{h_1} &= 1.3 \\ Z_{h_2} &= 0.7 \end{aligned} \quad (\text{Fig. A-29})$$

Therefore,

$$\begin{aligned} (\bar{h}_2 - \bar{h}_1)_{\text{H}_2} &= (\bar{h}_2 - \bar{h}_1)_{\text{H}_2, \text{ideal}} = 5,669.2 - 4,535.4 = 1,133.8 \text{ kJ/kmol} \\ (\bar{h}_2 - \bar{h}_1)_{\text{N}_2} &= R_u T_{\text{cr}} (Z_{h_1} - Z_{h_2}) + (\bar{h}_2 - \bar{h}_1)_{\text{ideal}} \\ &= (8.314 \text{ kPa} \cdot \text{m}^3 / \text{kmol} \cdot \text{K})(126.2 \text{ K})(1.3 - 0.7) + (5,810 - 4,648) \text{ kJ/kmol} = 1,791.5 \text{ kJ/kmol} \end{aligned}$$

Substituting,

$$Q_{\text{in}} = (3 \text{ kmol})(1,133.8 \text{ kJ/kmol}) + (0.75 \text{ kmol})(1,791.5 \text{ kJ/kmol}) = \mathbf{4745 \text{ kJ}}$$

13-73 Heat is transferred to a gas mixture contained in a piston cylinder device discussed in previous problem. The total entropy change and the exergy destruction are to be determined for two cases.

Analysis The entropy generated during this process is determined by applying the entropy balance on an *extended system* that includes the piston-cylinder device and its immediate surroundings so that the boundary temperature of the extended system is the environment temperature at all times. It gives

$$S_{\text{in}} - S_{\text{out}} + S_{\text{gen}} = \Delta S_{\text{system}}$$

$$\frac{Q_{\text{in}}}{T_{\text{boundary}}} + S_{\text{gen}} = \Delta S_{\text{water}} \rightarrow S_{\text{gen}} = m(s_2 - s_1) - \frac{Q_{\text{in}}}{T_{\text{surr}}}$$

Then the exergy destroyed during a process can be determined from its definition $X_{\text{destroyed}} = T_0 S_{\text{gen}}$.

(a) Noting that the total mixture pressure, and thus the partial pressure of each gas, remains constant, the entropy change of a component in the mixture during this process is

$$\Delta S_i = m_i \left(c_p \ln \frac{T_2}{T_1} - R \ln \frac{P_2}{P_1} \right) = m_i c_{p,i} \ln \frac{T_2}{T_1}$$

Assuming ideal gas behavior and using c_p values at the average temperature, the ΔS of H_2 and N_2 are determined from

$$\Delta S_{\text{H}_2, \text{ideal}} = (6 \text{ kg}) (13.60 \text{ kJ/kg} \cdot \text{K}) \ln \frac{200 \text{ K}}{160 \text{ K}} = 18.21 \text{ kJ/K}$$

$$\Delta S_{\text{N}_2, \text{ideal}} = (21 \text{ kg}) (1.039 \text{ kJ/kg} \cdot \text{K}) \ln \frac{200 \text{ K}}{160 \text{ K}} = 4.87 \text{ kJ/K}$$

and

$$S_{\text{gen}} = 18.21 \text{ kJ/K} + 4.87 \text{ kJ/K} - \frac{4273 \text{ kJ}}{303 \text{ K}} = \mathbf{8.98 \text{ kJ/K}}$$

$$X_{\text{destroyed}} = T_0 S_{\text{gen}} = (303 \text{ K})(8.98 \text{ kJ/K}) = \mathbf{2721 \text{ kJ}}$$

(b) Using Amagat's law and the generalized entropy departure chart, the entropy change of each gas is determined to be

$$\text{H}_2: \left. \begin{aligned} T_{R_1, \text{H}_2} &= \frac{T_{m,1}}{T_{\text{cr}, \text{H}_2}} = \frac{160}{33.3} = 4.805 \\ P_{R_1, \text{H}_2} &= P_{R_2, \text{H}_2} = \frac{P_m}{P_{\text{cr}, \text{H}_2}} = \frac{5}{1.30} = 3.846 \\ T_{R_2, \text{H}_2} &= \frac{T_{m,2}}{T_{\text{cr}, \text{H}_2}} = \frac{200}{33.3} = 6.006 \end{aligned} \right\} \begin{aligned} Z_{s_1} &\cong 1 \\ Z_{s_2} &\cong 1 \end{aligned} \quad (\text{Table A-30})$$

Thus H_2 can be treated as an ideal gas during this process.

$$\text{N}_2: \left. \begin{aligned} T_{R_1, \text{N}_2} &= \frac{T_{m,1}}{T_{\text{cr}, \text{N}_2}} = \frac{160}{126.2} = 1.268 \\ P_{R_1, \text{N}_2} &= P_{R_2, \text{N}_2} = \frac{P_m}{P_{\text{cr}, \text{N}_2}} = \frac{5}{3.39} = 1.475 \\ T_{R_2, \text{N}_2} &= \frac{T_{m,2}}{T_{\text{cr}, \text{N}_2}} = \frac{200}{126.2} = 1.585 \end{aligned} \right\} \begin{aligned} Z_{s_1} &= 0.8 \\ Z_{s_2} &= 0.4 \end{aligned} \quad (\text{Table A-30})$$

Therefore,

$$\Delta S_{\text{H}_2} = \Delta S_{\text{H}_2, \text{ideal}} = 18.21 \text{ kJ/K}$$

$$\begin{aligned}\Delta S_{\text{N}_2} &= N_{\text{N}_2} R_u (Z_{s_1} - Z_{s_2}) + \Delta S_{\text{N}_2, \text{ideal}} \\ &= (0.75 \text{ kmol})(8.314 \text{ kPa} \cdot \text{m}^3/\text{kmol} \cdot \text{K})(0.8 - 0.4) + (4.87 \text{ kJ/K}) \\ &= 7.37 \text{ kJ/K}\end{aligned}$$

$$\Delta S_{\text{surr}} = \frac{Q_{\text{surr}}}{T_0} = \frac{-4745 \text{ kJ}}{303 \text{ K}} = -15.66 \text{ kJ/K}$$

and

$$S_{\text{gen}} = 18.21 \text{ kJ/K} + 7.37 \text{ kJ/K} - \frac{4745 \text{ kJ}}{303 \text{ K}} = \mathbf{9.92 \text{ kJ/K}}$$

$$X_{\text{destroyed}} = T_0 S_{\text{gen}} = (303 \text{ K})(9.92 \text{ kJ/K}) = \mathbf{3006 \text{ kJ}}$$

13-74 Air is compressed isothermally in a steady-flow device. The power input to the compressor and the rate of heat rejection are to be determined for ideal and non-ideal gas cases.

Assumptions 1 This is a steady-flow process. **2** The kinetic and potential energy changes are negligible.

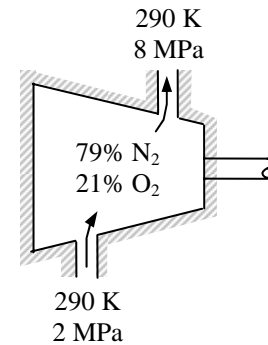
Properties The molar mass of air is 28.97 kg/kmol. (Table A-1).

Analysis The mass flow rate of air can be expressed in terms of the mole numbers as

$$\dot{N} = \frac{\dot{m}}{M} = \frac{1.75 \text{ kg/s}}{28.97 \text{ kg/kmol}} = 0.06041 \text{ kmol/s}$$

(a) Assuming ideal gas behavior, the Δh and Δs of air during this process is

$$\begin{aligned} \Delta \bar{h} &= 0 \text{ (isothermal process)} \\ \Delta \bar{s} &= \bar{c}_p \ln \frac{T_2}{T_1} - R_u \ln \frac{P_2}{P_1} = -R_u \ln \frac{P_2}{P_1} \\ &= -(8.314 \text{ kJ/kg} \cdot \text{K}) \ln \frac{8 \text{ MPa}}{2 \text{ MPa}} = -11.53 \text{ kJ/kmol} \cdot \text{K} \end{aligned}$$



Disregarding any changes in kinetic and potential energies, the steady-flow energy balance equation for the isothermal process of an ideal gas reduces to

$$\begin{aligned} \dot{E}_{\text{in}} - \dot{E}_{\text{out}} &= \Delta \dot{E}_{\text{system}}^{\text{0 (steady)}} = 0 \\ \dot{E}_{\text{in}} &= \dot{E}_{\text{out}} \\ \dot{W}_{\text{in}} + \dot{N}\bar{h}_1 &= \dot{Q}_{\text{out}} + \dot{N}\bar{h}_2 \\ \dot{W}_{\text{in}} - \dot{Q}_{\text{out}} &= \dot{N}\Delta \bar{h}^{\text{0}} = 0 \longrightarrow \dot{W}_{\text{in}} = \dot{Q}_{\text{out}} \end{aligned}$$

Also for an isothermal, internally reversible process the heat transfer is related to the entropy change by

$$Q = T\Delta S = NT\Delta \bar{s},$$

$$\dot{Q} = \dot{N}T\Delta \bar{s} = (0.06041 \text{ kmol/s})(290 \text{ K})(-11.53 \text{ kJ/kmol} \cdot \text{K}) = -201.9 \text{ kW} \rightarrow \dot{Q}_{\text{out}} = 201.9 \text{ kW}$$

Therefore,

$$\dot{W}_{\text{in}} = \dot{Q}_{\text{out}} = \mathbf{201.9 \text{ kW}}$$

(b) Using Amagat's law and the generalized charts, the enthalpy and entropy changes of each gas are determined from

$$\begin{aligned} \bar{h}_2 - \bar{h}_1 &= R_u T_{cr} (Z_{h_2} - Z_{h_1}) + (\bar{h}_2 - \bar{h}_1)_{\text{ideal}}^{\text{0}} \\ \bar{s}_2 - \bar{s}_1 &= R_u (Z_{s_1} - Z_{s_2}) + (\bar{s}_2 - \bar{s}_1)_{\text{ideal}} \end{aligned}$$

where

$$\text{N}_2: \left. \begin{aligned} P_{R_1} &= \frac{P_{m,1}}{P_{\text{cr},\text{N}_2}} = \frac{2}{3.39} = 0.59 \\ T_{R_1} = T_{R_2} &= \frac{T_m}{T_{\text{cr},\text{N}_2}} = \frac{290}{126.2} = 2.298 \\ P_{R_2} &= \frac{P_{m,2}}{P_{\text{cr},\text{N}_2}} = \frac{8}{3.39} = 2.36 \end{aligned} \right\} \begin{aligned} Z_{h_1} &= 0.1154, Z_{s_1} = 0.05136 \\ Z_{h_2} &= 0.4136, Z_{s_2} = 0.1903 \end{aligned} \quad \text{(Figures A-29 and A-30 or EES)}$$

Note that we used EES to obtain enthalpy and entropy departure factors. The accurate readings like these are not possible with Figures A-29 and A-30. EES has built-in functions for enthalpy departure and entropy departure factors in the following format:

$$\begin{aligned} Z_{h1} &= \text{ENTHDEP}(T_R1, P_R1) \text{ "the function that returns enthalpy departure factor"} \\ Z_{s1} &= \text{ENTRDEP}(T_R1, P_R1) \text{ "the function that returns entropy departure factor"} \end{aligned}$$

$$\text{O}_2: \left. \begin{aligned} P_{R_1} &= \frac{P_{m,1}}{P_{\text{cr,O}_2}} = \frac{2}{5.08} = 0.3937 \\ T_{R_1} &= T_{R_2} = \frac{T_m}{T_{\text{cr,O}_2}} = \frac{290}{154.8} = 1.873 \\ P_{R_2} &= \frac{P_{m,2}}{P_{\text{cr,O}_2}} = \frac{8}{5.08} = 1.575 \end{aligned} \right\} \begin{aligned} Z_{h_1} &= 0.1296, Z_{s_1} = 0.05967 \\ Z_{h_2} &= 0.4956, Z_{s_2} = 0.2313 \end{aligned} \quad \text{(Figures A-29 and A-30 or EES)}$$

Then,

$$\begin{aligned} \bar{h}_2 - \bar{h}_1 &= y_i \Delta \bar{h}_i = y_{\text{N}_2} (\bar{h}_2 - \bar{h}_1)_{\text{N}_2} + y_{\text{O}_2} (\bar{h}_2 - \bar{h}_1)_{\text{O}_2} \\ &= (0.79)(8.314)(126.2)(0.1154 - 0.4136) + (0.21)(8.314)(154.8)(0.1296 - 0.4956) + 0 \\ &= -346.1 \text{ kJ/kmol} \end{aligned}$$

$$\begin{aligned} \bar{s}_2 - \bar{s}_1 &= y_i \Delta \bar{s}_i = y_{\text{N}_2} (\bar{s}_2 - \bar{s}_1)_{\text{N}_2} + y_{\text{O}_2} (\bar{s}_2 - \bar{s}_1)_{\text{O}_2} \\ &= (0.79)(8.314)(0.05136 - 0.1903) + (0.21)(8.314)(0.05967 - 0.2313) + (-11.53) \\ &= -12.74 \text{ kJ/kmol} \cdot \text{K} \end{aligned}$$

Thus,

$$\dot{Q}_{\text{out}} = -\dot{N}T\Delta\bar{s} = -(0.06041 \text{ kmol/s})(290 \text{ K})(-12.74 \text{ kJ/kmol} \cdot \text{K}) = \mathbf{223.2 \text{ kW}}$$

$$\dot{E}_{\text{in}} - \dot{E}_{\text{out}} = \Delta \dot{E}_{\text{system}} \overset{\phi^0(\text{steady})}{=} 0$$

$$\dot{E}_{\text{in}} = \dot{E}_{\text{out}}$$

$$\dot{W}_{\text{in}} + \dot{N}\bar{h}_1 = \dot{Q}_{\text{out}} + \dot{N}\bar{h}_2$$

$$\dot{W}_{\text{in}} = \dot{Q}_{\text{out}} + \dot{N}(\bar{h}_2 - \bar{h}_1) \longrightarrow \dot{W}_{\text{in}} = 223.2 \text{ kW} + (0.06041 \text{ kmol/s})(-346.1 \text{ kJ/kmol}) = \mathbf{202.3 \text{ kW}}$$



13-75 Problem 13-74 is reconsidered. The results obtained by assuming ideal behavior, real gas behavior with Amagat's law, and real gas behavior with EES data are to be compared.

Analysis The problem is solved using EES, and the solution is given below.

"Given"

y_N2=0.79
y_O2=0.21
T=290 [K]
P1=2000 [kPa]
P2=8000 [kPa]
m_dot=1.75 [kg/s]

"Properties"

R_u=8.314 [kJ/kmol-K]
M_air=molar mass(air)
T_cr_N2=126.2 [K]
T_cr_O2=154.8 [K]
P_cr_N2=3390 [kPa]
P_cr_O2=5080 [kPa]

"Analysis"

"Ideal gas"

N_dot=m_dot/M_air
DELTAh_ideal=0 "isothermal process"
DELTA s_ideal=-R_u*ln(P2/P1) "isothermal process"
Q_dot_in_ideal=N_dot*T*DELTA s_ideal
W_dot_in_ideal=-Q_dot_in_ideal

"Amagad's law"

T_R1_N2=T/T_cr_N2
P_R1_N2=P1/P_cr_N2
Z_h1_N2=ENTHDEP(T_R1_N2, P_R1_N2) "the function that returns enthalpy departure factor"
Z_s1_N2=ENTRDEP(T_R1_N2, P_R1_N2) "the function that returns entropy departure factor"
T_R2_N2=T/T_cr_N2
P_R2_N2=P2/P_cr_N2
Z_h2_N2=ENTHDEP(T_R2_N2, P_R2_N2) "the function that returns enthalpy departure factor"
Z_s2_N2=ENTRDEP(T_R2_N2, P_R2_N2) "the function that returns entropy departure factor"
T_R1_O2=T/T_cr_O2
P_R1_O2=P1/P_cr_O2
Z_h1_O2=ENTHDEP(T_R1_O2, P_R1_O2) "the function that returns enthalpy departure factor"
Z_s1_O2=ENTRDEP(T_R1_O2, P_R1_O2) "the function that returns entropy departure factor"
T_R2_O2=T/T_cr_O2
P_R2_O2=P2/P_cr_O2
Z_h2_O2=ENTHDEP(T_R2_O2, P_R2_O2) "the function that returns enthalpy departure factor"
Z_s2_O2=ENTRDEP(T_R2_O2, P_R2_O2) "the function that returns entropy departure factor"
DELTAh=DELTAh_ideal-(y_N2*R_u*T_cr_N2*(Z_h2_N2-Z_h1_N2)+y_O2*R_u*T_cr_O2*(Z_h2_O2-Z_h1_O2))
DELTA s=DELTA s_ideal-(y_N2*R_u*(Z_s2_N2-Z_s1_N2)+y_O2*R_u*(Z_s2_O2-Z_s1_O2))
Q_dot_in_Amagad=N_dot*T*DELTA s
W_dot_in_Amagad=-Q_dot_in_Amagad+N_dot*DELTAh

"EES"

h_EES[1]=y_N2*enthalpy(Nitrogen,T=T,P=P1)+y_O2*enthalpy(Oxygen,T=T,P=P1)
h_EES[2]=y_N2*enthalpy(Nitrogen,T=T,P=P2)+y_O2*enthalpy(Oxygen,T=T,P=P2)
s_EES[1]=y_N2*entropy(Nitrogen,T=T,P=P1)+y_O2*entropy(Oxygen,T=T,P=P1)
s_EES[2]=y_N2*entropy(Nitrogen,T=T,P=P2)+y_O2*entropy(Oxygen,T=T,P=P2)
DELTAh_EES=h_EES[2]-h_EES[1]

$\Delta h_{EES} = s_{EES}[2] - s_{EES}[1]$
 $\dot{Q}_{in,EES} = \dot{N} T \Delta h_{EES}$
 $\dot{W}_{in,EES} = -\dot{Q}_{in,EES} + \dot{N} \Delta h_{EES}$

SOLUTION

$\Delta h = -346.1$ [kJ/kmol]
 $\Delta h_{ideal} = 0$ [kJ/kmol]
 $\Delta h_{EES} = -12.72$ [kJ/kmol-K]
 $h_{EES}[1] = 6473$ [kJ/kmol]
 $M_{air} = 28.97$ [kg/kmol]
 $\dot{N} = 0.06041$ [kmol/s]
 $P_2 = 8000$ [kPa]
 $P_{cr,O_2} = 5080$ [kPa]
 $P_{R1,O_2} = 0.3937$
 $P_{R2,O_2} = 1.575$
 $\dot{Q}_{in,EES} = -222.9$ [kW]
 $R_u = 8.314$ [kPa·m³/kmol-K]
 $s_{EES}[2] = 112.5$ [kJ/kmol-K]
 $T_{cr,N_2} = 126.2$ [K]
 $T_{R1,N_2} = 2.298$
 $T_{R2,N_2} = 2.298$
 $\dot{W}_{in,Amagad} = 202.3$ [kW]
 $\dot{W}_{in,ideal} = 201.9$ [kW]
 $y_{O_2} = 0.21$
 $Z_{h1,O_2} = 0.1296$
 $Z_{h2,O_2} = 0.4956$
 $Z_{s1,O_2} = 0.05967$
 $Z_{s2,O_2} = 0.2313$

$\Delta h_{EES} = -384.3$ [kJ/kmol]
 $\Delta h_{EES} = -12.74$ [kJ/kmol-K]
 $\Delta h_{ideal} = -11.53$ [kJ/kmol-K]
 $h_{EES}[2] = 6089$ [kJ/kmol]
 $\dot{m} = 1.75$ [kg/s]
 $P_1 = 2000$ [kPa]
 $P_{cr,N_2} = 3390$ [kPa]
 $P_{R1,N_2} = 0.59$
 $P_{R2,N_2} = 2.36$
 $\dot{Q}_{in,Amagad} = -223.2$ [kW]
 $\dot{Q}_{in,ideal} = -201.9$ [kW]
 $s_{EES}[1] = 125.3$ [kJ/kmol-K]
 $T = 290$ [K]
 $T_{cr,O_2} = 154.8$ [K]
 $T_{R1,O_2} = 1.873$
 $T_{R2,O_2} = 1.873$
 $\dot{W}_{in,EES} = 199.7$ [kW]
 $y_{N_2} = 0.79$
 $Z_{h1,N_2} = 0.1154$
 $Z_{h2,N_2} = 0.4136$
 $Z_{s1,N_2} = 0.05136$
 $Z_{s2,N_2} = 0.1903$

13-76 Two mass streams of two different ideal gases are mixed in a steady-flow chamber while receiving energy by heat transfer from the surroundings. Expressions for the final temperature and the exit volume flow rate are to be obtained and two special cases are to be evaluated.

Assumptions Kinetic and potential energy changes are negligible.

Analysis (a) Mass and Energy Balances for the mixing process:

$$\dot{m}_1 + \dot{m}_2 = \dot{m}_3$$

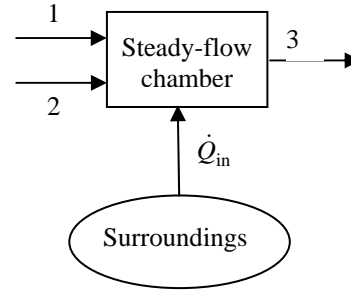
$$\dot{m}_1 h_1 + \dot{m}_2 h_2 + \dot{Q}_{in} = \dot{m}_3 h_3$$

$$h = C_p T$$

$$\dot{m}_1 C_{p,1} T_1 + \dot{m}_2 C_{p,2} T_2 + \dot{Q}_{in} = \dot{m}_3 C_{p,m} T_3$$

$$C_{p,m} = \frac{\dot{m}_1}{\dot{m}_3} C_{p,1} + \frac{\dot{m}_2}{\dot{m}_3} C_{p,2}$$

$$T_3 = \frac{\dot{m}_1 C_{p,1}}{\dot{m}_3 C_{p,m}} T_1 + \frac{\dot{m}_2 C_{p,2}}{\dot{m}_3 C_{p,m}} T_2 + \frac{\dot{Q}_{in}}{\dot{m}_3 C_{p,m}}$$



(b) The expression for the exit volume flow rate is obtained as follows:

$$\dot{V}_3 = \dot{m}_3 v_3 = \dot{m}_3 \frac{R_3 T_3}{P_3}$$

$$\dot{V}_3 = \frac{\dot{m}_3 R_3}{P_3} \left[\frac{\dot{m}_1 C_{p,1}}{\dot{m}_3 C_{p,m}} T_1 + \frac{\dot{m}_2 C_{p,2}}{\dot{m}_3 C_{p,m}} T_2 + \frac{\dot{Q}_{in}}{\dot{m}_3 C_{p,m}} \right]$$

$$\dot{V}_3 = \frac{C_{p,1} R_3}{C_{p,m} R_1} \frac{\dot{m}_1 R_1 T_1}{P_3} + \frac{C_{p,2} R_3}{C_{p,m} R_2} \frac{\dot{m}_2 R_2 T_2}{P_3} + \frac{R_3 \dot{Q}_{in}}{P_3 C_{p,m}}$$

$$P_3 = P_1 = P_2$$

$$\dot{V}_3 = \frac{C_{p,1} R_3}{C_{p,m} R_1} \dot{V}_1 + \frac{C_{p,2} R_3}{C_{p,m} R_2} \dot{V}_2 + \frac{R_3 \dot{Q}_{in}}{P_3 C_{p,m}}$$

$$R = \frac{R_u}{M}, \quad \frac{R_3}{R_1} = \frac{R_u M_1}{M_3 R_u} = \frac{M_1}{M_3}, \quad \frac{R_3}{R_2} = \frac{M_2}{M_3}$$

$$\dot{V}_3 = \frac{C_{p,1} M_1}{C_{p,m} M_3} \dot{V}_1 + \frac{C_{p,2} M_2}{C_{p,m} M_3} \dot{V}_2 + \frac{R_u \dot{Q}_{in}}{P_3 M_3 C_{p,m}}$$

The mixture molar mass M_3 is found as follows:

$$M_3 = \sum y_i M_i, \quad y_i = \frac{m_{fi} / M_i}{\sum m_{fi} / M_i}, \quad m_{fi} = \sum \dot{m}_i$$

(c) For adiabatic mixing \dot{Q}_{in} is zero, and the mixture volume flow rate becomes

$$\dot{V}_3 = \frac{C_{p,1} M_1}{C_{p,m} M_3} \dot{V}_1 + \frac{C_{p,2} M_2}{C_{p,m} M_3} \dot{V}_2$$

(d) When adiabatically mixing the same two ideal gases, the mixture volume flow rate becomes

$$M_3 = M_1 = M_2$$

$$C_{p,3} = C_{p,1} = C_{p,2}$$

$$\dot{V}_3 = \dot{V}_1 + \dot{V}_2$$

Special Topic: Chemical Potential and the Separation Work of Mixtures

13-77C No, a process that separates a mixture into its components without requiring any work (exergy) input is impossible since such a process would violate the 2nd law of thermodynamics.

13-78C Yes, the volume of the mixture can be more or less than the sum of the initial volumes of the mixing liquids because of the attractive or repulsive forces acting between dissimilar molecules.

13-79C The person who claims that the temperature of the mixture can be higher than the temperatures of the components is right since the total enthalpy of the mixture of two components at the same pressure and temperature, in general, is not equal to the sum of the total enthalpies of the individual components before mixing, the difference being the enthalpy (or heat) of mixing, which is the heat released or absorbed as two or more components are mixed isothermally.

13-80C Mixtures or solutions in which the effects of molecules of different components on each other are negligible are called ideal solutions (or ideal mixtures). The ideal-gas mixture is just one category of ideal solutions. For ideal solutions, the enthalpy change and the volume change due to mixing are zero, but the entropy change is not. The chemical potential of a component of an ideal mixture is independent of the identity of the other constituents of the mixture. The chemical potential of a component in an ideal mixture is equal to the Gibbs function of the pure component.

13-81 Brackish water is used to produce fresh water. The minimum power input and the minimum height the brackish water must be raised by a pump for reverse osmosis are to be determined.

Assumptions 1 The brackish water is an ideal solution since it is dilute. **2** The total dissolved solids in water can be treated as table salt (NaCl). **3** The environment temperature is also 12°C.

Properties The molar masses of water and salt are $M_w = 18.0 \text{ kg/kmol}$ and $M_s = 58.44 \text{ kg/kmol}$. The gas constant of pure water is $R_w = 0.4615 \text{ kJ/kg}\cdot\text{K}$ (Table A-1). The density of fresh water is 1000 kg/m^3 .

Analysis First we determine the mole fraction of pure water in brackish water using Eqs. 13-4 and 13-5. Noting that $mf_s = 0.00078$ and $mf_w = 1 - mf_s = 0.99922$,

$$M_m = \frac{1}{\sum \frac{mf_i}{M_i}} = \frac{1}{\frac{mf_s}{M_s} + \frac{mf_w}{M_w}} = \frac{1}{\frac{0.00078}{58.44} + \frac{0.99922}{18.0}} = 18.01 \text{ kg/kmol}$$

$$y_i = mf_i \frac{M_m}{M_i} \rightarrow y_w = mf_w \frac{M_m}{M_w} = (0.99922) \frac{18.01 \text{ kg/kmol}}{18.0 \text{ kg/kmol}} = 0.99976$$

The minimum work input required to produce 1 kg of freshwater from brackish water is

$$w_{\min, \text{in}} = R_w T_0 \ln(1/y_w) = (0.4615 \text{ kJ/kg}\cdot\text{K})(285.15 \text{ K}) \ln(1/0.99976) = 0.03159 \text{ kJ/kg fresh water}$$

Therefore, 0.03159 kJ of work is needed to produce 1 kg of fresh water is mixed with seawater reversibly. Therefore, the required power input to produce fresh water at the specified rate is

$$\dot{W}_{\min, \text{in}} = \rho \dot{V} w_{\min, \text{in}} = (1000 \text{ kg/m}^3)(0.280 \text{ m}^3/\text{s})(0.03159 \text{ kJ/kg}) \left(\frac{1 \text{ kW}}{1 \text{ kJ/s}} \right) = \mathbf{8.85 \text{ kW}}$$

The minimum height to which the brackish water must be pumped is

$$\Delta z_{\min} = \frac{w_{\min, \text{in}}}{g} = \left(\frac{0.03159 \text{ kJ/kg}}{9.81 \text{ m/s}^2} \right) \left(\frac{1 \text{ kg}\cdot\text{m/s}^2}{1 \text{ N}} \right) \left(\frac{1000 \text{ N}\cdot\text{m}}{1 \text{ kJ}} \right) = \mathbf{3.22 \text{ m}}$$

13-82 A river is discharging into the ocean at a specified rate. The amount of power that can be generated is to be determined.

Assumptions 1 The seawater is an ideal solution since it is dilute. **2** The total dissolved solids in water can be treated as table salt (NaCl). **3** The environment temperature is also 15°C.

Properties The molar masses of water and salt are $M_w = 18.0 \text{ kg/kmol}$ and $M_s = 58.44 \text{ kg/kmol}$. The gas constant of pure water is $R_w = 0.4615 \text{ kJ/kg}\cdot\text{K}$ (Table A-1). The density of river water is 1000 kg/m^3 .

Analysis First we determine the mole fraction of pure water in ocean water using Eqs. 13-4 and 13-5. Noting that $mf_s = 0.025$ and $mf_w = 1 - mf_s = 0.975$,

$$M_m = \frac{1}{\sum \frac{mf_i}{M_i}} = \frac{1}{\frac{mf_s}{M_s} + \frac{mf_w}{M_w}} = \frac{1}{\frac{0.025}{58.44} + \frac{0.975}{18.0}} = 18.32 \text{ kg/kmol}$$

$$y_i = mf_i \frac{M_m}{M_i} \rightarrow y_w = mf_w \frac{M_m}{M_w} = (0.975) \frac{18.32 \text{ kg/kmol}}{18.0 \text{ kg/kmol}} = 0.9922$$

The maximum work output associated with mixing 1 kg of seawater (or the minimum work input required to produce 1 kg of freshwater from seawater) is

$$w_{\max, \text{out}} = R_w T_0 \ln(1/y_w) = (0.4615 \text{ kJ/kg}\cdot\text{K})(288.15 \text{ K})\ln(1/0.9922) = 1.046 \text{ kJ/kg fresh water}$$

Therefore, 1.046 kJ of work can be produced as 1 kg of fresh water is mixed with seawater reversibly. Therefore, the power that can be generated as a river with a flow rate of $400,000 \text{ m}^3/\text{s}$ mixes reversibly with seawater is

$$\dot{W}_{\max \text{ out}} = \rho \dot{V} w_{\max \text{ out}} = (1000 \text{ kg/m}^3)(1.5 \times 10^5 \text{ m}^3/\text{s})(1.046 \text{ kJ/kg}) \left(\frac{1 \text{ kW}}{1 \text{ kJ/s}} \right) = \mathbf{157 \times 10^6 \text{ kW}}$$

Discussion This is more power than produced by all nuclear power plants (112 of them) in the U.S., which shows the tremendous amount of power potential wasted as the rivers discharge into the seas.



13-83 Problem 13-82 is reconsidered. The effect of the salinity of the ocean on the maximum power generated is to be investigated.

Analysis The problem is solved using EES, and the solution is given below.

"Given"

$$V_{\dot{}}=150000 \text{ [m}^3\text{/s]}$$

"salinity=2.5"

$$T=(15+273.15) \text{ [K]}$$

"Properties"

$$M_w=18 \text{ [kg/kmol]} \text{ "molar mass (H}_2\text{O)"}$$

$$M_s=58.44 \text{ [kg/kmol]} \text{ "molar mass of salt"}$$

$$R_w=0.4615 \text{ [kJ/kg-K]} \text{ "gas constant of water"}$$

$$\rho=1000 \text{ [kg/m}^3\text{]}$$

"Analysis"

$$\text{mass}_w=100-\text{salinity}$$

$$\text{mf}_s=\text{salinity}/100$$

$$\text{mf}_w=\text{mass}_w/100$$

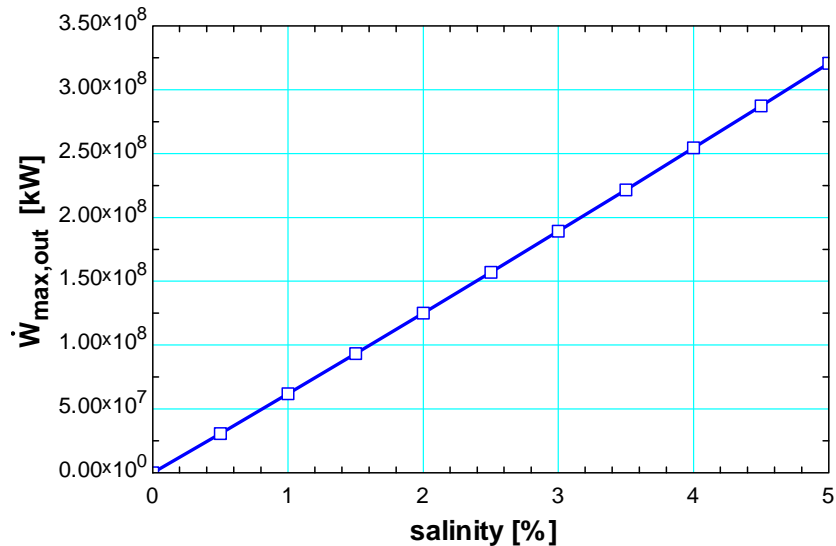
$$M_m=1/(\text{mf}_s/M_s+\text{mf}_w/M_w)$$

$$y_w=\text{mf}_w*M_m/M_w$$

$$w_{\text{max,out}}=R_w*T*\ln(1/y_w)$$

$$W_{\dot{\text{max,out}}}=\rho*V_{\dot{}}*w_{\text{max,out}}$$

Salinity [%]	$W_{\text{max,out}}$ [kW]
0	0
0.5	3.085E+07
1	6.196E+07
1.5	9.334E+07
2	1.249E+08
2.5	1.569E+08
3	1.891E+08
3.5	2.216E+08
4	2.544E+08
4.5	2.874E+08
5	3.208E+08



13-84E Brackish water is used to produce fresh water. The mole fractions, the minimum work inputs required to separate 1 lbm of brackish water and to obtain 1 lbm of fresh water are to be determined.

Assumptions 1 The brackish water is an ideal solution since it is dilute. **2** The total dissolved solids in water can be treated as table salt (NaCl). **3** The environment temperature is equal to the water temperature.

Properties The molar masses of water and salt are $M_w = 18.0$ lbm/lbmol and $M_s = 58.44$ lbm/lbmol. The gas constant of pure water is $R_w = 0.1102$ Btu/lbm·R (Table A-1E).

Analysis (a) First we determine the mole fraction of pure water in brackish water using Eqs. 13-4 and 13-5. Noting that $mf_s = 0.0012$ and $mf_w = 1 - mf_s = 0.9988$,

$$M_m = \frac{1}{\sum \frac{mf_i}{M_i}} = \frac{1}{\frac{mf_s}{M_s} + \frac{mf_w}{M_w}} = \frac{1}{\frac{0.0012}{58.44} + \frac{0.9988}{18.0}} = 18.015 \text{ lbm/lbmol}$$

$$y_i = mf_i \frac{M_m}{M_i} \rightarrow y_w = mf_w \frac{M_m}{M_w} = (0.9988) \frac{18.015 \text{ lbm/lbmol}}{18.0 \text{ lbm/lbmol}} = \mathbf{0.99963}$$

$$y_s = 1 - y_w = 1 - 0.99963 = \mathbf{0.00037}$$

(b) The minimum work input required to separate 1 lbmol of brackish water is

$$\begin{aligned} w_{\min, \text{in}} &= -R_w T_0 (y_w \ln y_w + y_s \ln y_s) \\ &= -(0.1102 \text{ Btu/lbmol}\cdot\text{R})(525 \text{ R})[0.99963 \ln(0.99963) + 0.00037 \ln(0.00037)] \\ &= \mathbf{-0.191 \text{ Btu/lbm brackish water}} \end{aligned}$$

(c) The minimum work input required to produce 1 lbm of freshwater from brackish water is

$$w_{\min, \text{in}} = R_w T_0 \ln(1/y_w) = (0.1102 \text{ Btu/lbm}\cdot\text{R})(525 \text{ R}) \ln(1/0.99963) = \mathbf{0.0214 \text{ Btu/lbm fresh water}}$$

Discussion Note that it takes about 9 times work to separate 1 lbm of brackish water into pure water and salt compared to producing 1 lbm of fresh water from a large body of brackish water.

13-85 A desalination plant produces fresh water from seawater. The second law efficiency of the plant is to be determined.

Assumptions 1 The seawater is an ideal solution since it is dilute. **2** The total dissolved solids in water can be treated as table salt (NaCl). **3** The environment temperature is equal to the seawater temperature.

Properties The molar masses of water and salt are $M_w = 18.0 \text{ kg/kmol}$ and $M_s = 58.44 \text{ kg/kmol}$. The gas constant of pure water is $R_w = 0.4615 \text{ kJ/kg}\cdot\text{K}$ (Table A-1). The density of river water is 1000 kg/m^3 .

Analysis First we determine the mole fraction of pure water in seawater using Eqs. 13-4 and 13-5. Noting that $mf_s = 0.032$ and $mf_w = 1 - mf_s = 0.968$,

$$M_m = \frac{1}{\sum \frac{mf_i}{M_i}} = \frac{1}{\frac{mf_s}{M_s} + \frac{mf_w}{M_w}} = \frac{1}{\frac{0.032}{58.44} + \frac{0.968}{18.0}} = 18.41 \text{ kg/kmol}$$

$$y_i = mf_i \frac{M_m}{M_i} \rightarrow y_w = mf_w \frac{M_m}{M_w} = (0.968) \frac{18.41 \text{ kg/kmol}}{18.0 \text{ kg/kmol}} = 0.9900$$

The maximum work output associated with mixing 1 kg of seawater (or the minimum work input required to produce 1 kg of freshwater from seawater) is

$$w_{\max, \text{out}} = R_w T_0 \ln(1/y_w) = (0.4615 \text{ kJ/kg}\cdot\text{K})(283.15 \text{ K})\ln(1/0.990) = 1.313 \text{ kJ/kg fresh water}$$

The power that can be generated as $1.4 \text{ m}^3/\text{s}$ fresh water mixes reversibly with seawater is

$$\dot{W}_{\max \text{ out}} = \rho \dot{V} w_{\max \text{ out}} = (1000 \text{ kg/m}^3)(1.4 \text{ m}^3/\text{s})(1.313 \text{ kJ/kg}) \left(\frac{1 \text{ kW}}{1 \text{ kJ/s}} \right) = 1.84 \text{ kW}$$

Then the second law efficiency of the plant becomes

$$\eta_{\text{II}} = \frac{\dot{W}_{\min, \text{in}}}{\dot{W}_{\text{in}}} = \frac{1.83 \text{ MW}}{8.5 \text{ MW}} = 0.216 = \mathbf{21.6\%}$$

13-86 The power consumption and the second law efficiency of a desalination plant are given. The power that can be produced if the fresh water produced is mixed with the seawater reversibly is to be determined.

Assumptions 1 This is a steady-flow process. **2** The kinetic and potential energy changes are negligible.

Analysis From the definition of the second law efficiency

$$\eta_{\text{II}} = \frac{\dot{W}_{\text{rev}}}{\dot{W}_{\text{actual}}} \rightarrow 0.25 = \frac{\dot{W}_{\text{rev}}}{11,500 \text{ kW}} \rightarrow \dot{W}_{\text{rev}} = \mathbf{2875 \text{ kW}}$$

which is the maximum power that can be generated.

13-87E It is to be determined if it is possible for an adiabatic liquid-vapor separator to separate wet steam at 100 psia and 90 percent quality, so that the pressure of the outlet streams is greater than 100 psia.

Analysis Because the separator divides the inlet stream into the liquid and vapor portions,

$$\dot{m}_2 = x\dot{m}_1 = 0.9\dot{m}_1$$

$$\dot{m}_3 = (1-x)\dot{m}_1 = 0.1\dot{m}_1$$

According to the water property tables at 100 psia (Table A-5E),

$$s_1 = s_f + xs_{fg} = 0.47427 + 0.9 \times 1.12888 = 1.4903 \text{ Btu/lbm} \cdot \text{R}$$

When the increase in entropy principle is adapted to this system, it becomes

$$\dot{m}_2 s_2 + \dot{m}_3 s_3 \geq \dot{m}_1 s_1$$

$$x\dot{m}_1 s_2 + (1-x)\dot{m}_1 s_3 \geq \dot{m}_1 s_1$$

$$0.9s_2 + 0.1s_3 \geq s_1$$

$$\geq 1.4903 \text{ Btu/lbm} \cdot \text{R}$$

To test this hypothesis, let's assume the outlet pressures are 110 psia. Then,

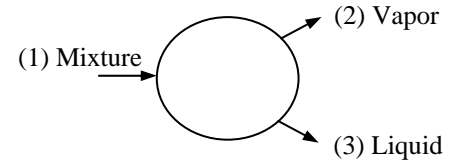
$$s_2 = s_g = 1.5954 \text{ Btu/lbm} \cdot \text{R}$$

$$s_3 = s_f = 0.48341 \text{ Btu/lbm} \cdot \text{R}$$

The left-hand side of the above equation is

$$0.9s_2 + 0.1s_3 = 0.9 \times 1.5954 + 0.1 \times 0.48341 = 1.4842 \text{ Btu/lbm} \cdot \text{R}$$

which is less than the minimum possible specific entropy. Hence, the outlet pressure cannot be 110 psia. Inspection of the water table in light of above equation proves that the pressure at the separator outlet cannot be greater than that at the inlet.



Review Problems

13-88 Using Dalton's law, it is to be shown that $Z_m = \sum_{i=1}^k y_i Z_i$ for a real-gas mixture.

Analysis Using the compressibility factor, the pressure of a component of a real-gas mixture and of the pressure of the gas mixture can be expressed as

$$P_i = \frac{Z_i N_i R_u T_m}{\nu_m} \quad \text{and} \quad P_m = \frac{Z_m N_m R_u T_m}{\nu_m}$$

Dalton's law can be expressed as $P_m = \sum P_i(T_m, \nu_m)$. Substituting,

$$\frac{Z_m N_m R_u T_m}{\nu_m} = \sum \frac{Z_i N_i R_u T_m}{\nu_m}$$

Simplifying,

$$Z_m N_m = \sum Z_i N_i$$

Dividing by N_m ,

$$Z_m = \sum y_i Z_i$$

where Z_i is determined at the mixture temperature and volume.

13-89 The volume fractions of components of a gas mixture are given. The mole fractions, the mass fractions, the partial pressures, the mixture molar mass, apparent gas constant, and constant-pressure specific heat are to be determined and compared to the values in Table A-2a.

Properties The molar masses of N_2 , O_2 and Ar are 28.0, 32.0, and 40.0 kg/kmol, respectively (Table A-1). The constant-pressure specific heats of these gases at 300 K are 1.039, 0.918, and 0.5203 kJ/kg·K, respectively (Table A-2a).

Analysis The volume fractions are equal to the mole fractions:

$$y_{N_2} = \mathbf{0.78}, \quad y_{O_2} = \mathbf{0.21}, \quad y_{Ar} = \mathbf{0.01}$$

The volume fractions are equal to the pressure fractions. The partial pressures are then

$$P_{N_2} = y_{N_2} P_{\text{total}} = (0.78)(100 \text{ kPa}) = \mathbf{78 \text{ kPa}}$$

$$P_{O_2} = y_{O_2} P_{\text{total}} = (0.21)(100 \text{ kPa}) = \mathbf{21 \text{ kPa}}$$

$$P_{Ar} = y_{Ar} P_{\text{total}} = (0.01)(100 \text{ kPa}) = \mathbf{1 \text{ kPa}}$$

<p>78% N_2 21% O_2 1% Ar (by volume)</p>
--

We consider 100 kmol of this mixture. Noting that volume fractions are equal to the mole fractions, mass of each component are

$$m_{N_2} = N_{N_2} M_{N_2} = (78 \text{ kmol})(28 \text{ kg/kmol}) = 2184 \text{ kg}$$

$$m_{O_2} = N_{O_2} M_{O_2} = (21 \text{ kmol})(32 \text{ kg/kmol}) = 672 \text{ kg}$$

$$m_{Ar} = N_{Ar} M_{Ar} = (1 \text{ kmol})(40 \text{ kg/kmol}) = 40 \text{ kg}$$

The total mass is

$$m_m = m_{N_2} + m_{O_2} + m_{Ar} = 2184 + 672 + 40 = 2896 \text{ kg}$$

Then the mass fractions are

$$\text{mf}_{N_2} = \frac{m_{N_2}}{m_m} = \frac{2184 \text{ kg}}{2896 \text{ kg}} = \mathbf{0.7541}$$

$$\text{mf}_{O_2} = \frac{m_{O_2}}{m_m} = \frac{672 \text{ kg}}{2896 \text{ kg}} = \mathbf{0.2320}$$

$$\text{mf}_{Ar} = \frac{m_{Ar}}{m_m} = \frac{40 \text{ kg}}{2896 \text{ kg}} = \mathbf{0.0138}$$

The apparent molecular weight of the mixture is

$$M_m = \frac{m_m}{N_m} = \frac{2896 \text{ kg}}{100 \text{ kmol}} = \mathbf{28.96 \text{ kg/kmol}}$$

The constant-pressure specific heat of the mixture is determined from

$$\begin{aligned} c_p &= \text{mf}_{N_2} c_{p,N_2} + \text{mf}_{O_2} c_{p,O_2} + \text{mf}_{Ar} c_{p,Ar} \\ &= 0.7541 \times 1.039 + 0.2320 \times 0.918 + 0.0138 \times 0.5203 \\ &= \mathbf{1.004 \text{ kJ/kg} \cdot \text{K}} \end{aligned}$$

The apparent gas constant of the mixture is

$$R = \frac{R_u}{M_m} = \frac{8.314 \text{ kJ/kmol} \cdot \text{K}}{28.96 \text{ kg/kmol}} = \mathbf{0.2871 \text{ kJ/kg} \cdot \text{K}}$$

This mixture closely correspond to the air, and the mixture properties determined (mixture molar mass, mixture gas constant and mixture specific heat) are practically the same as those listed for air in Tables A-1 and A-2a.

13-90 The mole numbers of combustion gases are given. The partial pressure of water vapor and the condensation temperature of water vapor are to be determined.

Properties The molar masses of CO_2 , H_2O , O_2 and N_2 are 44.0, 18.0, 32.0, and 28.0 kg/kmol, respectively (Table A-1).

Analysis The total mole of the mixture and the mole fraction of water vapor are

$$N_{\text{total}} = 8 + 9 + 12.5 + 94 = 123.5 \text{ kmol}$$

$$y_{\text{H}_2\text{O}} = \frac{N_{\text{H}_2\text{O}}}{N_{\text{total}}} = \frac{9}{123.5} = 0.07287$$

Noting that molar fraction is equal to pressure fraction, the partial pressure of water vapor is

$$P_{\text{H}_2\text{O}} = y_{\text{H}_2\text{O}} P_{\text{total}} = (0.07287)(100 \text{ kPa}) = \mathbf{7.29 \text{ kPa}}$$

The temperature at which the condensation starts is the saturation temperature of water at this pressure. This is called the dew-point temperature. Then,

$$T_{\text{cond}} = T_{\text{sat}@7.29 \text{ kPa}} = \mathbf{39.7^\circ\text{C}} \quad (\text{Table A-5})$$

Water vapor in the combustion gases will start to condense when the temperature of the combustion gases drop to 39.7°C .

13-91 The masses of gases forming a mixture at a specified pressure and temperature are given. The mass of the gas mixture is to be determined using four methods.

Properties The molar masses of O₂, CO₂, and He are 32.0, 44.0, and 4.0 kg/kmol, respectively (Table A-1).

Analysis (a) The given total mass of the mixture is

$$m_m = m_{\text{O}_2} + m_{\text{CO}_2} + m_{\text{He}} = 0.1 + 1 + 0.5 = 1.6 \text{ kg}$$

The mole numbers of each component are

$$N_{\text{O}_2} = \frac{m_{\text{O}_2}}{M_{\text{O}_2}} = \frac{0.1 \text{ kg}}{32 \text{ kg/kmol}} = 0.003125 \text{ kmol}$$

$$N_{\text{CO}_2} = \frac{m_{\text{CO}_2}}{M_{\text{CO}_2}} = \frac{1 \text{ kg}}{44 \text{ kg/kmol}} = 0.02273 \text{ kmol}$$

$$N_{\text{He}} = \frac{m_{\text{He}}}{M_{\text{He}}} = \frac{0.5 \text{ kg}}{4 \text{ kg/kmol}} = 0.125 \text{ kmol}$$

0.1 kg O₂
1 kg CO₂
0.5 kg He

The mole number of the mixture is

$$N_m = N_{\text{O}_2} + N_{\text{CO}_2} + N_{\text{He}} = 0.003125 + 0.02273 + 0.125 = 0.1509 \text{ kmol}$$

Then the apparent molecular weight of the mixture becomes

$$M_m = \frac{m_m}{N_m} = \frac{1.6 \text{ kg}}{0.1509 \text{ kmol}} = 10.61 \text{ kg/kmol}$$

The mass of this mixture in a 0.3 m³ tank is

$$m = \frac{M_m P V}{R_u T} = \frac{(10.61 \text{ kg/kmol})(17,500 \text{ kPa})(0.3 \text{ m}^3)}{(8.314 \text{ kPa} \cdot \text{m}^3/\text{kmol} \cdot \text{K})(293 \text{ K})} = \mathbf{22.87 \text{ kg}}$$

(b) To use the Amagat's law for this real gas mixture, we first need the mole fractions and the Z of each component at the mixture temperature and pressure.

$$y_{\text{O}_2} = \frac{N_{\text{O}_2}}{N_m} = \frac{0.003125 \text{ kmol}}{0.1509 \text{ kmol}} = 0.02071$$

$$y_{\text{CO}_2} = \frac{N_{\text{CO}_2}}{N_m} = \frac{0.02273 \text{ kmol}}{0.1509 \text{ kmol}} = 0.1506$$

$$y_{\text{He}} = \frac{N_{\text{He}}}{N_m} = \frac{0.125 \text{ kmol}}{0.1509 \text{ kmol}} = 0.8284$$

$$\left. \begin{aligned} T_{R,\text{O}_2} &= \frac{T_m}{T_{\text{cr},\text{O}_2}} = \frac{293 \text{ K}}{154.8 \text{ K}} = 1.893 \\ P_{R,\text{O}_2} &= \frac{P_m}{P_{\text{cr},\text{O}_2}} = \frac{17.5 \text{ MPa}}{5.08 \text{ MPa}} = 3.445 \end{aligned} \right\} Z_{\text{O}_2} = 0.93 \quad (\text{Fig. A-15})$$

$$\left. \begin{aligned} T_{R,\text{CO}_2} &= \frac{T_m}{T_{\text{cr},\text{CO}_2}} = \frac{293 \text{ K}}{304.2 \text{ K}} = 0.963 \\ P_{R,\text{CO}_2} &= \frac{P_m}{P_{\text{cr},\text{CO}_2}} = \frac{17.5 \text{ MPa}}{7.39 \text{ MPa}} = 2.368 \end{aligned} \right\} Z_{\text{CO}_2} = 0.33 \quad (\text{Fig. A-15})$$

$$\left. \begin{aligned} T_{R,\text{He}} &= \frac{T_m}{T_{\text{cr},\text{He}}} = \frac{293 \text{ K}}{5.3 \text{ K}} = 55.3 \\ P_{R,\text{He}} &= \frac{P_m}{P_{\text{cr},\text{He}}} = \frac{17.5 \text{ MPa}}{0.23 \text{ MPa}} = 76.1 \end{aligned} \right\} Z_{\text{He}} = 1.04 \quad (\text{from EES})$$

Then,

$$Z_m = \sum y_i Z_i = y_{O_2} Z_{O_2} + y_{CO_2} Z_{CO_2} + y_{He} Z_{He}$$

$$= (0.02071)(0.93) + (0.1506)(0.33) + (0.8284)(1.04) = 0.9305$$

$$m = \frac{M_m P V}{Z_m R_u T} = \frac{(10.61 \text{ kg/kmol})(17,500 \text{ kPa})(0.3 \text{ m}^3)}{(0.9305)(8.314 \text{ kPa} \cdot \text{m}^3/\text{kmol} \cdot \text{K})(293 \text{ K})} = \mathbf{24.57 \text{ kg}}$$

(c) To use Dalton's law with compressibility factors:

$$\left. \begin{aligned} T_{R,O_2} &= 1.893 \\ \nu_{R,O_2} &= \frac{\nu_m/m_{O_2}}{R_{O_2} T_{cr,O_2} / P_{cr,O_2}} = \frac{(0.3 \text{ m}^3)/(22.87 \times 0.1/1.6 \text{ kg})}{(0.2598 \text{ kPa} \cdot \text{m}^3/\text{kg} \cdot \text{K})(154.8 \text{ K})/(5080 \text{ kPa})} = 26.5 \end{aligned} \right\} Z_{O_2} = 1.0$$

$$\left. \begin{aligned} T_{R,CO_2} &= 0.963 \\ \nu_{R,CO_2} &= \frac{\nu_m/m_{CO_2}}{R_{CO_2} T_{cr,CO_2} / P_{cr,CO_2}} = \frac{(0.3 \text{ m}^3)/(22.87 \times 1.0/1.6 \text{ kg})}{(0.1889 \text{ kPa} \cdot \text{m}^3/\text{kg} \cdot \text{K})(304.2 \text{ K})/(7390 \text{ kPa})} = 2.70 \end{aligned} \right\} Z_{CO_2} = 0.86$$

$$\left. \begin{aligned} T_{R,He} &= 55.3 \\ \nu_{R,He} &= \frac{\nu_m/m_{He}}{R_{He} T_{cr,He} / P_{cr,He}} = \frac{(0.3 \text{ m}^3)/(22.87 \times 0.5/1.6 \text{ kg})}{(2.0769 \text{ kPa} \cdot \text{m}^3/\text{kg} \cdot \text{K})(5.3 \text{ K})/(230 \text{ kPa})} = 0.88 \end{aligned} \right\} Z_{He} = 1.0$$

Note that we used $m = 22.87 \text{ kg}$ in above calculations, the value obtained by ideal gas behavior. The solution normally requires iteration until the assumed and calculated mass values match. The mass of the component gas is obtained by multiplying the mass of the mixture by its mass fraction. Then,

$$Z_m = \sum y_i Z_i = y_{O_2} Z_{O_2} + y_{CO_2} Z_{CO_2} + y_{He} Z_{He}$$

$$= (0.02071)(1.0) + (0.1506)(0.86) + (0.8284)(1.0) = 0.9786$$

$$m = \frac{M_m P V}{Z_m R_u T} = \frac{(10.61 \text{ kg/kmol})(17,500 \text{ kPa})(0.3 \text{ m}^3)}{(0.9786)(8.314 \text{ kPa} \cdot \text{m}^3/\text{kmol} \cdot \text{K})(293 \text{ K})} = \mathbf{23.37 \text{ kg}}$$

This mass is sufficiently close to the mass value 22.87 kg. Therefore, there is no need to repeat the calculations at this calculated mass.

(d) To use Kay's rule, we need to determine the pseudo-critical temperature and pseudo-critical pressure of the mixture using the critical point properties of O_2 , CO_2 and He.

$$T'_{cr,m} = \sum y_i T_{cr,i} = y_{O_2} T_{cr,O_2} + y_{CO_2} T_{cr,CO_2} + y_{He} T_{cr,He}$$

$$= (0.02071)(154.8 \text{ K}) + (0.1506)(304.2 \text{ K}) + (0.8284)(5.3 \text{ K}) = 53.41 \text{ K}$$

$$P'_{cr,m} = \sum y_i P_{cr,i} = y_{O_2} P_{cr,O_2} + y_{CO_2} P_{cr,CO_2} + y_{He} P_{cr,He}$$

$$= (0.02071)(5.08 \text{ MPa}) + (0.1506)(7.39 \text{ MPa}) + (0.8284)(0.23 \text{ MPa}) = 1.409 \text{ MPa}$$

Then,

$$\left. \begin{aligned} T_R &= \frac{T_m}{T'_{cr,m}} = \frac{293 \text{ K}}{53.41 \text{ K}} = 5.486 \\ P_R &= \frac{P_m}{P'_{cr,m}} = \frac{17.5 \text{ MPa}}{1.409 \text{ MPa}} = 12.42 \end{aligned} \right\} Z_m = 1.194 \quad (\text{from EES})$$

$$m = \frac{M_m P V}{Z_m R_u T} = \frac{(10.61 \text{ kg/kmol})(17,500 \text{ kPa})(0.3 \text{ m}^3)}{(1.194)(8.314 \text{ kPa} \cdot \text{m}^3/\text{kmol} \cdot \text{K})(293 \text{ K})} = \mathbf{19.15 \text{ kg}}$$

13-92 A mixture of carbon dioxide and nitrogen flows through a converging nozzle. The required make up of the mixture on a mass basis is to be determined.

Assumptions Under specified conditions CO_2 and N_2 can be treated as ideal gases, and the mixture as an ideal gas mixture.

Properties The molar masses of CO_2 and N_2 are 44.0 and 28.0 kg/kmol, respectively (Table A-1). The specific heat ratios of CO_2 and N_2 at 500 K are $k_{\text{CO}_2} = 1.229$ and $k_{\text{N}_2} = 1.391$ (Table A-2).

Analysis The molar mass of the mixture is determined from

$$M_m = y_{\text{CO}_2} M_{\text{CO}_2} + y_{\text{N}_2} M_{\text{N}_2}$$

The molar fractions are related to each other by

$$y_{\text{CO}_2} + y_{\text{N}_2} = 1$$

The gas constant of the mixture is given by

$$R_m = \frac{R_u}{M_m}$$

The specific heat ratio of the mixture is expressed as

$$k = \text{mf}_{\text{CO}_2} k_{\text{CO}_2} + \text{mf}_{\text{N}_2} k_{\text{N}_2}$$

The mass fractions are

$$\text{mf}_{\text{CO}_2} = y_{\text{CO}_2} \frac{M_{\text{CO}_2}}{M_m}$$

$$\text{mf}_{\text{N}_2} = y_{\text{N}_2} \frac{M_{\text{N}_2}}{M_m}$$

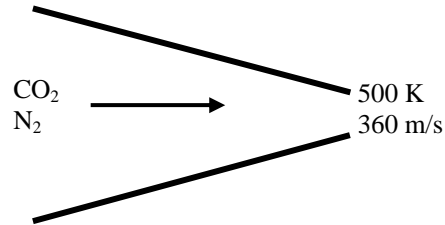
The exit velocity equals the speed of sound at 500 K

$$V_{\text{exit}} = \sqrt{k R_m T \left(\frac{1000 \text{ m}^2/\text{s}^2}{1 \text{ kJ/kg}} \right)}$$

Substituting the given values and known properties and solving the above equations simultaneously using EES, we find

$$\text{mf}_{\text{CO}_2} = \mathbf{0.838}$$

$$\text{mf}_{\text{N}_2} = \mathbf{0.162}$$



13-93E A mixture of nitrogen and oxygen is expanded isothermally. The work produced is to be determined.

Assumptions 1 Nitrogen and oxygen are ideal gases. 2 The process is reversible.

Properties The mole numbers of nitrogen and oxygen are 28.0 and 32.0 lbm/lbmol, respectively (Table A-1E).

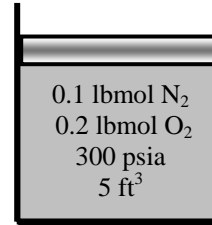
Analysis The mole fractions are

$$y_{\text{N}_2} = \frac{N_{\text{N}_2}}{N_{\text{total}}} = \frac{0.1 \text{ lbmol}}{0.3 \text{ lbmol}} = 0.3333$$

$$y_{\text{O}_2} = \frac{N_{\text{O}_2}}{N_{\text{total}}} = \frac{0.2 \text{ kmol}}{0.3 \text{ kmol}} = 0.6667$$

The gas constant for this mixture is then

$$\begin{aligned} R &= \frac{R_u}{y_{\text{N}_2} M_{\text{N}_2} + y_{\text{O}_2} M_{\text{O}_2}} \\ &= \frac{1.9858 \text{ Btu/lbmol} \cdot \text{R}}{(0.3333 \times 28 + 0.6667 \times 32) \text{ lbm/lbmol}} \\ &= 0.06475 \text{ Btu/lbm} \cdot \text{R} \\ &= (0.06475 \text{ Btu/lbm} \cdot \text{R}) \left(\frac{5.404 \text{ psia} \cdot \text{ft}^3}{1 \text{ Btu}} \right) \\ &= 0.3499 \text{ psia} \cdot \text{ft}^3 / \text{lbm} \cdot \text{R} \end{aligned}$$



The mass of this mixture of gases is

$$m = N_{\text{N}_2} M_{\text{N}_2} + N_{\text{O}_2} M_{\text{O}_2} = 0.1 \times 28 + 0.2 \times 32 = 9.2 \text{ lbm}$$

The temperature of the mixture is

$$T_1 = \frac{P_1 V_1}{mR} = \frac{(300 \text{ psia})(5 \text{ ft}^3)}{(9.2 \text{ lbm})(0.3499 \text{ psia} \cdot \text{ft}^3 / \text{lbm} \cdot \text{R})} = 466.0 \text{ R}$$

Noting that $Pv = RT$ for an ideal gas, the work done for this process is then

$$\begin{aligned} W_{\text{out}} &= m \int_1^2 P d v = mRT \int_1^2 \frac{d v}{v} = mRT \ln \frac{v_2}{v_1} \\ &= (9.2 \text{ lbm})(0.06475 \text{ Btu/lbm} \cdot \text{R})(466 \text{ R}) \ln \frac{10 \text{ ft}^3}{5 \text{ ft}^3} \\ &= \mathbf{192.4 \text{ Btu}} \end{aligned}$$

13-94 A mixture of nitrogen and carbon dioxide is compressed at constant temperature in a closed system. The work required is to be determined.

Assumptions 1 Nitrogen and carbon dioxide are ideal gases. 2 The process is reversible.

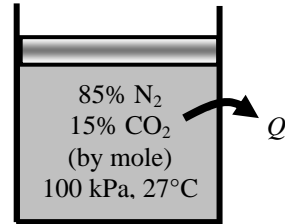
Properties The mole numbers of nitrogen and carbon dioxide are 28.0 and 44.0 kg/kmol, respectively (Table A-1).

Analysis The effective molecular weight of this mixture is

$$\begin{aligned} M &= y_{\text{N}_2}M_{\text{N}_2} + y_{\text{CO}_2}M_{\text{CO}_2} \\ &= (0.85)(28) + (0.15)(44) \\ &= 30.4 \text{ kg/kmol} \end{aligned}$$

The work done is determined from

$$\begin{aligned} w &= \int_1^2 P d\nu = RT \int_1^2 \frac{d\nu}{\nu} = RT \ln \frac{\nu_2}{\nu_1} = RT \ln \frac{P_2}{P_1} = \frac{R_u}{M} RT \ln \frac{P_2}{P_1} \\ &= \frac{8.314 \text{ kJ/kmol} \cdot \text{K}}{30.4 \text{ kg/kmol}} (300 \text{ K}) \ln \frac{500 \text{ kPa}}{100 \text{ kPa}} \\ &= \mathbf{132.0 \text{ kJ/kg}} \end{aligned}$$



13-95 The specific heat ratio and an apparent molecular weight of a mixture of ideal gases are given. The work required to compress this mixture isentropically in a closed system is to be determined.

Analysis For an isentropic process of an ideal gas with constant specific heats, the work is expressed as

$$\begin{aligned} w_{\text{out}} &= \int_1^2 P d\nu = P_1 \nu_1^k \int_1^2 \nu^{-k} d\nu \\ &= \frac{P_1 \nu_1^k}{1-k} (\nu_2^{1-k} - \nu_1^{1-k}) = \frac{P_1 \nu_1^k}{1-k} \left[\left(\frac{\nu_2}{\nu_1} \right)^{1-k} - 1 \right] \end{aligned}$$

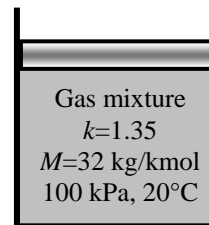
since $P_1 \nu_1^k = P \nu^k$ for an isentropic process. Also,

$$\begin{aligned} P_1 \nu_1 &= RT_1 \\ (\nu_2 / \nu_1)^k &= P_1 / P_2 \end{aligned}$$

Substituting, we obtain

$$\begin{aligned} w_{\text{out}} &= \frac{R_u T_1}{M(1-k)} \left[\left(\frac{P_2}{P_1} \right)^{(k-1)/k} - 1 \right] \\ &= \frac{(8.314 \text{ kJ/kmol} \cdot \text{K})(293 \text{ K})}{(32 \text{ kg/kmol})(1-1.35)} \left[\left(\frac{1000 \text{ kPa}}{100 \text{ kPa}} \right)^{(1.35-1)/1.35} - 1 \right] \\ &= \mathbf{-177.6 \text{ kJ/kg}} \end{aligned}$$

The negative sign shows that the work is done on the system.



13-96 A mixture of gases is placed in a spring-loaded piston-cylinder device. The device is now heated until the pressure rises to a specified value. The total work and heat transfer for this process are to be determined.

Properties The molar masses of Ne, O₂, and N₂ are 20.18, 32.0, 28.0 kg/kmol, respectively and the gas constants are 0.4119, 0.2598, and 0.2968 kJ/kg·K, respectively (Table A-1). The constant-volume specific heats are 0.6179, 0.658, and 0.743 kJ/kg·K, respectively (Table A-2a).

Analysis The total pressure is 200 kPa and the partial pressures are

$$P_{\text{Ne}} = y_{\text{Ne}} P_m = (0.25)(200 \text{ kPa}) = 50 \text{ kPa}$$

$$P_{\text{O}_2} = y_{\text{O}_2} P_m = (0.50)(200 \text{ kPa}) = 100 \text{ kPa}$$

$$P_{\text{N}_2} = y_{\text{N}_2} P_m = (0.25)(200 \text{ kPa}) = 50 \text{ kPa}$$

The mass of each constituent for a volume of 0.1 m³ and a temperature of 10°C are

$$m_{\text{Ne}} = \frac{P_{\text{Ne}} V_m}{R_{\text{Ne}} T} = \frac{(50 \text{ kPa})(0.1 \text{ m}^3)}{(0.4119 \text{ kPa} \cdot \text{m}^3/\text{kg} \cdot \text{K})(283 \text{ K})} = 0.04289 \text{ kg}$$

$$m_{\text{O}_2} = \frac{P_{\text{O}_2} V_m}{R_{\text{O}_2} T} = \frac{(100 \text{ kPa})(0.1 \text{ m}^3)}{(0.2598 \text{ kPa} \cdot \text{m}^3/\text{kg} \cdot \text{K})(283 \text{ K})} = 0.1360 \text{ kg}$$

$$m_{\text{N}_2} = \frac{P_{\text{N}_2} V_m}{R_{\text{N}_2} T} = \frac{(50 \text{ kPa})(0.1 \text{ m}^3)}{(0.2968 \text{ kPa} \cdot \text{m}^3/\text{kg} \cdot \text{K})(283 \text{ K})} = 0.05953 \text{ kg}$$

$$m_{\text{total}} = 0.04289 + 0.1360 + 0.05953 = 0.2384 \text{ kg}$$

The mass fractions are

$$\text{mf}_{\text{Ne}} = \frac{m_{\text{Ne}}}{m_m} = \frac{0.04289 \text{ kg}}{0.2384 \text{ kg}} = 0.1799$$

$$\text{mf}_{\text{O}_2} = \frac{m_{\text{O}_2}}{m_m} = \frac{0.1360 \text{ kg}}{0.2384 \text{ kg}} = 0.5705$$

$$\text{mf}_{\text{N}_2} = \frac{m_{\text{N}_2}}{m_m} = \frac{0.05953 \text{ kg}}{0.2384 \text{ kg}} = 0.2497$$

The constant-volume specific heat of the mixture is determined from

$$\begin{aligned} c_v &= \text{mf}_{\text{Ne}} c_{v,\text{Ne}} + \text{mf}_{\text{O}_2} c_{v,\text{O}_2} + \text{mf}_{\text{N}_2} c_{v,\text{N}_2} \\ &= 0.1799 \times 0.6179 + 0.5705 \times 0.658 + 0.2497 \times 0.743 = 0.672 \text{ kJ/kg} \cdot \text{K} \end{aligned}$$

The moles are

$$N_{\text{Ne}} = \frac{m_{\text{Ne}}}{M_{\text{Ne}}} = \frac{0.04289 \text{ kg}}{20.18 \text{ kg/kmol}} = 0.002126 \text{ kmol}$$

$$N_{\text{O}_2} = \frac{m_{\text{O}_2}}{M_{\text{O}_2}} = \frac{0.1360 \text{ kg}}{32 \text{ kg/kmol}} = 0.00425 \text{ kmol}$$

$$N_{\text{N}_2} = \frac{m_{\text{N}_2}}{M_{\text{N}_2}} = \frac{0.05953 \text{ kg}}{28 \text{ kg/kmol}} = 0.002126 \text{ kmol}$$

$$N_m = N_{\text{Ne}} + N_{\text{O}_2} + N_{\text{N}_2} = 0.008502 \text{ kmol}$$

Then the apparent molecular weight of the mixture becomes

$$M_m = \frac{m_m}{N_m} = \frac{0.2384 \text{ kg}}{0.008502 \text{ kmol}} = 28.04 \text{ kg/kmol}$$

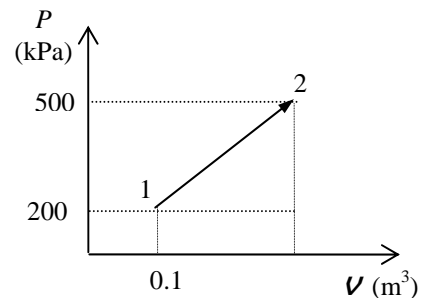
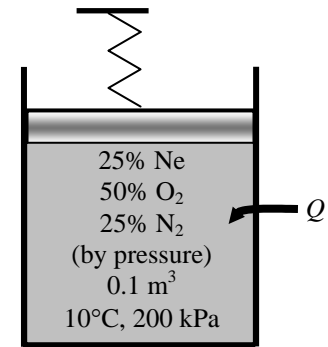
The apparent gas constant of the mixture is

$$R = \frac{R_u}{M_m} = \frac{8.314 \text{ kJ/kmol} \cdot \text{K}}{28.05 \text{ kg/kmol}} = 0.2964 \text{ kJ/kg} \cdot \text{K}$$

The mass contained in the system is

$$m = \frac{P_1 V_1}{RT_1} = \frac{(200 \text{ kPa})(0.1 \text{ m}^3)}{(0.2964 \text{ kPa} \cdot \text{m}^3/\text{kg} \cdot \text{K})(283 \text{ K})} = 0.2384 \text{ kg}$$

Noting that the pressure changes linearly with volume, the final volume is determined by linear interpolation to be



$$\frac{500 - 200}{1000 - 200} = \frac{V_2 - 0.1}{1.0 - 0.1} \longrightarrow V_2 = 0.4375 \text{ m}^3$$

The final temperature is

$$T_2 = \frac{P_2 V_2}{mR} = \frac{(500 \text{ kPa})(0.4375 \text{ m}^3)}{(0.2384 \text{ kg})(0.2964 \text{ kPa} \cdot \text{m}^3/\text{kg} \cdot \text{K})} = 3096 \text{ K}$$

The work done during this process is

$$W_{\text{out}} = \frac{P_1 + P_2}{2} (V_2 - V_1) = \frac{(500 + 200) \text{ kPa}}{2} (0.4375 - 0.1) \text{ m}^3 = \mathbf{118 \text{ kJ}}$$

An energy balance on the system gives

$$Q_{\text{in}} = W_{\text{out}} + mc_v (T_2 - T_1) = 118 + (0.2384 \text{ kg})(0.672 \text{ kJ/kg} \cdot \text{K})(3096 - 283) \text{ K} = \mathbf{569 \text{ kJ}}$$

13-97 A spring-loaded piston-cylinder device is filled with a mixture of nitrogen and carbon dioxide whose mass fractions are given. The gas is heated until the volume has doubled. The total work and heat transfer for this process are to be determined.

Properties The molar masses of N_2 and CO_2 are 28.0 and 44.0 kg/kmol, respectively (Table A-1). The constant-volume specific heats of these gases at room temperature are 0.743 and 0.657 kJ/kg·K, respectively (Table A-2a).

Analysis We consider 100 kg of this mixture. The mole numbers of each component are

$$N_{N_2} = \frac{m_{N_2}}{M_{N_2}} = \frac{55 \text{ kg}}{28 \text{ kg/kmol}} = 1.964 \text{ kmol}$$

$$N_{CO_2} = \frac{m_{CO_2}}{M_{CO_2}} = \frac{45 \text{ kg}}{44 \text{ kg/kmol}} = 1.023 \text{ kmol}$$

The mole number of the mixture is

$$N_m = N_{N_2} + N_{CO_2} = 1.964 + 1.023 = 2.987 \text{ kmol}$$

The apparent molecular weight of the mixture is

$$M_m = \frac{m_m}{N_m} = \frac{100 \text{ kg}}{2.987 \text{ kmol}} = 33.48 \text{ kg/kmol}$$

The constant-volume specific heat of the mixture is determined from

$$c_v = mf_{N_2} c_{v,N_2} + mf_{CO_2} c_{v,CO_2} = 0.55 \times 0.743 + 0.45 \times 0.657 = 0.7043 \text{ kJ/kg} \cdot \text{K}$$

The apparent gas constant of the mixture is

$$R = \frac{R_u}{M_m} = \frac{8.134 \text{ kJ/kmol} \cdot \text{K}}{33.48 \text{ kg/kmol}} = 0.2483 \text{ kJ/kg} \cdot \text{K}$$

Noting that the pressure changes linearly with volume, the initial volume is determined by linear interpolation using the data of the previous problem to be

$$\frac{200 - 200}{1000 - 200} = \frac{V_1 - 0.1}{1.0 - 0.1} \longrightarrow V_1 = 0.1 \text{ m}^3$$

The final volume is

$$V_2 = 2V_1 = 2(0.1 \text{ m}^3) = 0.2 \text{ m}^3$$

The final pressure is similarly determined by linear interpolation using the data of the previous problem to be

$$\frac{P_2 - 200}{1000 - 200} = \frac{0.2 - 0.1}{1.0 - 0.1} \longrightarrow P_2 = 288.9 \text{ kPa}$$

The mass contained in the system is

$$m = \frac{P_1 V_1}{RT_1} = \frac{(200 \text{ kPa})(0.1 \text{ m}^3)}{(0.2483 \text{ kPa} \cdot \text{m}^3/\text{kg} \cdot \text{K})(318 \text{ K})} = 0.2533 \text{ kg}$$

The final temperature is

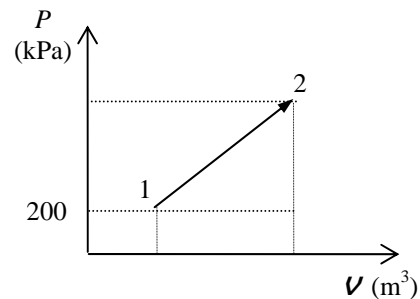
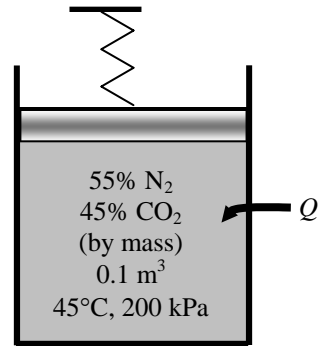
$$T_2 = \frac{P_2 V_2}{mR} = \frac{(288.9 \text{ kPa})(0.2 \text{ m}^3)}{(0.2533 \text{ kg})(0.2483 \text{ kPa} \cdot \text{m}^3/\text{kg} \cdot \text{K})} = 918.7 \text{ K}$$

The work done during this process is

$$W_{\text{out}} = \frac{P_1 + P_2}{2} (V_2 - V_1) = \frac{(200 + 288.9) \text{ kPa}}{2} (0.2 - 0.1) \text{ m}^3 = \mathbf{24.4 \text{ kJ}}$$

An energy balance on the system gives

$$Q_{\text{in}} = W_{\text{out}} + mc_v(T_2 - T_1) = 24.4 + (0.2533 \text{ kg})(0.7043 \text{ kJ/kg} \cdot \text{K})(918.7 - 318) \text{ K} = \mathbf{132 \text{ kJ}}$$



13-98 A spring-loaded piston-cylinder device is filled with a mixture of nitrogen and carbon dioxide whose mass fractions are given. The gas is heated until the pressure has tripled. The total work and heat transfer for this process are to be determined.

Properties The molar masses of N_2 and CO_2 are 28.0 and 44.0 kg/kmol, respectively (Table A-1). The constant-volume specific heats of these gases at room temperature are 0.743 and 0.657 kJ/kg·K, respectively (Table A-2a).

Analysis We consider 100 kg of this mixture. The mole numbers of each component are

$$N_{N_2} = \frac{m_{N_2}}{M_{N_2}} = \frac{55 \text{ kg}}{28 \text{ kg/kmol}} = 1.964 \text{ kmol}$$

$$N_{CO_2} = \frac{m_{CO_2}}{M_{CO_2}} = \frac{45 \text{ kg}}{44 \text{ kg/kmol}} = 1.023 \text{ kmol}$$

The mole number of the mixture is

$$N_m = N_{N_2} + N_{CO_2} = 1.964 + 1.023 = 2.987 \text{ kmol}$$

The apparent molecular weight of the mixture is

$$M_m = \frac{m_m}{N_m} = \frac{100 \text{ kg}}{2.987 \text{ kmol}} = 33.48 \text{ kg/kmol}$$

The constant-volume specific heat of the mixture is determined from

$$c_v = mf_{N_2}c_{v,N_2} + mf_{CO_2}c_{v,CO_2} = 0.55 \times 0.743 + 0.45 \times 0.657 = 0.7043 \text{ kJ/kg} \cdot \text{K}$$

The apparent gas constant of the mixture is

$$R = \frac{R_u}{M_m} = \frac{8.134 \text{ kJ/kmol} \cdot \text{K}}{33.48 \text{ kg/kmol}} = 0.2483 \text{ kJ/kg} \cdot \text{K}$$

Noting that the pressure changes linearly with volume, the initial volume is determined by linear interpolation using the data of the previous problem to be

$$\frac{200 - 200}{1000 - 200} = \frac{V_1 - 0.1}{1.0 - 0.1} \longrightarrow V_1 = 0.1 \text{ m}^3$$

The final pressure is

$$P_2 = 3P_1 = 3(200 \text{ kPa}) = 600 \text{ kPa}$$

The final volume is similarly determined by linear interpolation using the data of the previous problem to be

$$\frac{600 - 200}{1000 - 200} = \frac{V_2 - 0.1}{1.0 - 0.1} \longrightarrow V_2 = 0.55 \text{ m}^3$$

The mass contained in the system is

$$m = \frac{P_1 V_1}{RT_1} = \frac{(200 \text{ kPa})(0.1 \text{ m}^3)}{(0.2483 \text{ kPa} \cdot \text{m}^3/\text{kg} \cdot \text{K})(318 \text{ K})} = 0.2533 \text{ kg}$$

The final temperature is

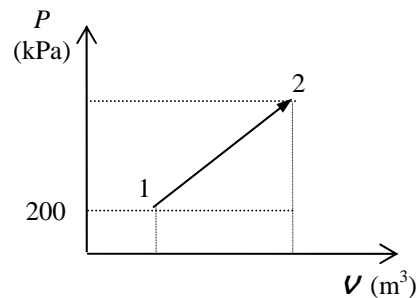
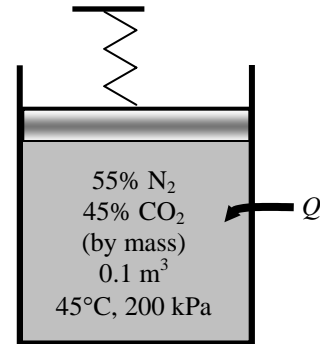
$$T_2 = \frac{P_2 V_2}{mR} = \frac{(600 \text{ kPa})(0.55 \text{ m}^3)}{(0.2533 \text{ kg})(0.2483 \text{ kPa} \cdot \text{m}^3/\text{kg} \cdot \text{K})} = 5247 \text{ K}$$

The work done during this process is

$$W_{\text{out}} = \frac{P_1 + P_2}{2} (V_2 - V_1) = \frac{(200 + 600) \text{ kPa}}{2} (0.55 - 0.1) \text{ m}^3 = \mathbf{180 \text{ kJ}}$$

An energy balance on the system gives

$$Q_{\text{in}} = W_{\text{out}} + mc_v(T_2 - T_1) = 180 + (0.2533 \text{ kg})(0.7043 \text{ kJ/kg} \cdot \text{K})(5247 - 318) \text{ K} = \mathbf{1059 \text{ kJ}}$$



13-99 The masses, pressures, and temperatures of the constituents of a gas mixture in a tank are given. Heat is transferred to the tank. The final pressure of the mixture and the heat transfer are to be determined.

Assumptions He is an ideal gas and O₂ is a nonideal gas.

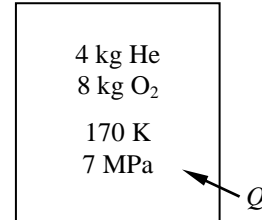
Properties The molar masses of He and O₂ are 4.0 and 32.0 kg/kmol. (Table A-1)

Analysis (a) The number of moles of each gas is

$$N_{\text{He}} = \frac{m_{\text{He}}}{M_{\text{He}}} = \frac{4 \text{ kg}}{4.0 \text{ kg/kmol}} = 1 \text{ kmol}$$

$$N_{\text{O}_2} = \frac{m_{\text{O}_2}}{M_{\text{O}_2}} = \frac{8 \text{ kg}}{32 \text{ kg/kmol}} = 0.25 \text{ kmol}$$

$$N_m = N_{\text{He}} + N_{\text{O}_2} = 1 \text{ kmol} + 0.25 \text{ kmol} = 1.25 \text{ kmol}$$



Then the partial volume of each gas and the volume of the tank are

He:

$$V_{\text{He}} = \frac{N_{\text{He}} R_u T_1}{P_{m,1}} = \frac{(1 \text{ kmol})(8.314 \text{ kPa} \cdot \text{m}^3/\text{kmol} \cdot \text{K})(170 \text{ K})}{7000 \text{ kPa}} = 0.202 \text{ m}^3$$

O₂:

$$\left. \begin{aligned} P_{R_1} &= \frac{P_{m,1}}{P_{\text{cr},\text{O}_2}} = \frac{7}{5.08} = 1.38 \\ T_{R_1} &= \frac{T_1}{T_{\text{cr},\text{O}_2}} = \frac{170}{154.8} = 1.10 \end{aligned} \right\} Z_1 = 0.53 \quad (\text{Fig. A-15})$$

$$V_{\text{O}_2} = \frac{Z_1 N_{\text{O}_2} R_u T_1}{P_{m,1}} = \frac{(0.53)(0.25 \text{ kmol})(8.314 \text{ kPa} \cdot \text{m}^3/\text{kg} \cdot \text{K})(170 \text{ K})}{7000 \text{ kPa}} = 0.027 \text{ m}^3$$

$$V_{\text{tank}} = V_{\text{He}} + V_{\text{O}_2} = 0.202 \text{ m}^3 + 0.027 \text{ m}^3 = 0.229 \text{ m}^3$$

The partial pressure of each gas and the total final pressure is

He:

$$P_{\text{He},2} = \frac{N_{\text{He}} R_u T_2}{V_{\text{tank}}} = \frac{(1 \text{ kmol})(8.314 \text{ kPa} \cdot \text{m}^3/\text{kmol} \cdot \text{K})(220 \text{ K})}{0.229 \text{ m}^3} = 7987 \text{ kPa}$$

O₂:

$$\left. \begin{aligned} T_{R_2} &= \frac{T_2}{T_{\text{cr},\text{O}_2}} = \frac{220}{154.8} = 1.42 \\ P_{R,\text{O}_2} &= \frac{\bar{v}_{\text{O}_2}}{R_u T_{\text{cr},\text{O}_2} / P_{\text{cr},\text{O}_2}} = \frac{V_m / N_{\text{O}_2}}{R_u T_{\text{cr},\text{O}_2} / P_{\text{cr},\text{O}_2}} \\ &= \frac{(0.229 \text{ m}^3)/(0.25 \text{ kmol})}{(8.314 \text{ kPa} \cdot \text{m}^3/\text{kmol} \cdot \text{K})(154.8 \text{ K})/(5080 \text{ kPa})} = 3.616 \end{aligned} \right\} P_R = 0.39 \quad (\text{Fig. A-15})$$

$$P_{\text{O}_2} = (P_R P_{\text{cr}})_{\text{O}_2} = (0.39)(5080 \text{ kPa}) = 1981 \text{ kPa} = 1.981 \text{ MPa}$$

$$P_{m,2} = P_{\text{He}} + P_{\text{O}_2} = 7.987 \text{ MPa} + 1.981 \text{ MPa} = \mathbf{9.97 \text{ MPa}}$$

(b) We take both gases as the system. No work or mass crosses the system boundary, therefore this is a closed system with no work interactions. Then the energy balance for this closed system reduces to

$$E_{\text{in}} - E_{\text{out}} = \Delta E_{\text{system}}$$

$$Q_{\text{in}} = \Delta U = \Delta U_{\text{He}} + \Delta U_{\text{O}_2}$$

He:

$$\Delta U_{\text{He}} = mc_v(T_m - T_1) = (4 \text{ kg})(3.1156 \text{ kJ/kg} \cdot \text{K})(220 - 170)\text{K} = 623.1 \text{ kJ}$$

O₂:

$$\left. \begin{array}{l} T_{R_1} = 1.10 \\ P_{R_1} = 1.38 \end{array} \right\} Z_{h_1} = 2.2$$

$$\left. \begin{array}{l} T_{R_2} = 1.42 \\ P_{R_2} = \frac{9.97}{5.08} = 1.963 \end{array} \right\} Z_{h_2} = 1.2 \quad (\text{Fig. A-29})$$

$$\begin{aligned} \bar{h}_2 - \bar{h}_1 &= R_u T_{\text{cr}} (Z_{h_1} - Z_{h_2}) + (\bar{h}_2 - \bar{h}_1)_{\text{ideal}} \\ &= (8.314 \text{ kJ/kmol} \cdot \text{K})(154.8 \text{ K})(2.2 - 1.2) + (6404 - 4949) \text{ kJ/kmol} \\ &= 2742 \text{ kJ/kmol} \end{aligned}$$

Also,

$$P_{\text{He},1} = \frac{N_{\text{He}} R_u T_1}{V_{\text{tank}}} = \frac{(1 \text{ kmol})(8.314 \text{ kPa} \cdot \text{m}^3/\text{kg} \cdot \text{K})(170 \text{ K})}{0.229 \text{ m}^3} = 6,172 \text{ kPa}$$

$$P_{\text{O}_2,1} = P_{m,1} - P_{\text{He},1} = 7000 \text{ kPa} - 6172 \text{ kPa} = 828 \text{ kPa}$$

Thus,

$$\begin{aligned} \Delta U_{\text{O}_2} &= N_{\text{O}_2} (\bar{h}_2 - \bar{h}_1) - (P_2 V_2 - P_1 V_1) = N_{\text{O}_2} (\bar{h}_2 - \bar{h}_1) - (P_{\text{O}_2,2} - P_{\text{O}_2,1}) V_{\text{tank}} \\ &= (0.25 \text{ kmol})(2742 \text{ kJ/kmol}) - (1981 - 828)(0.229) \text{ kPa} \cdot \text{m}^3 \\ &= 421.5 \text{ kJ} \end{aligned}$$

Substituting,

$$Q_{\text{in}} = 623.1 \text{ kJ} + 421.5 \text{ kJ} = \mathbf{1045 \text{ kJ}}$$

13-100 A mixture of carbon dioxide and methane expands through a turbine. The power produced by the mixture is to be determined using ideal gas approximation and Kay's rule.

Assumptions The expansion process is reversible and adiabatic (isentropic).

Properties The molar masses of CO₂ and CH₄ are 44.0 and 16.0 kg/kmol and respectively. The critical properties are 304.2 K, 7390 kPa for CO₂ and 191.1 K and 4640 kPa for CH₄ (Table A-1). EES may use slightly different values.

Analysis The molar mass of the mixture is determined to be

$$M_m = y_{\text{CO}_2} M_{\text{CO}_2} + y_{\text{CH}_4} M_{\text{CH}_4} = (0.75)(44) + (0.25)(16) = 37.0 \text{ kg/kmol}$$

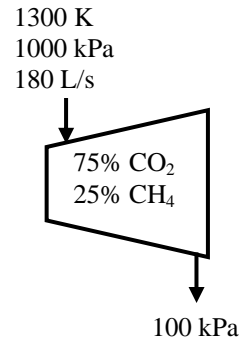
The gas constant is

$$R = \frac{R_u}{M_m} = \frac{8.314 \text{ kJ/kmol}\cdot\text{K}}{37.0 \text{ kg/kmol}} = 0.2246 \text{ kJ/kg}\cdot\text{K}$$

The mass fractions are

$$\text{mf}_{\text{CO}_2} = y_{\text{CO}_2} \frac{M_{\text{CO}_2}}{M_m} = (0.75) \frac{44 \text{ kg/kmol}}{37.0 \text{ kg/kmol}} = 0.8917$$

$$\text{mf}_{\text{CH}_4} = y_{\text{CH}_4} \frac{M_{\text{CH}_4}}{M_m} = (0.25) \frac{16 \text{ kg/kmol}}{37.0 \text{ kg/kmol}} = 0.1083$$



Ideal gas solution:

Using Dalton's law to find partial pressures, the entropies at the initial state are determined from EES to be:

$$T = 1300 \text{ K}, P = (0.75 \times 1000) = 750 \text{ kPa} \longrightarrow s_{\text{CO}_2,1} = 6.068 \text{ kJ/kg}\cdot\text{K}$$

$$T = 1600 \text{ K}, P = (0.25 \times 1000) = 250 \text{ kPa} \longrightarrow s_{\text{CH}_4,1} = 16.22 \text{ kJ/kg}\cdot\text{K}$$

The final state entropies cannot be determined at this point since the final temperature is not known. However, for an isentropic process, the entropy change is zero and the final temperature may be determined from

$$\begin{aligned} \Delta s_{\text{total}} &= \text{mf}_{\text{CO}_2} \Delta s_{\text{CO}_2} + \text{mf}_{\text{CH}_4} \Delta s_{\text{CH}_4} \\ 0 &= \text{mf}_{\text{CO}_2} (s_{\text{CO}_2,2} - s_{\text{CO}_2,1}) + \text{mf}_{\text{CH}_4} (s_{\text{CH}_4,2} - s_{\text{CH}_4,1}) \end{aligned}$$

The solution is obtained using EES to be

$$T_2 = 947.1 \text{ K}$$

The initial and final enthalpies and the changes in enthalpy are (from EES)

$$T_1 = 1300 \text{ K} \longrightarrow \begin{aligned} h_{\text{CO}_2,1} &= -7803 \text{ kJ/kg} \\ h_{\text{CH}_4,1} &= -831 \text{ kJ/kg} \end{aligned} \quad T_2 = 947.1 \text{ K} \longrightarrow \begin{aligned} h_{\text{CO}_2,2} &= -8248 \text{ kJ/kg} \\ h_{\text{CH}_4,2} &= -2503 \text{ kJ/kg} \end{aligned}$$

Noting that the heat transfer is zero, an energy balance on the system gives

$$\dot{Q}_{\text{in}} - \dot{W}_{\text{out}} = \dot{m} \Delta h_m \longrightarrow \dot{W}_{\text{out}} = -\dot{m} \Delta h_m$$

where

$$\begin{aligned} \Delta h_m &= \text{mf}_{\text{CO}_2} (h_{\text{CO}_2,2} - h_{\text{CO}_2,1}) + \text{mf}_{\text{CH}_4} (h_{\text{CH}_4,2} - h_{\text{CH}_4,1}) \\ &= (0.8917)[(-8248) - (-7803)] + (0.1083)[(-2503) - (-831)] = -577.7 \text{ kJ/kg} \end{aligned}$$

The mass flow rate is

$$\dot{m} = \frac{P_1 \dot{V}_1}{RT_1} = \frac{(1000 \text{ kPa})(0.180 \text{ m}^3/\text{s})}{(0.2246 \text{ kJ/kg}\cdot\text{K})(1300 \text{ K})} = 0.6165 \text{ kg/s}$$

$$\text{Substituting, } \dot{W}_{\text{out}} = \dot{m} \Delta h_m = -(0.6165)(-577.7 \text{ kJ/kg}) = \mathbf{356 \text{ kW}}$$

Kay's rule solution:

The critical temperature and pressure of the mixture is

$$T_{cr} = y_{CO_2} T_{cr,CO_2} + y_{CH_4} T_{cr,CH_4} = (0.75)(304.2 \text{ K}) + (0.25)(191.1 \text{ K}) = 276 \text{ K}$$

$$P_{cr} = y_{CO_2} P_{cr,CO_2} + y_{CH_4} P_{cr,CH_4} = (0.75)(7390 \text{ kPa}) + (0.25)(4640 \text{ kPa}) = 6683 \text{ kPa}$$

State 1 properties:

$$\left. \begin{aligned} T_{R1} &= \frac{T_1}{T_{cr}} = \frac{1300 \text{ K}}{276 \text{ K}} = 4.715 \\ P_{R1} &= \frac{P_1}{P_{cr}} = \frac{1000 \text{ kPa}}{6683 \text{ kPa}} = 0.1496 \end{aligned} \right\} \begin{aligned} Z_1 &= 1.003 \\ Z_{h1} &= -0.0064 \\ Z_{s1} &= 0.001197 \end{aligned} \quad (\text{from EES})$$

$$\Delta h_1 = Z_{h1} RT_{cr} = (-0.0064)(0.2246 \text{ kJ/kg}\cdot\text{K})(276 \text{ K}) = -0.399 \text{ kJ/kg}$$

$$\begin{aligned} h_1 &= mf_{CO_2} h_{CO_2,1} + mf_{CH_4} h_{CH_4,1} - \Delta h_1 \\ &= (0.8917)(-7803) + (0.1083)(-831) - (-0.399) = -7047 \text{ kJ/kg} \end{aligned}$$

$$\Delta s_1 = Z_{s1} R = (0.001197)(0.2246 \text{ kJ/kg}\cdot\text{K}) = 0.0002688 \text{ kJ/kg}\cdot\text{K}$$

$$\begin{aligned} s_1 &= mf_{CO_2} s_{CO_2,1} + mf_{CH_4} s_{CH_4,1} - \Delta s_1 \\ &= (0.8917)(6.068) + (0.1083)(16.22) - (0.0002688) = 7.1679 \text{ kJ/kg}\cdot\text{K} \end{aligned}$$

The final state entropies cannot be determined at this point since the final temperature is not known. However, for an isentropic process, the entropy change is zero and the final temperature may be determined from

$$\begin{aligned} \Delta s_{total} &= mf_{CO_2} \Delta s_{CO_2} + mf_{CH_4} \Delta s_{CH_4} \\ 0 &= mf_{CO_2} (s_{CO_2,2} - s_{CO_2,1}) + mf_{CH_4} (s_{CH_4,2} - s_{CH_4,1}) \end{aligned}$$

The solution is obtained using EES to be

$$T_2 = 947 \text{ K}$$

The initial and final enthalpies and the changes in enthalpy are

$$\left. \begin{aligned} T_{R2} &= \frac{T_2}{T_{cr}} = \frac{947 \text{ K}}{276 \text{ K}} = 3.434 \\ P_{R2} &= \frac{P_2}{P_{cr}} = \frac{100 \text{ kPa}}{6683 \text{ kPa}} = 0.015 \end{aligned} \right\} \begin{aligned} Z_{h2} &= -0.0004869 \\ Z_{s2} &= 0.0004057 \end{aligned} \quad (\text{from EES})$$

$$\Delta h_2 = Z_{h2} RT_{cr} = (-0.0004869)(0.2246 \text{ kJ/kg}\cdot\text{K})(276 \text{ K}) = -0.03015 \text{ kJ/kg}$$

$$\begin{aligned} h_2 &= mf_{CO_2} h_{CO_2,2} + mf_{CH_4} h_{CH_4,2} - \Delta h_2 \\ &= (0.8917)(-8248) + (0.1083)(-2503) - (-0.03015) = -7625 \text{ kJ/kg} \end{aligned}$$

Noting that the heat transfer is zero, an energy balance on the system gives

$$\dot{Q}_{in} - \dot{W}_{out} = \dot{m} \Delta h_m \longrightarrow \dot{W}_{out} = -\dot{m}(h_2 - h_1)$$

where the mass flow rate is

$$\dot{m} = \frac{P_1 \dot{V}_1}{Z_1 RT_1} = \frac{(1000 \text{ kPa})(0.180 \text{ m}^3/\text{s})}{(1.003)(0.2246 \text{ kJ/kg}\cdot\text{K})(1300 \text{ K})} = 0.6149 \text{ kg/s}$$

Substituting,

$$\dot{W}_{out} = -(0.6149 \text{ kg/s})[(-7625) - (-7047) \text{ kJ/kg}] = \mathbf{356 \text{ kW}}$$

13-101 A stream of gas mixture at a given pressure and temperature is to be separated into its constituents steadily. The minimum work required is to be determined.

Assumptions **1** Both the N_2 and CO_2 gases and their mixture are ideal gases. **2** This is a steady-flow process. **3** The kinetic and potential energy changes are negligible.

Properties The molar masses of N_2 and CO_2 are 28.0 and 44.0 kg/kmol. (Table A-1).

Analysis The minimum work required to separate a gas mixture into its components is equal to the reversible work associated with the mixing process, which is equal to the exergy destruction (or irreversibility) associated with the mixing process since

$$X_{\text{destroyed}} = W_{\text{rev,out}} - W_{\text{act,u}}^{\text{70}} = W_{\text{rev,out}} = T_0 S_{\text{gen}}$$

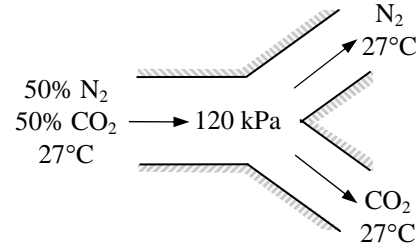
where S_{gen} is the entropy generation associated with the steady-flow mixing process. The entropy change associated with a constant pressure and temperature adiabatic mixing process is determined from

$$\begin{aligned}\bar{s}_{\text{gen}} &= \sum \Delta \bar{s}_i = -R_u \sum y_i \ln y_i = -(8.314 \text{ kJ/kmol} \cdot \text{K}) [0.5 \ln(0.5) + 0.5 \ln(0.5)] \\ &= 5.763 \text{ kJ/kmol} \cdot \text{K}\end{aligned}$$

$$M_m = \sum y_i M_i = (0.5)(28 \text{ kg/kmol}) + (0.5)(44 \text{ kg/kmol}) = 36 \text{ kg/kmol}$$

$$s_{\text{gen}} = \frac{\bar{s}_{\text{gen}}}{M_m} = \frac{5.763 \text{ kJ/kmol} \cdot \text{K}}{36 \text{ kg/kmol}} = 0.1601 \text{ kJ/kg} \cdot \text{K}$$

$$x_{\text{destroyed}} = T_0 s_{\text{gen}} = (300 \text{ K})(0.1601 \text{ kJ/kg} \cdot \text{K}) = \mathbf{48.0 \text{ kJ/kg}}$$



13-102E The mass percentages of a gas mixture are given. This mixture is expanded in an adiabatic, steady-flow turbine of specified isentropic efficiency. The second law efficiency and the exergy destruction during this expansion process are to be determined.

Assumptions All gases will be modeled as ideal gases with constant specific heats.

Properties The molar masses of N_2 , He, CH_4 , and C_2H_6 are 28.0, 4.0, 16.0, and 30.0 lbm/lbmol, respectively (Table A-1E). The constant-pressure specific heats of these gases at room temperature are 0.248, 1.25, 0.532, and 0.427 Btu/lbm·R, respectively (Table A-2Ea).

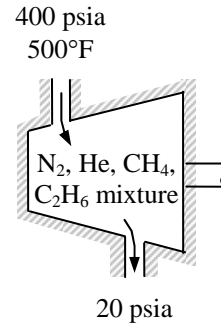
Analysis For 1 lbm of mixture, the mole numbers of each component are

$$N_{N_2} = \frac{m_{N_2}}{M_{N_2}} = \frac{0.15 \text{ lbm}}{28 \text{ lbm/lbmol}} = 0.005357 \text{ lbmol}$$

$$N_{He} = \frac{m_{He}}{M_{He}} = \frac{0.05 \text{ lbm}}{4 \text{ lbm/lbmol}} = 0.0125 \text{ lbmol}$$

$$N_{CH_4} = \frac{m_{CH_4}}{M_{CH_4}} = \frac{0.6 \text{ lbm}}{16 \text{ lbm/lbmol}} = 0.0375 \text{ lbmol}$$

$$N_{C_2H_6} = \frac{m_{C_2H_6}}{M_{C_2H_6}} = \frac{0.20 \text{ lbm}}{30 \text{ lbm/lbmol}} = 0.006667 \text{ lbmol}$$



The mole number of the mixture is

$$N_m = N_{O_2} + N_{CO_2} + N_{He} = 0.005357 + 0.0125 + 0.0375 + 0.006667 = 0.06202 \text{ lbmol}$$

The apparent molecular weight of the mixture is

$$M_m = \frac{m_m}{N_m} = \frac{1 \text{ lbm}}{0.06202 \text{ lbmol}} = 16.12 \text{ lbm/lbmol}$$

The apparent gas constant of the mixture is

$$R = \frac{R_u}{M_m} = \frac{1.9858 \text{ lbm/lbmol} \cdot \text{R}}{16.12 \text{ lbm/lbmol}} = 0.1232 \text{ Btu/lbm} \cdot \text{R}$$

The constant-pressure specific heat of the mixture is determined from

$$\begin{aligned} c_p &= mf_{N_2} c_{p,N_2} + mf_{He} c_{p,He} + mf_{CH_4} c_{p,CH_4} + mf_{C_2H_6} c_{p,C_2H_6} \\ &= 0.15 \times 0.248 + 0.05 \times 1.25 + 0.60 \times 0.532 + 0.20 \times 0.427 \\ &= 0.5043 \text{ Btu/lbm} \cdot \text{R} \end{aligned}$$

Then the constant-volume specific heat is

$$c_v = c_p - R = 0.5043 - 0.1232 = 0.3811 \text{ Btu/lbm} \cdot \text{R}$$

The specific heat ratio is

$$k = \frac{c_p}{c_v} = \frac{0.5043}{0.3811} = 1.323$$

The temperature at the end of the expansion for the isentropic process is

$$T_{2s} = T_1 \left(\frac{P_2}{P_1} \right)^{(k-1)/k} = (960 \text{ R}) \left(\frac{20 \text{ psia}}{400 \text{ psia}} \right)^{0.323/1.323} = 462.0 \text{ R}$$

Using the definition of turbine isentropic efficiency, the actual outlet temperature is

$$T_2 = T_1 - \eta_{\text{turb}} (T_1 - T_{2s}) = (960 \text{ R}) - (0.85)(960 - 462.0) = 536.7 \text{ R}$$

The entropy change of the gas mixture is

$$s_2 - s_1 = c_p \ln \frac{T_2}{T_1} - R \ln \frac{P_2}{P_1} = (0.5043) \ln \frac{536.7}{960} - (0.1232) \ln \frac{20}{400} = 0.07583 \text{ Btu/lbm} \cdot \text{R}$$

The actual work produced is

$$w_{\text{out}} = h_1 - h_2 = c_p (T_1 - T_2) = (0.5043 \text{ Btu/lbm} \cdot \text{R})(960 - 536.7) \text{ R} = 213.5 \text{ Btu/lbm}$$

The reversible work output is

$$w_{\text{rev,out}} = h_1 - h_2 - T_0 (s_1 - s_2) = 213.5 \text{ Btu/lbm} - (537 \text{ R})(-0.07583 \text{ Btu/lbm} \cdot \text{R}) = 254.2 \text{ Btu/lbm}$$

The second-law efficiency and the exergy destruction are then

$$\eta_{\text{II}} = \frac{w_{\text{out}}}{w_{\text{rev,out}}} = \frac{213.5}{254.2} = 0.840 = \mathbf{84.0\%}$$

$$x_{\text{dest}} = w_{\text{rev,out}} - w_{\text{out}} = 254.2 - 213.5 = \mathbf{40.7 \text{ Btu/lbm}}$$



13-103 A program is to be written to determine the mole fractions of the components of a mixture of three gases with known molar masses when the mass fractions are given, and to determine the mass fractions of the components when the mole fractions are given. Also, the program is to be run for a sample case.

Analysis The problem is solved using EES, and the solution is given below.

```

Procedure Fractions(Type$,A$,B$,C$,A,B,C:mf_A,mf_B,mf_C,y_A,y_B,y_C)
{If Type$ <> ('mass fraction' OR 'mole fraction') then
Call ERROR('Type$ must be set equal to "mass fraction" or "mole fraction".')
GOTO 10
endif}
Sum = A+B+C
If ABS(Sum - 1) > 0 then goto 20
MM_A = molarmass(A$)
MM_B = molarmass(B$)
MM_C = molarmass(C$)
If Type$ = 'mass fraction' then
mf_A = A
mf_B = B
mf_C = C
sumM_mix = mf_A/MM_A+ mf_B/MM_B+ mf_C/MM_C
y_A = mf_A/MM_A/sumM_mix
y_B = mf_B/MM_B/sumM_mix
y_C = mf_C/MM_C/sumM_mix
GOTO 10
endif
if Type$ = 'mole fraction' then
y_A = A
y_B = B
y_C = C
MM_mix = y_A*MM_A+ y_B*MM_B+ y_C*MM_C
mf_A = y_A*MM_A/MM_mix
mf_B = y_B*MM_B/MM_mix
mf_C = y_C*MM_C/MM_mix
GOTO 10
endif
Call ERROR('Type$ must be either mass fraction or mole fraction.')
GOTO 10
20:
Call ERROR('The sum of the mass or mole fractions must be 1')
10:
END

```

"Either the mole fraction y_i or the mass fraction mf_i may be given by setting the parameter $Type$='mole fraction'$ when the mole fractions are given or $Type$='mass fraction'$ is given"

```

{Input Data in the Diagram Window}
{Type$='mole fraction'
A$ = 'N2'
B$ = 'O2'
C$ = 'Argon'
A = 0.71 "When Type$='mole fraction' A, B, C are the mole fractions"
B = 0.28 "When Type$='mass fraction' A, B, C are the mass fractions"
C = 0.01}
Call Fractions(Type$,A$,B$,C$,A,B,C:mf_A,mf_B,mf_C,y_A,y_B,y_C)

```

SOLUTION

A=0.71	A\$='N2'	B=0.28	B\$='O2'
C=0.01	C\$='Argon'	mf_A=0.680	mf_B=0.306
mf_C=0.014	Type\$='mole fraction'		y_A=0.710
y_B=0.280	y_C=0.010		



13-104 A program is to be written to determine the entropy change of a mixture of 3 ideal gases when the mole fractions and other properties of the constituent gases are given. Also, the program is to be run for a sample case.

Analysis The problem is solved using EES, and the solution is given below.

```

T1=300 [K]
T2=600 [K]
P1=100 [kPa]
P2=500 [kPa]
A$ = 'N2'
B$ = 'O2'
C$ = 'Argon'
y_A = 0.71
y_B = 0.28
y_C = 0.01
MM_A = molarmass(A$)
MM_B = molarmass(B$)
MM_C = molarmass(C$)
MM_mix = y_A*MM_A+ y_B*MM_B+ y_C*MM_C
mf_A = y_A*MM_A/MM_mix
mf_B = y_B*MM_B/MM_mix
mf_C = y_C*MM_C/MM_mix
DELTAs_mix=mf_A*(entropy(A$,T=T2,P=y_B*P2)-
entropy(A$,T=T1,P=y_A*P1))+mf_B*(entropy(B$,T=T2,P=y_B*P2)-
entropy(B$,T=T1,P=y_B*P1))+mf_C*(entropy(C$,T=T2,P=y_C*P2)-entropy(C$,T=T1,P=y_C*P1))

```

SOLUTION

```

A$='N2'
B$='O2'
C$='Argon'
DELTAs_mix=12.41 [kJ/kg-K]
mf_A=0.68
mf_B=0.3063
mf_C=0.01366
MM_A=28.01 [kg/kmol]
MM_B=32 [kg/kmol]
MM_C=39.95 [kg/kmol]
MM_mix=29.25 [kJ/kmol]
P1=100 [kPa]
P2=500 [kPa]
T1=300 [K]
T2=600 [K]
y_A=0.71
y_B=0.28
y_C=0.01

```

Fundamentals of Engineering (FE) Exam Problems

13-105 An ideal gas mixture whose apparent molar mass is 20 kg/kmol consists of nitrogen N_2 and three other gases. If the mole fraction of nitrogen is 0.55, its mass fraction is

- (a) 0.15 (b) 0.23 (c) 0.39 (d) 0.55 (e) 0.77

Answer (e) 0.77

Solution Solved by EES Software. Solutions can be verified by copying-and-pasting the following lines on a blank EES screen. (Similar problems and their solutions can be obtained easily by modifying numerical values).

```
M_mix=20 "kg/kmol"
M_N2=28 "kg/kmol"
y_N2=0.55
mf_N2=(M_N2/M_mix)*y_N2
```

"Some Wrong Solutions with Common Mistakes:"

```
W1_mf = y_N2 "Taking mass fraction to be equal to mole fraction"
W2_mf= y_N2*(M_mix/M_N2) "Using the molar mass ratio backwards"
W3_mf= 1-mf_N2 "Taking the complement of the mass fraction"
```

13-106 An ideal gas mixture consists of 2 kmol of N_2 and 6 kmol of CO_2 . The mass fraction of CO_2 in the mixture is

- (a) 0.175 (b) 0.250 (c) 0.500 (d) 0.750 (e) 0.825

Answer (e) 0.825

Solution Solved by EES Software. Solutions can be verified by copying-and-pasting the following lines on a blank EES screen. (Similar problems and their solutions can be obtained easily by modifying numerical values).

```
N1=2 "kmol"
N2=6 "kmol"
N_mix=N1+N2
MM1=28 "kg/kmol"
MM2=44 "kg/kmol"
m_mix=N1*MM1+N2*MM2
mf2=N2*MM2/m_mix
```

"Some Wrong Solutions with Common Mistakes:"

```
W1_mf = N2/N_mix "Using mole fraction"
W2_mf = 1-mf2 "The wrong mass fraction"
```

13-107 An ideal gas mixture consists of 2 kmol of N_2 and 4 kmol of CO_2 . The apparent gas constant of the mixture is

- (a) 0.215 kJ/kg·K (b) 0.225 kJ/kg·K (c) 0.243 kJ/kg·K (d) 0.875 kJ/kg·K (e) 1.24 kJ/kg·K

Answer (a) 0.215 kJ/kg·K

Solution Solved by EES Software. Solutions can be verified by copying-and-pasting the following lines on a blank EES screen. (Similar problems and their solutions can be obtained easily by modifying numerical values).

```
Ru=8.314 "kJ/kmol.K"
N1=2 "kmol"
N2=4 "kmol"
MM1=28 "kg/kmol"
MM2=44 "kg/kmol"
R1=Ru/MM1
R2=Ru/MM2
N_mix=N1+N2
y1=N1/N_mix
y2=N2/N_mix
MM_mix=y1*MM1+y2*MM2
R_mix=Ru/MM_mix
```

"Some Wrong Solutions with Common Mistakes:"

W1_Rmix=(R1+R2)/2 "Taking the arithmetic average of gas constants"

W2_Rmix= y1*R1+y2*R2 "Using wrong relation for Rmixture"

13-108 A rigid tank is divided into two compartments by a partition. One compartment contains 3 kmol of N_2 at 400 kPa pressure and the other compartment contains 7 kmol of CO_2 at 200 kPa. Now the partition is removed, and the two gases form a homogeneous mixture at 250 kPa. The partial pressure of N_2 in the mixture is

- (a) 75 kPa (b) 90 kPa (c) 125 kPa (d) 175 kPa (e) 250 kPa

Answer (a) 75 kPa

Solution Solved by EES Software. Solutions can be verified by copying-and-pasting the following lines on a blank EES screen. (Similar problems and their solutions can be obtained easily by modifying numerical values).

```
P1 = 400 "kPa"
P2 = 200 "kPa"
P_mix=250 "kPa"
N1=3 "kmol"
N2=7 "kmol"
MM1=28 "kg/kmol"
MM2=44 "kg/kmol"
N_mix=N1+N2
y1=N1/N_mix
y2=N2/N_mix
P_N2=y1*P_mix
```

"Some Wrong Solutions with Common Mistakes:"

W1_P1= P_mix/2 "Assuming equal partial pressures"

W2_P1= mf1*P_mix; mf1=N1*MM1/(N1*MM1+N2*MM2) "Using mass fractions"

W3_P1 = P_mix*N1*P1/(N1*P1+N2*P2) "Using some kind of weighed averaging"

13-109 An 80-L rigid tank contains an ideal gas mixture of 5 g of N₂ and 5 g of CO₂ at a specified pressure and temperature. If N₂ were separated from the mixture and stored at mixture temperature and pressure, its volume would be

- (a) 32 L (b) 36 L (c) 40 L (d) 49 L (e) 80 L

Answer (d) 49 L

Solution Solved by EES Software. Solutions can be verified by copying-and-pasting the following lines on a blank EES screen. (Similar problems and their solutions can be obtained easily by modifying numerical values).

```
V_mix=80 "L"
m1=5 "g"
m2=5 "g"
MM1=28 "kg/kmol"
MM2=44 "kg/kmol"
N1=m1/MM1
N2=m2/MM2
N_mix=N1+N2
y1=N1/N_mix
V1=y1*V_mix "L"
```

"Some Wrong Solutions with Common Mistakes:"

```
W1_V1=V_mix*m1/(m1+m2) "Using mass fractions"
W2_V1= V_mix "Assuming the volume to be the mixture volume"
```

13-110 An ideal gas mixture consists of 3 kg of Ar and 6 kg of CO₂ gases. The mixture is now heated at constant volume from 250 K to 350 K. The amount of heat transfer is

- (a) 374 kJ (b) 436 kJ (c) 488 kJ (d) 525 kJ (e) 664 kJ

Answer (c) 488 kJ

Solution Solved by EES Software. Solutions can be verified by copying-and-pasting the following lines on a blank EES screen. (Similar problems and their solutions can be obtained easily by modifying numerical values).

```
T1=250 "K"
T2=350 "K"
Cv1=0.3122; Cp1=0.5203 "kJ/kg.K"
Cv2=0.657; Cp2=0.846 "kJ/kg.K"
m1=3 "kg"
m2=6 "kg"
MM1=39.95 "kg/kmol"
MM2=44 "kg/kmol"
"Applying Energy balance gives Q=DeltaU=DeltaU_Ar+DeltaU_CO2"
Q=(m1*Cv1+m2*Cv2)*(T2-T1)
```

"Some Wrong Solutions with Common Mistakes:"

```
W1_Q = (m1+m2)*(Cv1+Cv2)/2*(T2-T1) "Using arithmetic average of properties"
W2_Q = (m1*Cp1+m2*Cp2)*(T2-T1)"Using Cp instead of Cv"
W3_Q = (m1*Cv1+m2*Cv2)*T2 "Using T2 instead of T2-T1"
```

13-111 An ideal gas mixture consists of 60% helium and 40% argon gases by mass. The mixture is now expanded isentropically in a turbine from 400°C and 1.2 MPa to a pressure of 200 kPa. The mixture temperature at turbine exit is

- (a) 56°C (b) 195°C (c) 130°C (d) 112°C (e) 400°C

Answer (a) 56°C

Solution Solved by EES Software. Solutions can be verified by copying-and-pasting the following lines on a blank EES screen. (Similar problems and their solutions can be obtained easily by modifying numerical values).

T1=400+273"K"

P1=1200 "kPa"

P2=200 "kPa"

mf_He=0.6

mf_Ar=0.4

k1=1.667

k2=1.667

"The specific heat ratio k of the mixture is also 1.667 since k=1.667 for all componet gases"

k_mix=1.667

T2=T1*(P2/P1)^((k_mix-1)/k_mix)-273

"Some Wrong Solutions with Common Mistakes:"

W1_T2 = (T1-273)*(P2/P1)^((k_mix-1)/k_mix) "Using C for T1 instead of K"

W2_T2 = T1*(P2/P1)^((k_air-1)/k_air)-273; k_air=1.4 "Using k value for air"

W3_T2 = T1*P2/P1 "Assuming T to be proportional to P"

13-112 One compartment of an insulated rigid tank contains 2 kmol of CO₂ at 20°C and 150 kPa while the other compartment contains 5 kmol of H₂ gas at 35°C and 300 kPa. Now the partition between the two gases is removed, and the two gases form a homogeneous ideal gas mixture. The temperature of the mixture is

- (a) 25°C (b) 29°C (c) 22°C (d) 32°C (e) 34°C

Answer (b) 29°C

Solution Solved by EES Software. Solutions can be verified by copying-and-pasting the following lines on a blank EES screen. (Similar problems and their solutions can be obtained easily by modifying numerical values).

```
N_H2=5 "kmol"
T1_H2=35 "C"
P1_H2=300 "kPa"
N_CO2=2 "kmol"
T1_CO2=20 "C"
P1_CO2=150 "kPa"
Cv_H2=10.183; Cp_H2=14.307 "kJ/kg.K"
Cv_CO2=0.657; Cp_CO2=0.846 "kJ/kg.K"
MM_H2=2 "kg/kmol"
MM_CO2=44 "kg/kmol"
m_H2=N_H2*MM_H2
m_CO2=N_CO2*MM_CO2
"Applying Energy balance gives 0=DeltaU=DeltaU_H2+DeltaU_CO2"
0=m_H2*Cv_H2*(T2-T1_H2)+m_CO2*Cv_CO2*(T2-T1_CO2)
```

"Some Wrong Solutions with Common Mistakes:"

```
0=m_H2*Cp_H2*(W1_T2-T1_H2)+m_CO2*Cp_CO2*(W1_T2-T1_CO2) "Using Cp instead of Cv"
0=N_H2*Cv_H2*(W2_T2-T1_H2)+N_CO2*Cv_CO2*(W2_T2-T1_CO2) "Using N instead of mass"
W3_T2 = (T1_H2+T1_CO2)/2 "Assuming average temperature"
```

13-113 A piston-cylinder device contains an ideal gas mixture of 3 kmol of He gas and 7 kmol of Ar gas at 50°C and 400 kPa. Now the gas expands at constant pressure until its volume doubles. The amount of heat transfer to the gas mixture is

- (a) 6.2 MJ (b) 42 MJ (c) 27 MJ (d) 10 MJ (e) 67 MJ

Answer (e) 67 MJ

Solution Solved by EES Software. Solutions can be verified by copying-and-pasting the following lines on a blank EES screen. (Similar problems and their solutions can be obtained easily by modifying numerical values).

```
N_He=3 "kmol"
N_Ar=7 "kmol"
T1=50+273 "C"
P1=400 "kPa"
P2=P1
"T2=2T1 since PV/T=const for ideal gases and it is given that P=constant"
T2=2*T1 "K"
MM_He=4 "kg/kmol"
MM_Ar=39.95 "kg/kmol"
m_He=N_He*MM_He
m_Ar=N_Ar*MM_Ar
Cp_Ar=0.5203; Cv_Ar = 3122 "kJ/kg.C"
Cp_He=5.1926; Cv_He = 3.1156 "kJ/kg.K"
"For a P=const process, Q=DeltaH since DeltaU+Wb is DeltaH"
Q=m_Ar*Cp_Ar*(T2-T1)+m_He*Cp_He*(T2-T1)
```

"Some Wrong Solutions with Common Mistakes:"

```
W1_Q =m_Ar*Cv_Ar*(T2-T1)+m_He*Cv_He*(T2-T1) "Using Cv instead of Cp"
W2_Q=N_Ar*Cp_Ar*(T2-T1)+N_He*Cp_He*(T2-T1) "Using N instead of mass"
W3_Q=m_Ar*Cp_Ar*(T2-T1)+m_He*Cp_He*(T2-T1); T2=2*(T1-273)+273 "Using C for T1"
W4_Q=(m_Ar+m_He)*0.5*(Cp_Ar+Cp_He)*(T2-T1) "Using arithmetic average of Cp"
```

13-114 An ideal gas mixture of helium and argon gases with identical mass fractions enters a turbine at 1500 K and 1 MPa at a rate of 0.12 kg/s, and expands isentropically to 100 kPa. The power output of the turbine is

- (a) 253 kW (b) 310 kW (c) 341 kW (d) 463 kW (e) 550 kW

Answer (b) 310 kW

Solution Solved by EES Software. Solutions can be verified by copying-and-pasting the following lines on a blank EES screen. (Similar problems and their solutions can be obtained easily by modifying numerical values).

```
m=0.12 "kg/s"
T1=1500 "K"
P1=1000 "kPa"
P2=100 "kPa"
mf_He=0.5
mf_Ar=0.5
k_He=1.667
k_Ar=1.667
Cp_Ar=0.5203
Cp_He=5.1926
Cp_mix=mf_He*Cp_He+mf_Ar*Cp_Ar
"The specific heat ratio k of the mixture is also 1.667 since k=1.667 for all componet gases"
k_mix=1.667
T2=T1*(P2/P1)^((k_mix-1)/k_mix)
-W_out=m*Cp_mix*(T2-T1)
```

"Some Wrong Solutions with Common Mistakes:"

```
W1_Wout= - m*Cp_mix*(T22-T1); T22 = (T1-273)*(P2/P1)^((k_mix-1)/k_mix)+273 "Using C for T1 instead of K"
W2_Wout= - m*Cp_mix*(T222-T1); T222 = T1*(P2/P1)^((k_air-1)/k_air)-273; k_air=1.4 "Using k value for air"
W3_Wout= - m*Cp_mix*(T2222-T1); T2222 = T1*P2/P1 "Assuming T to be proportional to P"
W4_Wout= - m*0.5*(Cp_Ar+Cp_He)*(T2-T1) "Using arithmetic average for Cp"
```

13-115 ... 13-118 Design and Essay Problem

